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FILE LAST UPDATED: 2 May 2006 (20060502/ED)

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=> d stat que
L1      795 SEA FILE=REGISTRY ABB=ON  PLU=ON  TETRAHYDROBORATE OR HYDROBORA
              TE
L2      11  SEA FILE=REGISTRY ABB=ON  PLU=ON  SODIUM BOROHYDRATE?/CN OR
              LITHIUM BOROHYDR?/CN OR POTASSIUM BOROHYDR?/CN
L3      104214 SEA FILE=REGISTRY ABB=ON  PLU=ON  METAL OR METALS OR MAGNESIUM

L4      128483 SEA FILE=REGISTRY ABB=ON  PLU=ON  BORATE
L5      15523 SEA FILE=HCAPLUS ABB=ON  PLU=ON  L1 OR TETRAHYDROBORATE OR
              HYDROBORATE
L6      SEL  PLU=ON  L2 1- CHEM :      52 TERMS
L7      15422 SEA FILE=HCAPLUS ABB=ON  PLU=ON  L6
L8      15583 SEA FILE=HCAPLUS ABB=ON  PLU=ON  L7 OR (SODIUM OR LITHIUM OR
              POTASSIUM) (2A) BOROHYDR?
L9      4137887 SEA FILE=HCAPLUS ABB=ON  PLU=ON  L3 OR METAL OR MAGNESIUM
L10     166954 SEA FILE=HCAPLUS ABB=ON  PLU=ON  L4 OR BORATE
L12     15254 SEA FILE=HCAPLUS ABB=ON  PLU=ON  L10 (L) L9
L14     38473 SEA FILE=HCAPLUS ABB=ON  PLU=ON  L10 (L) (RCT/RL OR RACT/RL)
L15     250013 SEA FILE=HCAPLUS ABB=ON  PLU=ON  L9 (L) (RCT/RL OR RACT/RL)
L16     2579  SEA FILE=HCAPLUS ABB=ON  PLU=ON  L12 AND L14
L17     1952  SEA FILE=HCAPLUS ABB=ON  PLU=ON  L16 AND L15
L18     104576 SEA FILE=REGISTRY ABB=ON  PLU=ON  PALLADIUM/BI
L19      208 SEA FILE=REGISTRY ABB=ON  PLU=ON  PT/MF
L21     513099 SEA FILE=HCAPLUS ABB=ON  PLU=ON  L18 OR L19 OR PALLADIUM OR
              PLATINUM OR PD OR PT
L25     3041  SEA FILE=HCAPLUS ABB=ON  PLU=ON  (L5 OR L8) (L) (PREP?/RL OR
              PREP OR PREPARTION)
L26      78  SEA FILE=HCAPLUS ABB=ON  PLU=ON  L17 AND L25
L27      11  SEA FILE=HCAPLUS ABB=ON  PLU=ON  L26 AND L21
L28      9  SEA FILE=HCAPLUS ABB=ON  PLU=ON  L27 AND PD=<NOVEMBER 25, 2003
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Valenrod 10_721479

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=>
=> d ibib abs hitstr l28 1-9

L28 ANSWER 1 OF 9 HCAPLUS COPYRIGHT 2006 ACS on STN
 ACCESSION NUMBER: 2002:475940 HCAPLUS
 DOCUMENT NUMBER: 138:26390
 TITLE: Hydrogen generation using sodium borohydride solution
 and metal catalyst coated on metal oxide
 AUTHOR(S): Kojima, Yoshitsugu; Suzuki, Ken-ichirou; Fukumoto,
 Kazuhiro; Sasaki, Megumi; Yamamoto, Toshio; Kawai,
 Yasuaki; Hayashi, Hiroaki
 CORPORATE SOURCE: Toyota Central Research & Development Laboratories,
 Inc., Nagakute-cho, Aichi-gun, Aichi, 480-1192, Japan
 SOURCE: International Journal of Hydrogen Energy (2002
), 27(10), 1029-1034
 CODEN: IJHEDX; ISSN: 0360-3199
 PUBLISHER: Elsevier Science Ltd.
 DOCUMENT TYPE: Journal
 LANGUAGE: English
 AB Sodium borohydride (NaBH4) reacted slowly with water to liberate 4 mol of
 hydrogen/mol of the compound at room temperature. Hydrogen generation was
 accelerated by applying metal-metal oxide catalysts such as Pt
 -TiO2, Pt-CoO and Pt-LiCoO2. As the size of metal
 crystallites decreased and the amount increased, the hydrogen generation
 rate increased. The hydrogen generation rates using Pt and
 LiCoO2 were high compared with those using other metal and metal oxide,
 resp. It seemed that a key finding was that use of the supercrit. CO2
 method produced a superior catalyst. Borohydride ion was stabilized in
 alkaline solution containing at least 5 weight% of NaOH. Alkaline stabilized
 solution of NaBH4
 can be applied as a hydrogen source. We found that Pt-LiCoO2
 worked as an excellent catalyst for releasing hydrogen from the stabilized
 NaBH4 solution
 IT 7440-05-3, Palladium, uses 7440-06-4,
 Platinum, uses
 RL: CAT (Catalyst use); USES (Uses)
 (hydrogen generation by reaction of sodium borohydride with water using
 metal catalyst coated on metal oxide)
 RN 7440-05-3 HCAPLUS
 CN Palladium (8CI, 9CI) (CA INDEX NAME)

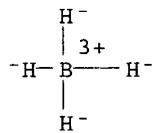
Pd

RN 7440-06-4 HCAPLUS
 CN Platinum (8CI, 9CI) (CA INDEX NAME)

Pt

IT 16940-66-2, Sodium borohydride
 RL: CPS (Chemical process); PEP (Physical, engineering or chemical
 process); RCT (Reactant); PROC (Process); RACT (Reactant or
 reagent)
 (hydrogen generation by reaction of sodium borohydride with water using
 metal catalyst coated on metal oxide)
 RN 16940-66-2 HCAPLUS
 CN Borate(1-), tetrahydro-, sodium (8CI, 9CI) (CA INDEX NAME)

Valenrod 10_721479



● Na⁺

REFERENCE COUNT: 17 THERE ARE 17 CITED REFERENCES AVAILABLE FOR THIS RECORD. ALL CITATIONS AVAILABLE IN THE RE FORMAT

L28 ANSWER 2 OF 9 HCAPLUS COPYRIGHT 2006 ACS on STN
 ACCESSION NUMBER: 1994:568925 HCAPLUS
 DOCUMENT NUMBER: 121:168925
 TITLE: Homoscorpionate (Tris(pyrazolyl)borate) Ligands
 Containing Tethered 3-Phenyl Groups
 AUTHOR(S): Rheingold, Arnold L.; Ostrander, Robert L.; Haggerty,
 Brian S.; Trofimenko, Swiatoslaw
 CORPORATE SOURCE: Department of Chemistry, University of Delaware,
 Newark, DE, 19716-2522, USA
 SOURCE: Inorganic Chemistry (1994), 33(17), 3666-76
 CODEN: INOCAJ; ISSN: 0020-1669
 DOCUMENT TYPE: Journal
 LANGUAGE: English
 AB Four new homoscorpionate ligands were prepared, 3 of them involving a 3-Ph substituent subject to steric control via tethering to the the 4-position of the pyrazole ring. The tethering prevents major departure from parallel alignment of the Ph and pyrazolyl rings. The choice of the tether (methylene or 1,2-ethylene) results in the Ph 6'-CH being either pulled away from the metal in the former case or thrust toward the metal in the latter. The effect of the 5-Me substituent on the bite of the ligand was also explored. The new ligands were hydrotris(2H-benz[g]-4,5-dihydroindazol-2-yl)borate (=Tpa), hydrotris(3-methyl-2H-benz[g]-4,5-dihydroindazol-2-yl)borate (=Tpa,Me), hydrotris(1,4-dihydroindeno[1,2-c]pyrazol-1-yl)borate (=Tp_b), and hydrotris(3-phenyl-5-methylpyrazol-1-yl)borate (=TpPh,Me), each of which differed subtly in its coordinative behavior from the other 3. Complexes L2M, LMX, LRh(COD), LRh(CO)₂, LPd(η 3-methallyl), and LM_o(CO)₂(η 3-methallyl) were synthesized, and the structures of TpaRh(CO)₂, Tpa,MeZnI, TpbTl, and TpPh,MeZnI were determined by x-ray crystallog. TpaRh(CO)₂ crystallizes in the space group P.hivin.1, with a 8.574(2), b 20.113(6), c 20.188(6) Å, α 61.68(2), β 84.14(2), γ 85.15(2) $^\circ$ for Z = 4. Tpa,MeZnI crystallizes in the space group P212121, with a 10.002(1), b 16.237(3), c 19.952(3) Å for Z = 4. TpbTl crystallizes in the space group P21/n, with a 11.363(2), b = 11.096(2) Å, c 21.010(4) Å, β 99.33(2) $^\circ$ for Z = 4. The Ph and pyrazolyl planes are essentially coplanar. TpPh,MeZnI crystallizes in the space group Pna21, with a 32.645(6), b 11.327(3), c 16.180(3) Å for Z = 8. Cone and wedge angles were calculated for the new ligands and compared with revised previously reported values.

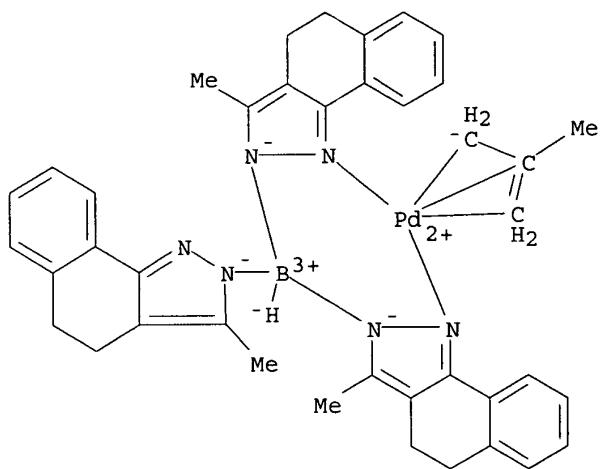
IT 157410-24-7P 157410-25-8P 157410-26-9P

157441-22-0P

RL: SPN (Synthetic preparation); PREP (Preparation)
 (preparation of)

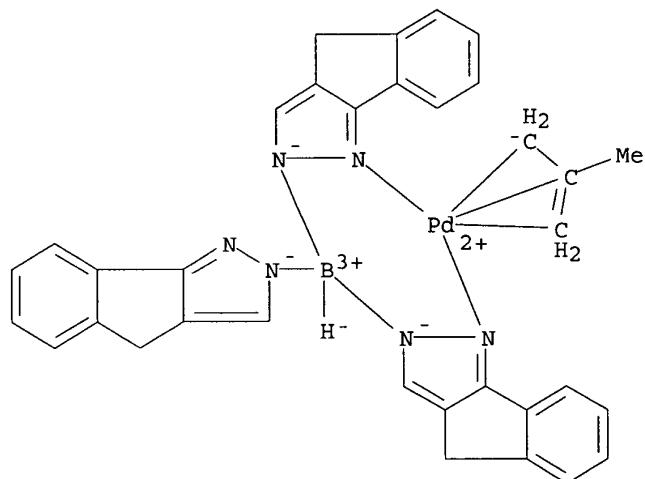
RN 157410-24-7 HCAPLUS

CN Palladium, [(1,2,3- η)-2-methyl-2-propenyl] [tris(4,5-dihydro-3-methyl-2H-benz[g]indazolato-N2)hydroborato(1-)-N1,N1']- (9CI) (CA INDEX NAME)



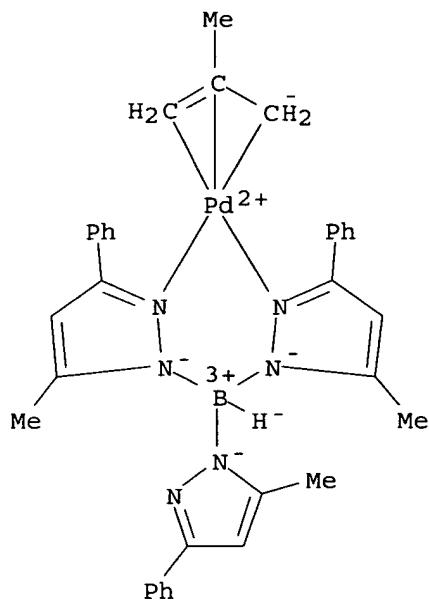
RN 157410-25-8 HCPLUS

CN Palladium, [(1,2,3- η)-2-methyl-2-propenyl][tris(2,4-dihydroinden[1,2-c]pyrazolato-N2)hydroborato(1-)-N1,N1']- (9CI) (CA INDEX NAME)

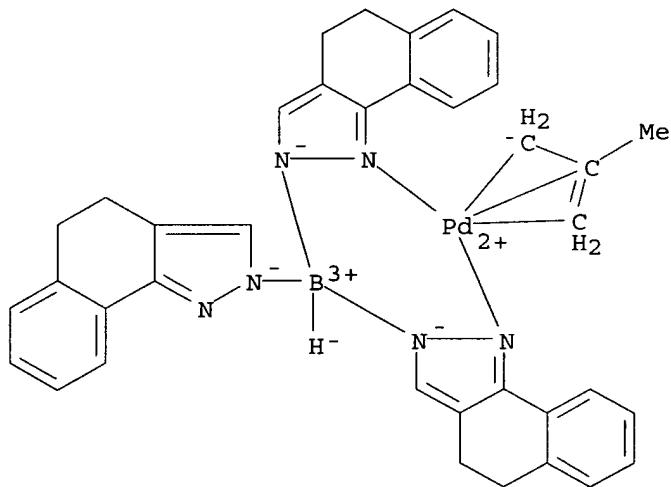


RN 157410-26-9 HCPLUS

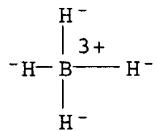
CN Palladium, [hydrotris(5-methyl-3-phenyl-1*H*-pyrazolato-N1)borato(1-)-N2,N2'][(1,2,3- η)-2-methyl-2-propenyl]- (9CI) (CA INDEX NAME)



RN 157441-22-0 HCAPLUS
 CN Palladium, [(1,2,3- η)-2-methyl-2-propenyl] [tris(4,5-dihydro-2H-benz[g]indazolato-N2)hydroborato(1-)-N1,N1',N1'']- (9CI) (CA INDEX NAME)

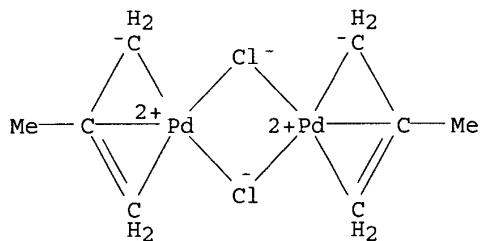


IT 13762-51-1, Potassium tetrahydroborate(1-)
 RL: RCT (Reactant); RACT (Reactant or reagent)
 (reaction of, with benzdihydroindazoles or dihydroindenopyrazole or
 phenyl(methyl)pyrazole)
 RN 13762-51-1 HCAPLUS
 CN Borate(1-), tetrahydro-, potassium (8CI, 9CI) (CA INDEX NAME)



● K^+

IT 12081-18-4, Bis(chloro(η 3-2-methylallyl) palladium)
RL: RCT (Reactant); RACT (Reactant or reagent)
(reaction of, with thallous tris(pyrazolyl)borate derivs.)
RN 12081-18-4 HCPLUS
CN Palladium, di- μ -chlorobis[(1,2,3- η)-2-methyl-2-propenyl]di- (9CI)
(CA INDEX NAME)



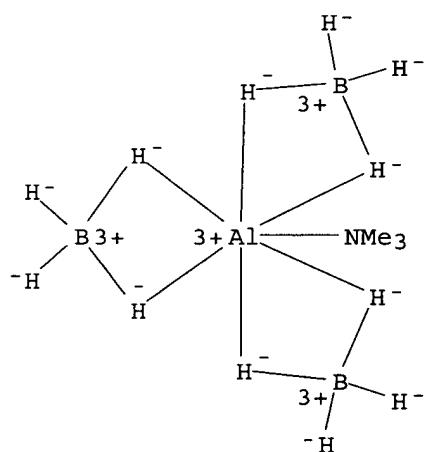
L28 ANSWER 3 OF 9 HCAPLUS COPYRIGHT 2006 ACS on STN
 ACCESSION NUMBER: 1994:437099 HCAPLUS
 DOCUMENT NUMBER: 121:37099
 TITLE: Palladium and aluminum thin film deposition
 on thermally sensitive substrates from organometallic
 complexes
 AUTHOR(S): Datta, Saswati; Kim, Yoon Gi; Dowben, P. A.; Glass,
 John A.; Kher, Shreyas S.; Peters, Scott A.; Spencer,
 James T.
 CORPORATE SOURCE: Int. Bus. Mach. Corp., Endicott, NY, 13760, USA
 SOURCE: Metallized Plastics (1992), 3, 65-72
 CODEN: MPFAEU; ISSN: 1068-7440
 DOCUMENT TYPE: Journal
 LANGUAGE: English

AB The UV decomposition of allylcyclopentadienyl palladium, $[(\eta_3\text{-C}_3\text{H}_5)(\eta_5\text{-C}_5\text{H}_5)\text{Pd}]$ (I) has been used to deposit relatively pure (>95% palladium) metal films with submicrometer spatial resolution. The films deposited on polyimide resins were found to be porous with no apparent damage to the substrate from the UV photoassisted decomposition technique. The fabrication of palladium thin films from I is consistent with the known energetics of the ligand to metal bonds in the complex. The room temperature metalization of thermally sensitive substrates, such as plastics, was also achieved from the decomposition of the volatile aluminum borohydride complex, $\text{AlH}_2(\text{BH}_4)\cdot\text{N}(\text{CH}_3)_3$. Films of thicknesses between 100 Å and 2 μm have been prepared and characterized by SEM, Auger electron spectroscopy (AES), laser microprobe mass anal. (LAMMA), x-ray diffraction (XRD), and energy dispersive x-ray anal. (EDXA).

IT 19567-28-3, Aluminum tetrahydroborate, compound with NMe_3 (1:1)
 RL: RCT (Reactant); RACT (Reactant or reagent)
 (decomposition of, chemical vacuum deposition of metal films on polyimides by)

RN 19567-28-3 HCAPLUS

CN Aluminum, (N,N-dimethylmethanamine)tris[tetrahydroborato(1-) -H, H']-, (PB-7-12-22'2'2''2'')- (9CI) (CA INDEX NAME)

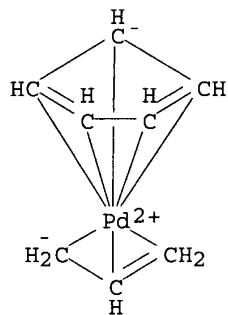


IT 1271-03-0, Allylcyclopentadienyl palladium
 RL: RCT (Reactant); RACT (Reactant or reagent)
 (decomposition of, laser radiation-induced, deposition of metal films on polyimides by)

RN 1271-03-0 HCAPLUS

Valenrod 10_721479

CN Palladium, (η 5-2,4-cyclopentadien-1-yl) (η 3-2-propenyl)- (9CI) (CA INDEX NAME)



IT 7440-05-3P, Palladium, preparation

RL: PREP (Preparation)

(films, deposition of, on polymeric substrates, by laser radiation-induced allylcyclopentadienyl palladium decomposition)

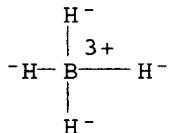
RN 7440-05-3 HCAPLUS

CN Palladium (8CI, 9CI) (CA INDEX NAME)

Pd

L28 ANSWER 4 OF 9 HCPLUS COPYRIGHT 2006 ACS on STN
 ACCESSION NUMBER: 1990:525153 HCPLUS
 DOCUMENT NUMBER: 113:125153
 TITLE: Preparation of magnetic powder containing rare earth elements
 INVENTOR(S): Buchkov, D.; Dragieva, I.; Iliev, I.; Todorov, O.;
 Slavcheva-Staikova, M.
 PATENT ASSIGNEE(S): Higher Institute of Electrical and Mechanical
 Engineering, Sofia, Bulg.
 SOURCE: Ger. (East), 3 pp.
 CODEN: GEXXA8
 DOCUMENT TYPE: Patent
 LANGUAGE: German
 FAMILY ACC. NUM. COUNT: 1
 PATENT INFORMATION:

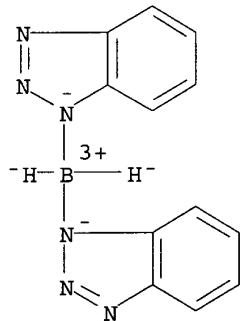
PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
DD 271612	A7	19890913	DD 1986-291413	19860618 <--
PRIORITY APPLN. INFO.:			DD 1986-291413	19860618
AB	A solution of water-soluble salts of a rare earth metal, Fe, Ni, Co, Cu, and Pd at concns. of 1 + 10-3 to 1 + 10-1 mol/L is mixed with aqueous Na borohydride or Na hypophosphite (reducing agent) at 0.05-1 mol/L and 10-90° for 1-30 min in a magnetic field of (1-8) + 105 A/m to give a magnetic powder.			
IT	16940-66-2, Sodium borohydride RL: RCT (Reactant); RACT (Reactant or reagent) (reduction by, of metal salt solns., in preparation of magnetic powders)			
RN	16940-66-2 HCPLUS			
CN	Borate(1-), tetrahydro-, sodium (8CI, 9CI) (CA INDEX NAME)			



● Na⁺

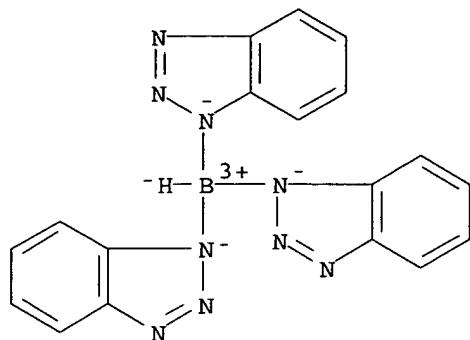
Valenrod 10_721479

L28 ANSWER 5 OF 9 HCPLUS COPYRIGHT 2006 ACS on STN
ACCESSION NUMBER: 1989:534235 HCPLUS
DOCUMENT NUMBER: 111:134235
TITLE: Unexpected regiochemistry in the reaction of benzotriazole with KBH4: synthesis of a new class of poly(azolyl)borate ligands
AUTHOR(S): Lalor, F. J.; Miller, S.; Garvey, N.
CORPORATE SOURCE: Dep. Chem., Univ. Coll., Cork, Ire.
SOURCE: Journal of Organometallic Chemistry (1988), 356(2), C57-60
CODEN: JORCAI; ISSN: 0022-328X
DOCUMENT TYPE: Journal
LANGUAGE: English
OTHER SOURCE(S): CASREACT 111:134235
AB The reaction of KBH4 with molten benzotriazole forms a series of poly(benzotriazolyl)borate anions $[HnB(C_6H_4N_3)_4-n]^-$ (n = 0-3) depending on the temperature. The regiospecificity of the process differs from that of pyrazole/BH4- reactions in that B-N bond formation takes place in a manner that maximizes steric crowding at boron (i.e. at the triazole N(1) atoms). The complex-forming abilities of the new ligands have been investigated, and show some interesting differences from those of the poly(1-pyrazolyl)borate analogs.
IT 109782-45-8P 122627-92-3P 122658-64-4P
RL: RCT (Reactant); SPN (Synthetic preparation); PREP (Preparation); RACT (Reactant or reagent)
(preparation and reaction of, with transition metal complexes)
RN 109782-45-8 HCPLUS
CN Borate(1-), bis(1H-benzotriazolato- κN_1)dihydro-, potassium, (T-4)- (9CI) (CA INDEX NAME)



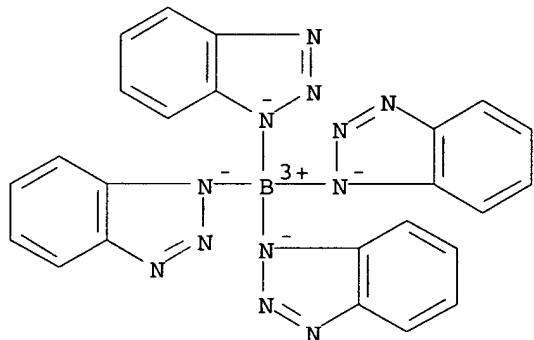
● K^+

RN 122627-92-3 HCPLUS
CN Borate(1-), tris(1H-benzotriazolato- κN_1)hydro-, potassium, (T-4)- (9CI) (CA INDEX NAME)



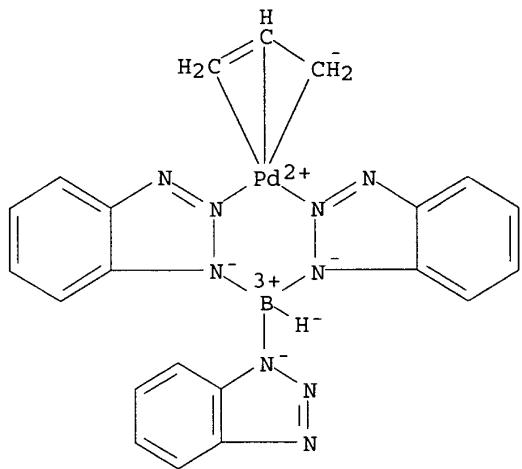
● K⁺

RN 122658-64-4 HCAPLUS
CN Borate(1-), tetrakis(1H-benzotriazolato-N1)-, potassium (9CI) (CA INDEX NAME)

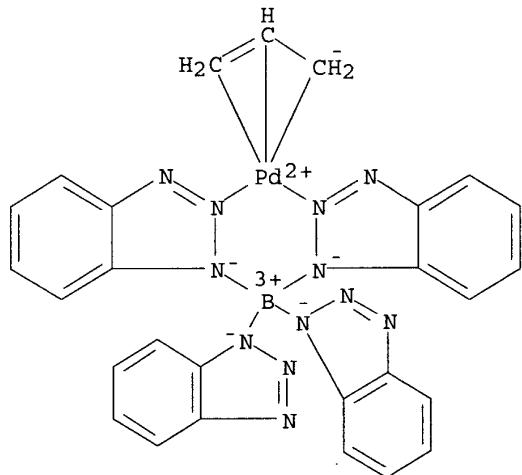


● K⁺

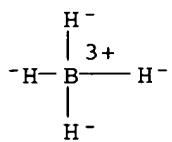
IT 122651-40-5P 122651-42-7P
RL: SPN (Synthetic preparation); PREP (Preparation)
(preparation of)
RN 122651-40-5 HCAPLUS
CN Palladium, (η₃-2-propenyl) [tris(1H-benzotriazolato-N1)hydroborato(1-) - N₂,N₂'] - (9CI) (CA INDEX NAME)



RN 122651-42-7 HCAPLUS
CN Palladium, (η₃-2-propenyl) [tetrakis(1H-benzotriazolato-N1)borato(1-) - N₂,N₂']- (9CI) (CA INDEX NAME)



IT 13762-51-1, Potassium tetrahydroborate
RL: RCT (Reactant); RACT (Reactant or reagent)
(reaction of, with molten benzotriazole, regiochem. of)
RN 13762-51-1 HCAPLUS
CN Borate(1-), tetrahydro-, potassium (8CI, 9CI) (CA INDEX NAME)



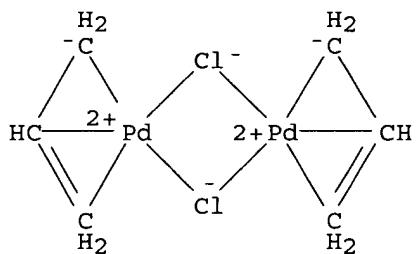
● K^+

IT 12012-95-2

RL: RCT (Reactant); RACT (Reactant or reagent)
(reaction of, with poly(benzotriazolyl)borate anions)

RN 12012-95-2 HCPLUS

CN Palladium, di- μ -chlorobis(η 3-2-propenyl)di- (9CI) (CA INDEX NAME)



Valenrod 10_721479

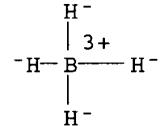
L28 ANSWER 6 OF 9 HCAPLUS COPYRIGHT 2006 ACS on STN
ACCESSION NUMBER: 1984:124631 HCAPLUS
DOCUMENT NUMBER: 100:124631
TITLE: Concentration of **platinum** metals by reducing
with sodium tetrahydroborate in the presence of
manganese dioxide
AUTHOR(S): Khain, V. S.; Volkov, A. A.; Martynova, V. F.
CORPORATE SOURCE: Ukhtinsk. Ind. Inst., Ukhta, USSR
SOURCE: Zhurnal Prikladnoi Khimii (Sankt-Peterburg, Russian
Federation) (1984), 57(1), 30-3
CODEN: ZPKHAB; ISSN: 0044-4618
DOCUMENT TYPE: Journal
LANGUAGE: Russian
AB Pt, Pd, and Rh were recovered from H₂PtCl₆, PdCl₂, and
RhCl₃ solns. by reduction with NaBH₄ at pH 4-10 in the presence of KMnO₄.
MnO₂ formed by the reaction of KMnO₄ with NaBH₄ served as a collector for
Pt-group metals.
IT 7440-06-4P, preparation
RL: PUR (Purification or recovery); PREP (Preparation)
(recovery of, from hydrogen hexachloroplatinate solns., by reduction with
sodium borohydride, manganese dioxide collector in)
RN 7440-06-4 HCAPLUS
CN Platinum (8CI, 9CI) (CA INDEX NAME)

Pt

IT 7440-05-3P, preparation
RL: PUR (Purification or recovery); PREP (Preparation)
(recovery of, from palladium chloride solns., by reduction with
sodium borohydride, manganese dioxide collector in)
RN 7440-05-3 HCAPLUS
CN Palladium (8CI, 9CI) (CA INDEX NAME)

Pd

IT 16940-66-2
RL: RCT (Reactant); RACT (Reactant or reagent)
(reduction by, of platinum-group metal chlorides, in
aqueous solns.)
RN 16940-66-2 HCAPLUS
CN Borate(1-), tetrahydro-, sodium (8CI, 9CI) (CA INDEX NAME)



● Na⁺

IT 7647-10-1

Valenrod 10_721479

RL: RCT (Reactant); RACT (Reactant or reagent)
(reduction of, with sodium borohydride in aqueous media, palladium
recovery by)

RN 7647-10-1 HCAPLUS

CN Palladium chloride (PdCl₂) (6CI, 8CI, 9CI) (CA INDEX NAME)

Cl-Pd-Cl

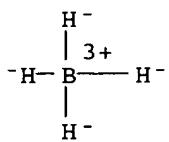
L28 ANSWER 7 OF 9 HCPLUS COPYRIGHT 2006 ACS on STN
ACCESSION NUMBER: 1979:179411 HCPLUS
DOCUMENT NUMBER: 90:179411
TITLE: Reaction of sodium tetrahydroborate with precious metal salts
AUTHOR(S): Mal'tseva, N. N.; Sterlyadkina, Z. K.; Erusalimchik, I. G.; Mikheeva, V. I.
CORPORATE SOURCE: USSR
SOURCE: Zhurnal Neorganicheskoi Khimii (1979), 24(3), 822-4
CODEN: ZNOKAQ; ISSN: 0044-457X
DOCUMENT TYPE: Journal
LANGUAGE: Russian
AB During the reaction of NaBH4 with aqueous solns. of Ag, Au, Pd, and Pt salts, highly pure free metal precipitated. The stoichiometry of the oxidation-reduction reaction was established on the basis of the H evolved, solution pH, and amount of reducing agent added. In acidic solution hydrolysis of NaBH4 occurred simultaneously with the oxidation-reduction process. The precipitated metal is highly dispersed and absorbs a significant amount of H.
IT 7440-05-3P, preparation
RL: SPN (Synthetic preparation); PREP (Preparation)
(preparation of, by reduction of palladium salts with sodium tetrahydroborate)
RN 7440-05-3 HCPLUS
CN Palladium (8CI, 9CI) (CA INDEX NAME)

Pd

IT 7440-06-4P, preparation
RL: SPN (Synthetic preparation); PREP (Preparation)
(preparation of, by reduction of platinum salts with sodium tetrahydroborate)
RN 7440-06-4 HCPLUS
CN Platinum (8CI, 9CI) (CA INDEX NAME)

Pt

IT 16940-66-2
RL: RCT (Reactant); RACT (Reactant or reagent)
(reduction by, of precious metal salts)
RN 16940-66-2 HCPLUS
CN Borate(1-), tetrahydro-, sodium (8CI, 9CI) (CA INDEX NAME)



● Na^+

IT 7440-05-3D, salts 7440-06-4D, salts 7440-22-4D
, salts 7440-57-5D, salts
RL: RCT (Reactant); RACT (Reactant or reagent)
(reduction of, by sodium tetrahydroborate)
RN 7440-05-3 HCAPLUS
CN Palladium (8CI, 9CI) (CA INDEX NAME)

Pd

RN 7440-06-4 HCAPLUS
CN Platinum (8CI, 9CI) (CA INDEX NAME)

Pt

RN 7440-22-4 HCAPLUS
CN Silver (8CI, 9CI) (CA INDEX NAME)

Ag

RN 7440-57-5 HCAPLUS
CN Gold (8CI, 9CI) (CA INDEX NAME)

Au

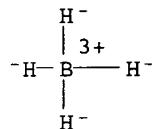
L28 ANSWER 8 OF 9 HCPLUS COPYRIGHT 2006 ACS on STN
 ACCESSION NUMBER: 1972:493257 HCPLUS
 DOCUMENT NUMBER: 77:93257
 TITLE: Catalysts formed by the reduction of Group VIII metal salts with sodium borohydride
 AUTHOR(S): Sterlyadkina, Z. K.; Mal'tseva, N. N.; Frangulyan, G. D.; Polkovnikov, B. D.; Bakulina, V. M.
 CORPORATE SOURCE: Inst. Org. Khim. im. Zelinskogo, Moscow, USSR
 SOURCE: Izvestiya Akademii Nauk SSSR, Seriya Khimicheskaya (1972), (6), 1240-5
 CODEN: IASKA6; ISSN: 0002-3353
 DOCUMENT TYPE: Journal
 LANGUAGE: Russian
 AB The ppt. formed by reduction of NiCl_2 and CoCl_2 by NaBH_4 contain the free metal and M_2B ($\text{M} = \text{Co, Ni}$) after being heated in Ar at 400° . Similar reduction of PdCl_2 and H_2PtCl_6 leads to Pt and α -Pd.
 IT 7440-06-4P, uses and miscellaneous
 RL: CAT (Catalyst use); PREP (Preparation); USES (Uses)
 (catalysts, reduction of chloroplatinic acid by sodium borohydride in preparation of)
 RN 7440-06-4 HCPLUS
 CN Platinum (8CI, 9CI) (CA INDEX NAME)

Pt

IT 7440-05-3P, uses and miscellaneous
 RL: CAT (Catalyst use); PREP (Preparation); USES (Uses)
 (catalysts, reduction of palladium chloride by sodium borohydride in preparation of)
 RN 7440-05-3 HCPLUS
 CN Palladium (8CI, 9CI) (CA INDEX NAME)

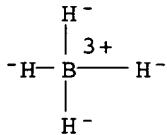
Pd

IT 16940-66-2
 RL: RCT (Reactant); RACT (Reactant or reagent)
 (reduction by, of Group VIII metal salts in catalyst preparation)
 RN 16940-66-2 HCPLUS
 CN Borate(1-), tetrahydro-, sodium (8CI, 9CI) (CA INDEX NAME)

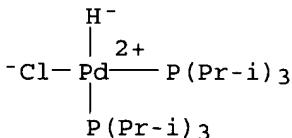


● Na^+

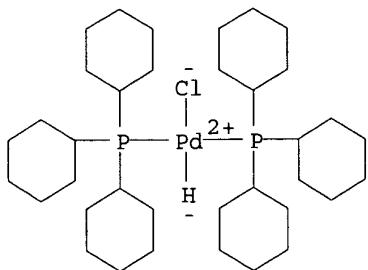
L28 ANSWER 9 OF 9 HCAPLUS COPYRIGHT 2006 ACS on STN
 ACCESSION NUMBER: 1971:71098 HCAPLUS
 DOCUMENT NUMBER: 74:71098
 TITLE: Tertiary phosphine-hydride and -hydridoborohydride
 compounds of nickel and palladium
 AUTHOR(S): Green, Malcolm L. H.; Munakata, H.; Saito, Tato
 CORPORATE SOURCE: Inorg. Chem. Lab., Oxford, UK
 SOURCE: Journal of the Chemical Society [Section] A:
 Inorganic, Physical, Theoretical (1971),
 (3), 469-74
 CODEN: JCSIAP; ISSN: 0022-4944
 DOCUMENT TYPE: Journal
 LANGUAGE: English
 AB The isolation and properties of the complexes $\text{trans-}(\text{R13P})_2\text{M1H}(\text{BH}_4)$ and
 $\text{trans-}(\text{R13P})_2\text{M2HX}$, where $\text{M1} = \text{Ni}$ or Pd ; $\text{R1} = \text{iso-Pr}$, cyclohexyl;
 $\text{M2} = \text{Pd}$; $\text{X} = \text{Cl}$, Br , I , or NCS , are described. Phosphine
 exchange reactions of $(\text{R13P})_2\text{M1H}$ with R23P , where $\text{R2} = \text{Et}$, Pr , and Bu ,
 were studied by ^1H NMR.
 IT 16971-29-2DP, Borate(1-), tetrahydro-, Group VIII
 metal complexes 27900-91-0P 28016-71-9P
 28016-73-1P 30916-07-5P 31178-50-4P
 31178-51-5P
 RL: SPN (Synthetic preparation); PREP (Preparation)
 (preparation of)
 RN 16971-29-2 HCAPLUS
 CN Borate(1-), tetrahydro- (8CI, 9CI) (CA INDEX NAME)



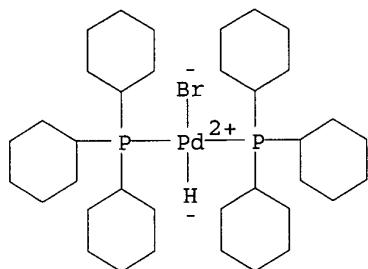
RN 27900-91-0 HCAPLUS
 CN Palladium, chlorohydrobis[tris(1-methylethyl)phosphine]-, (SP-4-3)- (9CI)
 (CA INDEX NAME)



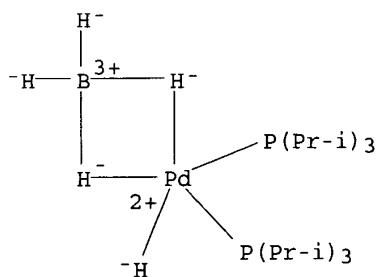
RN 28016-71-9 HCAPLUS
 CN Palladium, chlorohydrobis(tricyclohexylphosphine)-, (SP-4-3)- (9CI) (CA
 INDEX NAME)



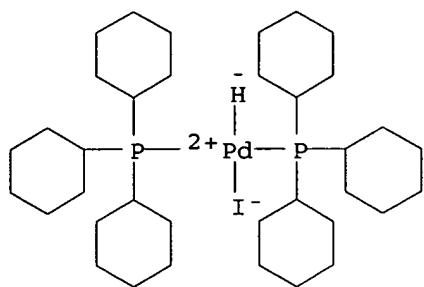
RN 28016-73-1 HCPLUS
CN Palladium, bromohydrobis(tricyclohexylphosphine)-, trans- (8CI) (CA INDEX
NAME)



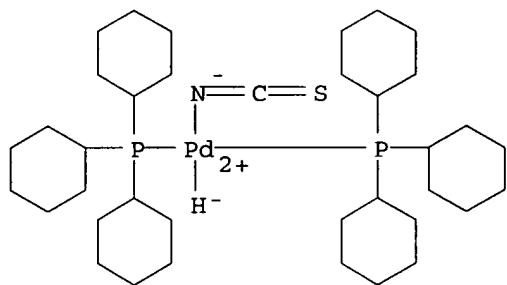
RN 30916-07-5 HCPLUS
CN Palladium, hydro[tetrahydroborate(1-)]bis(triisopropylphosphine)-, trans-
(8CI) (CA INDEX NAME)



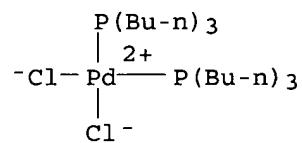
RN 31178-50-4 HCPLUS
CN Palladium, hydroiodobis(tricyclohexylphosphine)-, trans- (8CI) (CA INDEX
NAME)



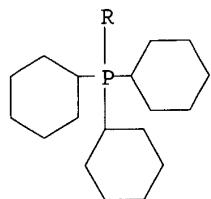
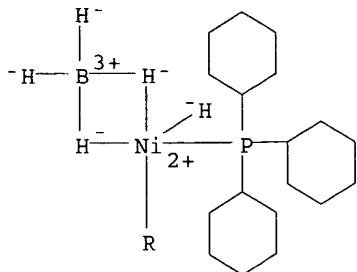
RN 31178-51-5 HCAPLUS
 CN Palladium, hydro(isothiocyanato)bis(tricyclohexylphosphine)-, trans- (8CI)
 (CA INDEX NAME)



IT 17523-47-6 24899-12-5
 RL: RCT (Reactant); RACT (Reactant or reagent)
 (redistribution reaction of, with Group VIII metal
 trialkylphosphine complexes, NMR in relation to)
 RN 17523-47-6 HCAPLUS
 CN Palladium, dichlorobis(tributylphosphine)-, (SP-4-1)- (9CI) (CA INDEX
 NAME)



RN 24899-12-5 HCAPLUS
 CN Nickel, hydro[tetrahydroborato(1-)-H,H']bis(tricyclohexylphosphine)-,
 (TB-5-11)- (9CI) (CA INDEX NAME)

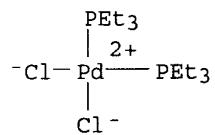


IT 15642-19-0

RL: RCT (Reactant); RACT (Reactant or reagent)
 (redistribution reaction of, with Group VIII metal
 tricyclohexylphosphine complexes, NMR in relation to)

RN 15642-19-0 HCPLUS

CN Palladium, dichlorobis(triethylphosphine)-, (SP-4-1)- (9CI) (CA INDEX NAME)

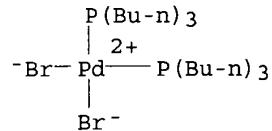


IT 17523-48-7

RL: RCT (Reactant); RACT (Reactant or reagent)
 (redistribution reaction of, with nickel tricyclohexylphosphine
 complexes, NMR in relation to)

RN 17523-48-7 HCPLUS

CN Palladium, dibromobis(tributylphosphine)-, trans- (8CI) (CA INDEX NAME)

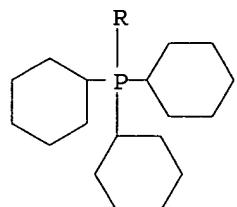
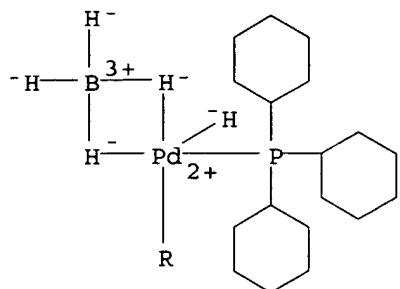


IT 30916-06-4

RL: RCT (Reactant); RACT (Reactant or reagent)
 (redistribution reaction of, with palladium trialkyl
 phosphine complexes, NMR in relation to)

RN 30916-06-4 HCPLUS

CN Palladium, hydro[tetrahydroborato(1-) -H,H']bis(tricyclohexylphosphine) - ,
(TB-5-11) - (9CI) (CA INDEX NAME)

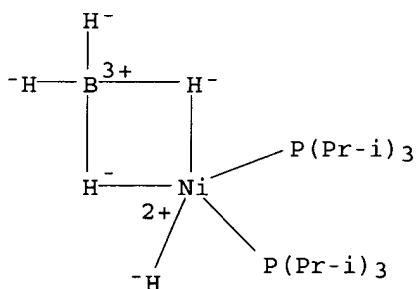


IT 31095-23-5

RL: RCT (Reactant); RACT (Reactant or reagent)
(redistribution reaction of, with palladium tributyl
phosphine complexes, NMR in relation to)

RN 31095-23-5 HCAPLUS

CN Nickel, hydro[tetrahydroborato(1-) -H,H']bis[tris(1-methylethyl)phosphine] - ,
(TB-5-11) - (9CI) (CA INDEX NAME)



Valenrod 10_721479

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L3      104214 SEA FILE=REGISTRY ABB=ON  PLU=ON  METAL OR METALS OR MAGNESIUM
L4      128483 SEA FILE=REGISTRY ABB=ON  PLU=ON  BORATE
L5      15523 SEA FILE=HCAPLUS ABB=ON  PLU=ON  L1 OR TETRAHYDROBORATE OR
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L6      SEL  PLU=ON  L2 1- CHERM :      52 TERMS
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L8      15583 SEA FILE=HCAPLUS ABB=ON  PLU=ON  L7 OR (SODIUM OR LITHIUM OR
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L9      4137887 SEA FILE=HCAPLUS ABB=ON  PLU=ON  L3 OR METAL OR MAGNESIUM
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L17     1952 SEA FILE=HCAPLUS ABB=ON  PLU=ON  L16 AND L15
L18     104576 SEA FILE=REGISTRY ABB=ON  PLU=ON  PALLADIUM/BI
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L21     513099 SEA FILE=HCAPLUS ABB=ON  PLU=ON  L18 OR L19 OR PALLADIUM OR
              PLATINUM OR PD OR PT
L25     3041 SEA FILE=HCAPLUS ABB=ON  PLU=ON  (L5 OR L8) (L) (PREP?/RL OR
              PREP OR PREPARTION)
L26     78 SEA FILE=HCAPLUS ABB=ON  PLU=ON  L17 AND L25
L27     11 SEA FILE=HCAPLUS ABB=ON  PLU=ON  L26 AND L21
L28     9 SEA FILE=HCAPLUS ABB=ON  PLU=ON  L27 AND PD=<NOVEMBER 25, 2003
L29     2 SEA FILE=HCAPLUS ABB=ON  PLU=ON  L27 NOT L28
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L29 ANSWER 1 OF 2 HCAPLUS COPYRIGHT 2006 ACS on STN
 ACCESSION NUMBER: 2004:676920 HCAPLUS
 DOCUMENT NUMBER: 142:6582
 TITLE: Hydrocarbon C-H bond activation with $\text{Tp}'\text{Pt}$ complexes
 AUTHOR(S): Norris, Cynthia M.; Templeton, Joseph L.
 CORPORATE SOURCE: Department of Chemistry, University of North Carolina,
 Chapel Hill, NC, 27599-3290, USA
 SOURCE: ACS Symposium Series (2004), 885 (Activation and
 Functionalization of C-H Bonds), 303-318
 CODEN: ACSMC8; ISSN: 0097-6156
 PUBLISHER: American Chemical Society
 DOCUMENT TYPE: Journal; General Review
 LANGUAGE: English

AB A review. Stabilization of Pt complexes resembling key intermediates in C-H bond activation has been achieved through the use of the strongly electron-donating hydridotris(3,5-dimethyl-1H-pyrazol-1-yl)borate (Tp') ligand. Isolation of stable alkyl(hydrido)platinum(IV) complexes, $\text{Tp}'\text{PtMe}_2\text{H}$ and $\text{Tp}'\text{PtMeH}_2$, allowed the authors to study the mechanism of reductive alkane elimination from these complexes via (1) thermolysis, (2) Lewis acid addition, and (3) low temperature protonation. By replacing the alkyl group with SiR_3 and C_6R_5 in the low temperature protonation of $\text{Tp}'\text{PtRH}_2$, isolation of five-coordinate Pt (IV) complexes and Pt (II) η^2 -arene adducts, resp., was achieved. These Pt (II)/ Pt (IV) interconversions provide insight into the mechanisms by which these reagents activate strong C-H bonds and also provide a foundation for future plans to functionalize hydrocarbons.

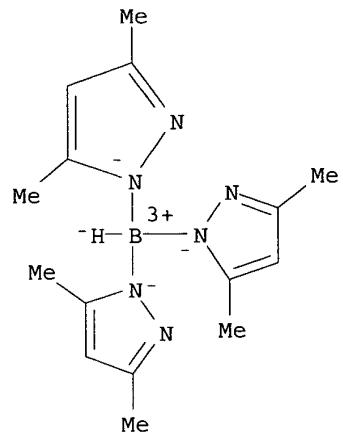
IT 7440-06-4DP, Platinum, hydridotris(3,5-dimethylpyrazol-1-yl)borate complexes 83534-02-5DP, Hydridotris(3,5-dimethylpyrazol-1-yl)borate, platinum complexes
 RL: CPS (Chemical process); PEP (Physical, engineering or chemical process); RCT (Reactant); SPN (Synthetic preparation); PREP (Preparation); PROC (Process); RACT (Reactant or reagent)
 (hydrocarbon carbon-hydrogen bond activation with platinum hydridotripyrazolylborato complexes)

RN 7440-06-4 HCAPLUS
 CN Platinum (8CI, 9CI) (CA INDEX NAME)

Pt

RN 83534-02-5 HCAPLUS
 CN Borate(1-), tris(3,5-dimethyl-1H-pyrazolato- $\kappa\text{N}1$)hydro-, (T-4)- (9CI)
 (CA INDEX NAME)

Valenrod 10_721479



REFERENCE COUNT:

91

THERE ARE 91 CITED REFERENCES AVAILABLE FOR THIS RECORD. ALL CITATIONS AVAILABLE IN THE RE FORMAT

L29 ANSWER 2 OF 2 HCAPLUS COPYRIGHT 2006 ACS on STN
 ACCESSION NUMBER: 2004:651206 HCAPLUS
 DOCUMENT NUMBER: 141:192634
 TITLE: method to produce tetrahydroborate salts
 INVENTOR(S): Suda, Seijiro; Li, Chou-Peng; Iwase, Yasuyoshi;
 Morigasaki, Nobuto
 PATENT ASSIGNEE(S): Materials & Energy Research Institute Tokyo, Ltd.
 MERIT., Japan
 SOURCE: Jpn. Kokai Tokkyo Koho, 9 pp.
 CODEN: JKXXAF
 DOCUMENT TYPE: Patent
 LANGUAGE: Japanese
 FAMILY ACC. NUM. COUNT: 1
 PATENT INFORMATION:

PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
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JP 2004224593	A2	20040812	JP 2003-11299	20030120
PRIORITY APPLN. INFO.:			JP 2003-11299	20030120

AB The method includes hydriding reaction of alkali or alkaline earth metal meta-borate by supplying mixed gas of O₂-accepting gas (e.g. CO, C_{≤5} hydrocarbons), H₂, and O₂ with a hydriding catalyst selected from ≥1 of Ni, Co, Fe, Pt, Cu, Pd, Ru, and Rh. The borate raw material is obtained from spent fuel cells. The method enhances the hydriding reaction under mild conditions, and produces NaBH₄, LiBH₄, and KBH₄ in low operation costs.

IT 14100-65-3, Meta-borate
 RL: RCT (Reactant); RACT (Reactant or reagent)
 (alkali metal salt, alkaline metal salt; method to produce tetrahydroborate salts)
 RN 14100-65-3 HCAPLUS
 CN Borate (B021-) (8CI, 9CI) (CA INDEX NAME)

-O—B=O

IT 7440-05-3, Palladium, uses 7440-06-4, Platinum, uses
 RL: CAT (Catalyst use); USES (Uses)
 (method to produce tetrahydroborate salts)
 RN 7440-05-3 HCAPLUS
 CN Palladium (8CI, 9CI) (CA INDEX NAME)

Pd

RN 7440-06-4 HCAPLUS
 CN Platinum (8CI, 9CI) (CA INDEX NAME)

Pt

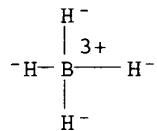
IT 13762-51-1P, Potassium Tetrahydroborate
 16940-66-2P, Sodium Tetrahydroborate
 16949-15-8P, Lithium Tetrahydroborate
 RL: PUR (Purification or recovery); PREP (Preparation)

Valenrod 10_721479

(method to produce **tetrahydroborate salts**)

RN 13762-51-1 HCPLUS

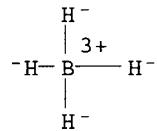
CN Borate(1-), tetrahydro-, potassium (8CI, 9CI) (CA INDEX NAME)



● K^+

RN 16940-66-2 HCPLUS

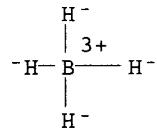
CN Borate(1-), tetrahydro-, sodium (8CI, 9CI) (CA INDEX NAME)



● Na^+

RN 16949-15-8 HCPLUS

CN Borate(1-), tetrahydro-, lithium (8CI, 9CI) (CA INDEX NAME)



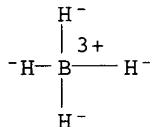
● Li^+

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L3      104214 SEA FILE=REGISTRY ABB=ON  PLU=ON  METAL OR METALS OR MAGNESIUM
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L5      15523 SEA FILE=HCAPLUS ABB=ON  PLU=ON  L1 OR TETRAHYDROBORATE OR
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L7      15422 SEA FILE=HCAPLUS ABB=ON  PLU=ON  L6
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L12     15254 SEA FILE=HCAPLUS ABB=ON  PLU=ON  L10(L) L9
L14     38473 SEA FILE=HCAPLUS ABB=ON  PLU=ON  L10(L) (RCT/RL OR RACT/RL)
L15     250013 SEA FILE=HCAPLUS ABB=ON  PLU=ON  L9(L) (RCT/RL OR RACT/RL)
L16     2579 SEA FILE=HCAPLUS ABB=ON  PLU=ON  L12 AND L14
L17     1952 SEA FILE=HCAPLUS ABB=ON  PLU=ON  L16 AND L15
L18     104576 SEA FILE=REGISTRY ABB=ON  PLU=ON  PALLADIUM/BI
L19     208 SEA FILE=REGISTRY ABB=ON  PLU=ON  PT/MF
L21     513099 SEA FILE=HCAPLUS ABB=ON  PLU=ON  L18 OR L19 OR PALLADIUM OR
              PLATINUM OR PD OR PT
L25     3041 SEA FILE=HCAPLUS ABB=ON  PLU=ON  (L5 OR L8) (L) (PREP?/RL OR
              PREP OR PREPARTION)
L26     78 SEA FILE=HCAPLUS ABB=ON  PLU=ON  L17 AND L25
L27     11 SEA FILE=HCAPLUS ABB=ON  PLU=ON  L26 AND L21
L28     9 SEA FILE=HCAPLUS ABB=ON  PLU=ON  L27 AND PD=<NOVEMBER 25, 2003
L30     524 SEA FILE=HCAPLUS ABB=ON  PLU=ON  (L5 OR L8) (L) (PURIFI?/RL OR
              PURIFI? OR RECOVER?)
L31     152 SEA FILE=HCAPLUS ABB=ON  PLU=ON  L30 AND L10 AND L9
L32     38 SEA FILE=HCAPLUS ABB=ON  PLU=ON  L31 AND L9 AND L21
L33     36 SEA FILE=HCAPLUS ABB=ON  PLU=ON  L32 NOT (L27 OR L28)
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L33 ANSWER 1 OF 36 HCPLUS COPYRIGHT 2006 ACS on STN
ACCESSION NUMBER: 2005:652766 HCPLUS
DOCUMENT NUMBER: 144:153746
TITLE: Recovery of platinum from spent catalysts
from catalytic reforming unit using the Gemini process
AUTHOR(S): Soltan-Mohammadzadeh, Jafar-Sadeg; Fatehifar, Esmaeel;
Hamedi, Hassan
CORPORATE SOURCE: Chem. Engineering, Sahand Technical University,
Tabriz, Iran
SOURCE: Iranian National Chemical Engineering Congress, 8th,
Mashhad, Islamic Republic of Iran, Oct. 19-21, 2003
(2003), 227/1-227/11. Danishgah-i Ferdowsi Mashhad,
Danishkada Muhandisi: Mashhad, Iran.
CODEN: 69GMYP
DOCUMENT TYPE: Conference; (computer optical disk)
LANGUAGE: Persian
AB Studies were carried out to optimize the conditions for the solubilization
of base catalyst alumina and concentration and separation of Pt using formic
acid and NaBH4. Pt recovery amounted to 94.1% with formic acid
and 95.68% with NaBH4.
IT 16940-66-2, Sodium borohydride
RL: NUU (Other use, unclassified); USES (Uses)
(recovery of platinum from spent catalysts from
catalytic reforming unit using Gemini process)
RN 16940-66-2 HCPLUS
CN Borate(1-), tetrahydro-, sodium (8CI, 9CI) (CA INDEX NAME)



● Na+

IT 7440-06-4P, Platinum, preparation
RL: PUR (Purification or recovery); PREP (Preparation)
(recovery of platinum from spent catalysts from catalytic
reforming unit using Gemini process)
RN 7440-06-4 HCPLUS
CN Platinum (8CI, 9CI) (CA INDEX NAME)

Pt

L33 ANSWER 2 OF 36 HCAPLUS COPYRIGHT 2006 ACS on STN
 ACCESSION NUMBER: 2005:395865 HCAPLUS
 DOCUMENT NUMBER: 142:396971
 TITLE: Process for removal of heavy **metal** ions from wastewater using reducing agent composition
 INVENTOR(S): Covaliov, Victor; Covaliova, Olga; Duca, Gheorghe
 PATENT ASSIGNEE(S): Universitatea de Stat din Moldova, Moldova
 SOURCE: Mold. Unexam. Pat. Appl., 7 pp.
 CODEN: MDXXAX
 DOCUMENT TYPE: Patent
 LANGUAGE: Romanian
 FAMILY ACC. NUM. COUNT: 1
 PATENT INFORMATION:

PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
MD 990123	A	20001130	MD 1999-123	19990317
MD 1649	F2	20010430		

PRIORITY APPLN. INFO.: MD 1999-123 19990317
 AB The invention refers to processes for **purification** of sewage from. Heavy **metal** ions are removed from wastewater using a chemical reducing agent composition containing **sodium borohydride** 0.05-0.06, sodium hypophosphite 0.25-0.30, formaldehyde 0.60-0.70, and hydrazine 0.30-0.40 mol/L at pH 6-8. The process is realized in a flow at the linear velocity of 0.2-0.3 m/min in an electromagnetic field at a voltage of 20-30 W/dm³ and frequency of 60-74 kHz on the catalyst surface. A nickel-zinc alloy is used as catalyst and is coated by electrochem. deposition onto a steel grill with the cell dimension of 0.5-2.0 mm. The solution is added so that the weight ratio to the total heavy **metal** ions content is (1.5-2.0):1. Use of this composition results in increased sewage **purification** degree and decreased material and power consumption for the **purification** process.

IT 7440-02-0, Nickel, processes 7440-05-3,
 Palladium, processes 7440-22-4, Silver, processes
 7440-50-8, Copper, processes 7440-66-6, Zinc, processes
 RL: REM (Removal or disposal); PROC (Process)
 (removal of heavy **metal** ions from wastewater using reducing agent composition)
 RN 7440-02-0 HCAPLUS
 CN Nickel (8CI, 9CI) (CA INDEX NAME)

Ni

RN 7440-05-3 HCAPLUS
 CN Palladium (8CI, 9CI) (CA INDEX NAME)

Pd

RN 7440-22-4 HCAPLUS
 CN Silver (8CI, 9CI) (CA INDEX NAME)

Ag

RN 7440-50-8 HCAPLUS

Valenrod 10_721479

CN Copper (7CI, 8CI, 9CI) (CA INDEX NAME)

Cu

RN 7440-66-6 HCPLUS

CN Zinc (7CI, 8CI, 9CI) (CA INDEX NAME)

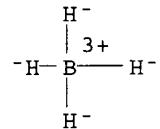
Zn

IT 16940-66-2, Sodium borohydride

RL: NUU (Other use, unclassified); USES (Uses)
(removal of heavy metal ions from wastewater using reducing
agent composition containing)

RN 16940-66-2 HCPLUS

CN Borate(1-), tetrahydro-, sodium (8CI, 9CI) (CA INDEX NAME)



● Na^+

L33 ANSWER 3 OF 36 HCAPLUS COPYRIGHT 2006 ACS on STN
 ACCESSION NUMBER: 2004:780106 HCAPLUS
 DOCUMENT NUMBER: 141:245652
 TITLE: Production and purification of hydrogen
 INVENTOR(S): Ring, Terry A.; Freise, William; Maylett, Brett;
 Fisher, Matthew
 PATENT ASSIGNEE(S): USA
 SOURCE: U.S. Pat. Appl. Publ., 8 pp.
 CODEN: USXXCO
 DOCUMENT TYPE: Patent
 LANGUAGE: English
 FAMILY ACC. NUM. COUNT: 1
 PATENT INFORMATION:

PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
US 2004184987	A1	20040923	US 2003-384961	20030310
PRIORITY APPLN. INFO.:			US 2003-384961	20030310

AB Highly pure hydrogen gas is produced by reaction of Na, NaH, NaAlH₄, NaBH₄, Li, LiH, or LiBH₄ with water and subsequent purification of the hydrogen using a hydrogen-selective metal membrane. The membrane can consist of Pd, Ta, V, Nb, Y, Th, Zr, or Ti. The membrane is supported by a rigid porous structural support.
 IT 7440-05-3, Palladium, uses 7440-32-6,
 Titanium, uses
 RL: DEV (Device component use); USES (Uses)
 (membrane; production and purification of hydrogen)
 RN 7440-05-3 HCAPLUS
 CN Palladium (8CI, 9CI) (CA INDEX NAME)

Pd

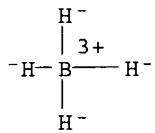
RN 7440-32-6 HCAPLUS
 CN Titanium (8CI, 9CI) (CA INDEX NAME)

Ti

IT 7440-23-5, Sodium, reactions 16940-66-2, Sodium
 borohydride (NaBH₄) 16949-15-8, Lithium
 borohydride (LiBH₄)
 RL: CPS (Chemical process); PEP (Physical, engineering or chemical
 process); RCT (Reactant); PROC (Process); RACT (Reactant or reagent)
 (production and purification of hydrogen)
 RN 7440-23-5 HCAPLUS
 CN Sodium (8CI, 9CI) (CA INDEX NAME)

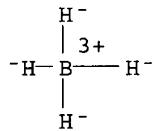
Na

RN 16940-66-2 HCAPLUS
 CN Borate(1-), tetrahydro-, sodium (8CI, 9CI) (CA INDEX NAME)



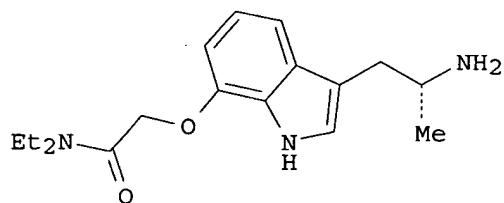
● Na^+

RN 16949-15-8 HCPLUS
CN Borate(1-), tetrahydro-, lithium (8CI, 9CI) (CA INDEX NAME)

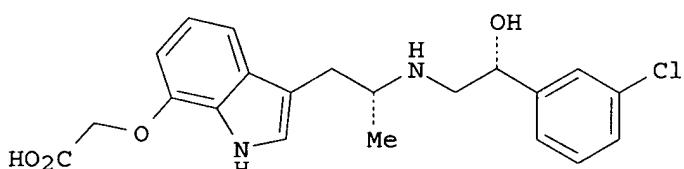


● Li^+

L33 ANSWER 4 OF 36 HCPLUS COPYRIGHT 2006 ACS on STN
 ACCESSION NUMBER: 2004:112383 HCPLUS
 DOCUMENT NUMBER: 140:321195
 TITLE: Process Development of a Scaleable Route to
 (2R)-[3-(2-Aminopropyl)-1H-indol-7-yloxy]-N,N-
 diethylacetamide: A Key Intermediate for AJ-9677, a
 Potent and Selective Human and Rat β 3-Adrenergic
 Receptor Agonist
 AUTHOR(S): Harada, Hiroshi; Fujii, Akihito; Odai, Osamu; Kato,
 Shiro
 CORPORATE SOURCE: Chemistry Research Laboratories, Dainippon
 Pharmaceutical Co., Ltd., Suita, Osaka, 564-0053,
 Japan
 SOURCE: Organic Process Research & Development (2004), 8(2),
 238-245
 CODEN: OPRDFK; ISSN: 1083-6160
 PUBLISHER: American Chemical Society
 DOCUMENT TYPE: Journal
 LANGUAGE: English
 OTHER SOURCE(S): CASREACT 140:321195
 GI



I



II



III

AB Nonracemic (aminopropylindoleoxy)acetamide I, an intermediate in the preparation of β 3-adrenergic receptor agonist AJ-9677 II, is prepared in five steps and 43% overall yield from 7-benzyloxy-1H-indole without chromatog. purification. The key step in the preparation of I is a regioselective and stereoselective acylation of 7-benzyloxy-1H-indole with N-Fmoc-D-alaninyl chloride; reduction of the acylindole intermediate with

sodium borohydride in a mixture of acetonitrile and isopropanol, and salt formation yields the nonracemic 1:1 oxalate salt of III. Protection of the terminal amine with Boc anhydride, a one-step debenzylation and alkylation using hydrogenation with **palladium** on carbon followed by alkylation of the phenol with N,N-di-Et chloroacetamide, removal of the Boc group with oxalic acid and treatment of the crystalline oxalate salt with potassium carbonate yields I. Deprotonation of 7-benzyloxy-1H-indole using methylmagnesium bromide is more effective than deprotonation with other Me Grignard reagents. N-Fmoc-D-alaninyl chloride is a more effective reagent in the preparation of I than other N-protected D-alaninyl chlorides; N-Cbz-D-alaninyl chloride does not react with the anion generated from 7-benzyloxy-1H-indole and methylmagnesium bromide, a phthaloyl-protected alaninyl chloride reacts but yields an acylindole which could not be reduced, and N-trifluoroacetyl-D-alaninyl chloride gives a lower yield in the acylation step. Reduction with **sodium borohydride** gives better results than other reducing agents such as **lithium borohydride**, borane-THF, and Vitride in the reduction of the acylindole intermediate in the preparation of I; using a mixed solvent system of acetonitrile and isopropanol gives better results than the use of either an alc. solvent alone or the use of other alcs. such as ethanol as cosolvents with acetonitrile. Oxalic acid salts of the aminopropylindole and the aminopropylindoleoxyacetamide intermediates in the preparation of I are crystalline, while salts with other inorg. and organic acids tried are amorphous.

Alkylation of the penultimate phenol intermediate with chloroacetic acid esters gives lactam products rather than the desired aminopropylindoleoxyacetate esters.

IT 676-58-4, Methylmagnesium chloride 917-64-6,
Methylmagnesium iodide
RL: RGT (Reagent); RACT (Reactant or reagent)
(less effective methylmagnesium halide reagent for the deprotonation of 7-benzyloxyindole in the regioselective acylation with N-protected D-alaninyl chlorides for the synthesis of the β 3-adrenergic receptor agonist AJ-9677)

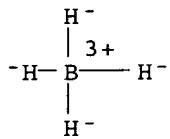
RN 676-58-4 HCPLUS
CN Magnesium, chloromethyl- (7CI, 8CI, 9CI) (CA INDEX NAME)

$\text{H}_3\text{C}-\text{Mg}-\text{Cl}$

RN 917-64-6 HCPLUS
CN Magnesium, iodomethyl- (7CI, 8CI, 9CI) (CA INDEX NAME)

$\text{H}_3\text{C}-\text{Mg}-\text{I}$

IT 16949-15-8, Lithium borohydride
RL: RGT (Reagent); RACT (Reactant or reagent)
(less effective reducing agent tried in the reduction of a nonracemic 3-acylindole in the preparation of an (aminopropylindoleoxy)acetamide intermediate for the synthesis of the β 3-adrenergic receptor agonist AJ-9677)
RN 16949-15-8 HCPLUS
CN Borate(1-), tetrahydro-, lithium (8CI, 9CI) (CA INDEX NAME)

● Li^+

IT 75-16-1, Methylmagnesium bromide

RL: RGT (Reagent); RACT (Reactant or reagent)

(optimal methylmagnesium halide reagent for the deprotonation of 7-benzyloxyindole in the regioselective acylation with N-protected D-alaninyl chlorides for the synthesis of the β_3 -adrenergic receptor agonist AJ-9677)

RN 75-16-1 HCPLUS

CN Magnesium, bromomethyl- (8CI, 9CI) (CA INDEX NAME)



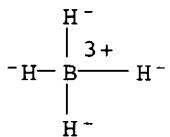
IT 16940-66-2, Sodium borohydride

RL: RGT (Reagent); RACT (Reactant or reagent)

(optimal reducing agent tried in the reduction of a nonracemic 3-acylindole in the preparation of an (aminopropylindoleoxy)acetamide intermediate for the synthesis of the β_3 -adrenergic receptor agonist AJ-9677)

RN 16940-66-2 HCPLUS

CN Borate(1-), tetrahydro-, sodium (8CI, 9CI) (CA INDEX NAME)

● Na^+

REFERENCE COUNT:

19

THERE ARE 19 CITED REFERENCES AVAILABLE FOR THIS RECORD. ALL CITATIONS AVAILABLE IN THE RE FORMAT

L33 ANSWER 5 OF 36 HCAPLUS COPYRIGHT 2006 ACS on STN

ACCESSION NUMBER: 2003:767751 HCAPLUS

DOCUMENT NUMBER: 139:276637

TITLE: Purification of ethylenic compounds having fluorinated organic group

INVENTOR(S): Matsuda, Takashi; Koike, Noriyuki; Oyama, Masayuki

PATENT ASSIGNEE(S): Shin-Etsu Chemical Industry Co., Ltd., Japan

SOURCE: Jpn. Kokai Tokkyo Koho, 5 pp.

CODEN: JKXXAF

DOCUMENT TYPE: Patent

LANGUAGE: Japanese

FAMILY ACC. NUM. COUNT: 1

PATENT INFORMATION:

PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
JP 2003277304	A2	20031002	JP 2002-75785	20020319
PRIORITY APPLN. INFO.:			JP 2002-75785	20020319

OTHER SOURCE(S): MARPAT 139:276637

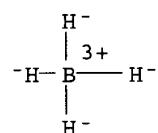
AB Rf(CH₂CH₂)_n (Rf= mono- or divalent F-containing organic group; n = 1, 2) are purified by treating with ≥1 selected from metal hydrides and metal-hydrogen complex compds. to eliminate I from iodine compds. contained as impurities and other components which interfere Pt-catalyzed hydrosilylation reaction. A mixture of C₄F₉CH₂CH₂ (I; 100 g, I₂ content 105 ppm), NaBH₄, NaOH solution, and Me₂CHOH was stirred at 60° for 1 h to give 97 g I containing ≤1 ppm I₂. The purified I was treated with MeSiHCl₂ in Me₂CHOH containing H₂PtCl₆ at 150° for 18 h to show conversion 98.2%.

IT 16940-66-2, Sodium borohydride

RL: RGT (Reagent); RACT (Reactant or reagent)
(purification of F-containing ethylenic compds. by eliminating I from impurity iodides using metal hydrides or metal -hydrogen complexes)

RN 16940-66-2 HCAPLUS

CN Borate(1-), tetrahydro-, sodium (8CI, 9CI) (CA INDEX NAME)

● Na⁺

L33 ANSWER 6 OF 36 HCAPLUS COPYRIGHT 2006 ACS on STN
 ACCESSION NUMBER: 2003:673831 HCAPLUS
 DOCUMENT NUMBER: 139:179794
 TITLE: Method for purification of methylal containing peroxides
 INVENTOR(S): Omori, Hideki; Shoji, Hiroshi; Nakamura, Tomizo
 PATENT ASSIGNEE(S): Maruzen Oil Co., Ltd., Japan
 SOURCE: Jpn. Kokai Tokkyo Koho, 4 pp.
 CODEN: JKXXAF
 DOCUMENT TYPE: Patent
 LANGUAGE: Japanese
 FAMILY ACC. NUM. COUNT: 1
 PATENT INFORMATION:

PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
JP 2003238471	A2	20030827	JP 2002-34153	20020212

PRIORITY APPLN. INFO.: JP 2002-34153 20020212
 AB The method includes reduction of methylal (I) containing peroxides and adsorption of the reduced products to remove HCO₂H (II). Thus, I containing 180 ppm peroxide was hydrogenolyzed over Raney Ni to give a product (peroxide content <2 ppm, II content 60 ppm), which was treated with Mol. Sieve 4A, showing II content <1 ppm.

IT 1318-93-0, Montmorillonite, uses
 RL: NUU (Other use, unclassified); USES (Uses)
 (adsorbent; purification of methylal containing peroxides by reduction and adsorption)
 RN 1318-93-0 HCAPLUS
 CN Montmorillonite ((Al_{1.33}-1.67Mg_{0.33}-0.67) (Ca₀₋₁Na₀₋₁)_{0.33}Si₄(OH)₂O₁₀.xH₂O) (9CI) (CA INDEX NAME)

*** STRUCTURE DIAGRAM IS NOT AVAILABLE ***

IT 1344-28-1, Alumina, uses
 RL: CAT (Catalyst use); NUU (Other use, unclassified); USES (Uses)
 (hydrogenolysis catalyst support and adsorbent; purification of methylal containing peroxides by reduction and adsorption)
 RN 1344-28-1 HCAPLUS
 CN Aluminum oxide (Al₂O₃) (8CI, 9CI) (CA INDEX NAME)

*** STRUCTURE DIAGRAM IS NOT AVAILABLE ***

IT 7440-02-0, Raney nickel, uses 7440-05-3, Palladium, uses 7440-06-4, Platinum, uses
 RL: CAT (Catalyst use); USES (Uses)
 (hydrogenolysis catalyst; purification of methylal containing peroxides by reduction and adsorption)
 RN 7440-02-0 HCAPLUS
 CN Nickel (8CI, 9CI) (CA INDEX NAME)

Ni

RN 7440-05-3 HCAPLUS
 CN Palladium (8CI, 9CI) (CA INDEX NAME)

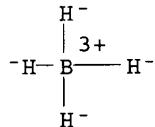
Pd

Valenrod 10_721479

RN 7440-06-4 HCAPLUS
CN Platinum (8CI, 9CI) (CA INDEX NAME)

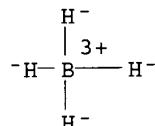
Pt

IT 16940-66-2, Sodium borohydride
16971-29-2D, Borohydride, trimethoxy-, salts 33195-00-5D
, Cyanoborohydride, salts
RL: RGT (Reagent); RACT (Reactant or reagent)
(reducing agent; purification of methylal containing peroxides by
reduction and adsorption)
RN 16940-66-2 HCAPLUS
CN Borate(1-), tetrahydro-, sodium (8CI, 9CI) (CA INDEX NAME)

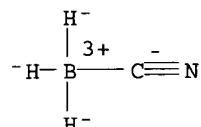


● Na⁺

RN 16971-29-2 HCAPLUS
CN Borate(1-), tetrahydro- (8CI, 9CI) (CA INDEX NAME)



RN 33195-00-5 HCAPLUS
CN Borate(1-), (cyano- κ C)trihydro-, (T-4)- (9CI) (CA INDEX NAME)



L33 ANSWER 7 OF 36 HCPLUS COPYRIGHT 2006 ACS on STN
 ACCESSION NUMBER: 2003:22609 HCPLUS
 DOCUMENT NUMBER: 138:94840
 TITLE: Extraction and recovery of ions from a solution
 INVENTOR(S): Strauss, Steven H.; Odom, Matthew A.; Clapsaddle, Brady
 PATENT ASSIGNEE(S): Colorado State University Research Foundation, USA
 SOURCE: PCT Int. Appl., 32 pp.
 CODEN: PIXXD2
 DOCUMENT TYPE: Patent
 LANGUAGE: English
 FAMILY ACC. NUM. COUNT: 1
 PATENT INFORMATION:

PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
WO 2003001895	A2	20030109	WO 2002-US20973	20020627
WO 2003001895	A3	20030227		
W: AE, AG, AL, AM, AT, AU, AZ, BA, BB, BG, BR, BY, BZ, CA, CH, CN, CO, CR, CU, CZ, DE, DK, DM, DZ, EC, EE, ES, FI, GB, GD, GE, GH, GM, HR, HU, ID, IL, IN, IS, JP, KE, KG, KP, KR, KZ, LC, LK, LR, LS, LT, LU, LV, MA, MD, MG, MK, MN, MW, MX, MZ, NO, NZ, OM, PH, PL, PT, RO, RU, SD, SE, SG, SI, SK, SL, TJ, TM, TN, TR, TT, TZ, UA, UG, US, UZ, VN, YU, ZA, ZM, ZW, AM, AZ, BY, KG, KZ, MD, RU, TJ, TM				
RW: GH, GM, KE, LS, MW, MZ, SD, SL, SZ, TZ, UG, ZM, ZW, AT, BE, CH, CY, DE, DK, ES, FI, FR, GB, GR, IE, IT, LU, MC, NL, PT, SE, TR, BF, BJ, CF, CG, CI, CM, GA, GN, GQ, GW, ML, MR, NE, SN, TD, TG				
US 2005145571	A1	20050707	US 2003-501884	20020627
PRIORITY APPLN. INFO.:			US 2001-302052P	P 20010629
			WO 2002-US20973	W 20020627

AB The present invention provides organometallic redox-recyclable solid ion-extractant and a method for using the same in extracting and recovering perfluoroalkyl carboxylate or perfluoroalkyl sulfonate ions in a solution. The solid ion extractant is a composition of the formula [ZX1Si(X2R2)(X3R3)R1Ar1MAr2]Ya, wherein: each of Ar1 and Ar2 is independently C4-C20 aryl selected from the group consisting of cyclopentadienyl, dicarbollide and Ph, each of which can be optionally substituted; M is a transition metal selected from Fe, Ru, Mn, Co, Ni, Cr, Os, Rh and Ir; R1 is C2-C20 alkylene; each of X1, X2 and X3 is independently a bond, O, S, or NR4; each of R2, R3 and R4 is independently H, or C1-C6 alkyl; Z is a solid support selected from glass, metal or a polymeric resin; Y is an anion selected from NO3-, halide, HSO4-, ClO4-, ReO4-, PF6-, carboxylate and CF3SO3-; and a is 0 when said ion coordinating agent is deactivated, and a is an integer from 1-3 when said ion coordinating agent is activated. In particular, the present invention provides a method for extracting and recovering ions which are relatively insol. in a solvent at a particular solvent temperature range. More specifically, the present invention provides a method for extracting and recovering perfluoroalkylsulfonates from an aqueous solution, such as

wastewaters
 from firefighting.

IT 7440-06-4D, Platinum, silyl compds., complexes of
 RL: CAT (Catalyst use); USES (Uses)
 (hydrosilylation catalyst; preparation and use of a recyclable solid extractant for extraction and recovery of perfluoroalkyl carboxylate or perfluoroalkyl sulfonate ions from wastewater or other aqueous solns.)

RN 7440-06-4 HCPLUS

CN Platinum (8CI, 9CI) (CA INDEX NAME)

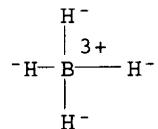
Pt

IT 16940-66-2, Sodium borohydride

RL: CPS (Chemical process); PEP (Physical, engineering or chemical process); PROC (Process)
(reductant to deactivate extractant; preparation and use of a recyclable solid extractant for extraction and **recovery** of perfluoroalkyl carboxylate or perfluoroalkyl sulfonate ions from wastewater or other aqueous solns.)

RN 16940-66-2 HCPLUS

CN Borate(1-), tetrahydro-, sodium (8CI, 9CI) (CA INDEX NAME)



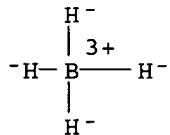
● Na⁺

L33 ANSWER 8 OF 36 HCPLUS COPYRIGHT 2006 ACS on STN
 ACCESSION NUMBER: 2002:827745 HCPLUS
 DOCUMENT NUMBER: 137:341465
 TITLE: Method and apparatus for generation of clean air moist
 for indoor air purification
 INVENTOR(S): Suda, Seijiro
 PATENT ASSIGNEE(S): Japan
 SOURCE: Jpn. Kokai Tokkyo Koho, 6 pp.
 CODEN: JKXXAF
 DOCUMENT TYPE: Patent
 LANGUAGE: Japanese
 FAMILY ACC. NUM. COUNT: 1
 PATENT INFORMATION:

PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
JP 2002317980	A2	20021031	JP 2001-117286	20010416
PRIORITY APPLN. INFO.:			JP 2001-117286	20010416
AB	The method is carried out by contacting alkaline aqueous solution containing metal-hydrogen complex compound (e.g., NaBH4) with H2-generation catalyst (e.g., fluorinated Ni-alloy) to form H2 gas, then mixing the H2 with O2-excess gas (e.g., air) in the presence of oxidation catalyst (e.g., Pd) for low-temperature combustion to produce air having water content, to be supplied to the indoor air via a pipeline.			
IT	7440-05-3, Palladium, uses RL: CAT (Catalyst use); DEV (Device component use); NUU (Other use, unclassified); USES (Uses) (method and apparatus for generation of clean air moist for indoor air purification)			
RN	7440-05-3 HCPLUS			
CN	Palladium (8CI, 9CI) (CA INDEX NAME)			

Pd

IT 16940-66-2, Sodium borohydride (NaBH4)
 RL: NUU (Other use, unclassified); RCT (Reactant); RACT (Reactant or reagent); USES (Uses)
 (method and apparatus for generation of clean air moist for indoor air purification)
 RN 16940-66-2 HCPLUS
 CN Borate(1-), tetrahydro-, sodium (8CI, 9CI) (CA INDEX NAME)

● Na⁺

L33 ANSWER 9 OF 36 HCAPLUS COPYRIGHT 2006 ACS on STN
ACCESSION NUMBER: 2002:638003 HCAPLUS
DOCUMENT NUMBER: 137:188572
TITLE: Acidic redox leaching for recovery of precious
metals from low-concentration sources
INVENTOR(S): Farone, William A.; Azad, Maryam H.
PATENT ASSIGNEE(S): USA
SOURCE: U.S. Pat. Appl. Publ., 11 pp.
CODEN: USXXCO
DOCUMENT TYPE: Patent
LANGUAGE: English
FAMILY ACC. NUM. COUNT: 1
PATENT INFORMATION:

PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
US 2002112569	A1	20020822	US 2001-789270	20010220
US 6551378	B2	20030422		

PRIORITY APPLN. INFO.: US 2001-789270 20010220

AB The Au and Pt-group metals in powdered low-concentration sources (especially ore tailings) are recovered by leaching with aqueous solution containing HCl and an oxidizing agent (typically Cl₂), followed by separation of the leach solution, extraction of the associated base metals, and reduction with precipitation of the precious metals for recovery. The residual solids from acidic leaching are neutralized with a base, and most of the base metals are precipitated as hydroxides. The process is environmentally acceptable, uses no organic solvents, and is suitable for leaching of ore tailings and wastes containing ppm amts. of precious metals.

IT 7440-06-4, Platinum, processes 7440-57-5,
Gold, processes
RL: EPR (Engineering process); PEP (Physical, engineering or chemical process); PROC (Process)
(leaching of; acidic redox leaching for recovery of precious metals from low-concentration sources)

RN 7440-06-4 HCAPLUS
CN Platinum (8CI, 9CI) (CA INDEX NAME)

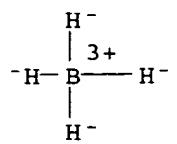
Pt

RN 7440-57-5 HCAPLUS
CN Gold (8CI, 9CI) (CA INDEX NAME)

Au

IT 16940-66-2, Sodium borohydride
RL: CPS (Chemical process); PEP (Physical, engineering or chemical process); PROC (Process)
(reduction with, in leach solution; acidic redox leaching for recovery of precious metals from low-concentration sources)

RN 16940-66-2 HCAPLUS
CN Borate(1-), tetrahydro-, sodium (8CI, 9CI) (CA INDEX NAME)



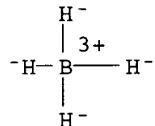
● Na^+

L33 ANSWER 10 OF 36 HCPLUS COPYRIGHT 2006 ACS on STN
 ACCESSION NUMBER: 2001:91311 HCPLUS
 DOCUMENT NUMBER: 134:118619
 TITLE: Method for recovery of palladium from spent catalyst solution for electroless coating
 INVENTOR(S): Nakamura, Tomonobu; Tsuchida, Kazunori
 PATENT ASSIGNEE(S): Matsuda Sangyo K. K., Japan
 SOURCE: Jpn. Kokai Tokkyo Koho, 7 pp.
 CODEN: JKXXAF
 DOCUMENT TYPE: Patent
 LANGUAGE: Japanese
 FAMILY ACC. NUM. COUNT: 1
 PATENT INFORMATION:

PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
JP 2001032025	A2	20010206	JP 1999-204878	19990719
PRIORITY APPLN. INFO.: JP 1999-204878 19990719				
AB	The method comprises reducing with Na borohydride.			
IT	7440-05-3P, Palladium, preparation RL: PUR (Purification or recovery); PREP (Preparation) (recovery of palladium from spent catalyst solution for electroless coating by reducing with sodium borohydride)			
RN	7440-05-3 HCPLUS			
CN	Palladium (8CI, 9CI) (CA INDEX NAME)			

Pd

IT 16940-66-2, Sodium borohydride
 RL: RCT (Reactant); RACT (Reactant or reagent)
 (recovery of palladium from spent catalyst solution
 for electroless coating by reducing with sodium
 borohydride)
 RN 16940-66-2 HCPLUS
 CN Borate(1-), tetrahydro-, sodium (8CI, 9CI) (CA INDEX NAME)

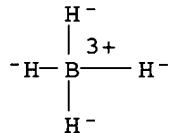
● Na⁺

L33 ANSWER 11 OF 36 HCAPLUS COPYRIGHT 2006 ACS on STN
 ACCESSION NUMBER: 1998:42117 HCAPLUS
 DOCUMENT NUMBER: 128:104629
 TITLE: Method for recovery of palladium from spent catalysts
 INVENTOR(S): Nakazawa, Hiroyuki
 PATENT ASSIGNEE(S): Japan
 SOURCE: Jpn. Kokai Tokkyo Koho, 3 pp.
 CODEN: JKXXAF
 DOCUMENT TYPE: Patent
 LANGUAGE: Japanese
 FAMILY ACC. NUM. COUNT: 1
 PATENT INFORMATION:

PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
JP 10008155	A2	19980113	JP 1996-192651	19960619
PRIORITY APPLN. INFO.:			JP 1996-192651	19960619
AB	The method comprises dissolving a spent catalyst with aqua regia, precipitating the dissolved Pd as iodide for separation, and reducing the iodide with a reducing agent selected from NaBH4, hydrazine hydrochloride, hydrazine sulfate, hydrazine, Na2SO3, Na2S2O6, formic acid, oxalic acid.			
IT	7440-05-3P, Palladium, preparation RL: PUR (Purification or recovery); PREP (Preparation) (recovery of palladium from spent catalysts by separation as iodide and reduction)			
RN	7440-05-3 HCAPLUS			
CN	Palladium (8CI, 9CI) (CA INDEX NAME)			

Pd

IT 16940-66-2, Sodium borohydride
 RL: TEM (Technical or engineered material use); USES (Uses)
 (recovery of palladium from spent catalysts by separation as iodide and reduction)
 RN 16940-66-2 HCAPLUS
 CN Borate(1-), tetrahydro-, sodium (8CI, 9CI) (CA INDEX NAME)

● Na⁺

L33 ANSWER 12 OF 36 HCPLUS COPYRIGHT 2006 ACS on STN
 ACCESSION NUMBER: 1997:93860 HCPLUS
 DOCUMENT NUMBER: 126:106889
 TITLE: Separation and recovery of platinum-group
 metals in solutions with high efficiency
 INVENTOR(S): Kawasaki, Hajime; Amino, Maki; Nishimura, Kenji
 PATENT ASSIGNEE(S): Mitsubishi Materials Corp, Japan
 SOURCE: Jpn. Kokai Tokkyo Koho, 10 pp.
 CODEN: JKXXAF
 DOCUMENT TYPE: Patent
 LANGUAGE: Japanese
 FAMILY ACC. NUM. COUNT: 1
 PATENT INFORMATION:

PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
JP 08291345	A2	19961105	JP 1995-95134	19950420
JP 3348808	B2	20021120		

PRIORITY APPLN. INFO.: JP 1995-95134 19950420

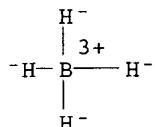
AB The process comprises (1) formation of complexes by adding complexing agents selected from ≥ 1 phosphines, phosphites, dialkylsulfides, cyclopentadiene, and/or CO in aqueous solns. or organic solns. containing Pt-group metal salts, (2) optional solid-liquid separation, (3) extraction of the complexes by contacting with supercrit. fluids, and (4) blowing into NH₄OH aqueous solns. containing N₂H₄, NaBH₄, thiourea, or NH₄Cl for recovery of Pt-group metals. The phosphines may be selected from trimethylphosphine, triethylphosphine, tripropylphosphine, or tributylphosphine. The phosphites may be selected from trimethylphosphite, triethylphosphite, tripropylphosphite, or tributylphosphite. The dialkylsulfides may be selected from butylsulfide, hexylsulfide, or octylsulfide. The supercrit. fluids. may be selected from CO₂, fluorocarbons, or N₂O.

IT 16940-66-2, Sodium borohydride (NaBH₄)

RL: NUU (Other use, unclassified); USES (Uses)
 (blowing of ammonia aqueous solns. containing; platinum-group
 metals separation and recovery by complexation and extraction
 with supercrit. fluids)

RN 16940-66-2 HCPLUS

CN Borate(1-), tetrahydro-, sodium (8CI, 9CI) (CA INDEX NAME)



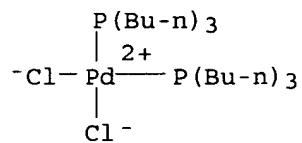
● Na⁺

IT 14977-08-3P, Dichlorobistributylphosphinepalladium
 RL: PEP (Physical, engineering or chemical process); PNU (Preparation,
 unclassified); PREP (Preparation); PROC (Process)
 (platinum-group metals separation and recovery by
 complexation and extraction with supercrit. fluids)

RN 14977-08-3 HCPLUS

CN Palladium, dichlorobis(tributylphosphine)- (7CI, 8CI, 9CI) (CA INDEX)

NAME)



IT 7440-05-3P, Palladium, preparation
RL: PUR (Purification or recovery); PREP (Preparation)
(platinum-group metals separation and recovery by
complexation and extraction with supercrit. fluids)
RN 7440-05-3 HCPLUS
CN Palladium (8CI, 9CI) (CA INDEX NAME)

Pd

L33 ANSWER 13 OF 36 HCPLUS COPYRIGHT 2006 ACS on STN
ACCESSION NUMBER: 1996:485373 HCPLUS
DOCUMENT NUMBER: 125:201342
TITLE: Precipitation behavior of metallic **palladium**
and **platinum** in aqueous solutions
AUTHOR(S): Sakurai, Hiroki; Hirokawa, Kichinosuke
CORPORATE SOURCE: Inst. Mater. Res., Tohoku Univ., Sendai, 980-77, Japan
SOURCE: Bunseki Kagaku (1996), 45(8), 795-798
CODEN: BNSKAK; ISSN: 0525-1931
PUBLISHER: Nippon Bunseki Kagakkai
DOCUMENT TYPE: Journal
LANGUAGE: Japanese

AB **Pd** and **Pt** were reduced and precipitated in aqueous acidic solns.
by **sodium tetrahydroborate**, sodium phosphinate and
ascorbic acid. **Pd** was reduced by all the reducing agents, but
Pt was reduced by **sodium tetrahydroborate**
only. Both **metal** could be reduced individually, but when they
were co-existed, complete separation could not be achieved by these reduction

for
the generation of active hydrogen during reduction of **Pd**. The
technic is applicable to recycle and **recover Pt** and
Pd from exhaust gas-purification waste catalysts.

IT 7440-05-3P, **Palladium**, preparation 7440-06-4P,
Platinum, preparation
RL: CAT (Catalyst use); PEP (Physical, engineering or chemical process);
PUR (Purification or recovery); PREP (Preparation); PROC (Process); USES
(Uses)
(precipitation behavior of metallic **palladium** and **platinum**
in aqueous acidic solns. for separation recovery)

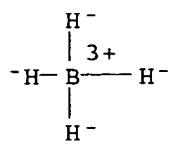
RN 7440-05-3 HCPLUS
CN Palladium (8CI, 9CI) (CA INDEX NAME)

Pd

RN 7440-06-4 HCPLUS
CN Platinum (8CI, 9CI) (CA INDEX NAME)

Pt

IT 16940-66-2, **Sodium tetrahydroborate**
RL: NUU (Other use, unclassified); USES (Uses)
(reducing agent; precipitation behavior of metallic **palladium** and
platinum in aqueous acidic solns. for separation **recovery**)
RN 16940-66-2 HCPLUS
CN Borate(1-), tetrahydro-, sodium (8CI, 9CI) (CA INDEX NAME)



● Na^+

L33 ANSWER 14 OF 36 HCPLUS COPYRIGHT 2006 ACS on STN
 ACCESSION NUMBER: 1996:148121 HCPLUS
 DOCUMENT NUMBER: 124:261743
 TITLE: Preparation of N-(long-chain acyl)amino acids and
 their salts
 INVENTOR(S): Kaneko, Yohei; Nishimoto, Yoshifumi
 PATENT ASSIGNEE(S): Kao Corp, Japan
 SOURCE: Jpn. Kokai Tokkyo Koho, 6 pp.
 CODEN: JKXXAF
 DOCUMENT TYPE: Patent
 LANGUAGE: Japanese
 FAMILY ACC. NUM. COUNT: 1
 PATENT INFORMATION:

PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
JP 07330699	A2	19951219	JP 1994-120924	19940602
PRIORITY APPLN. INFO.:			JP 1994-120924	19940602

OTHER SOURCE(S): MARPAT 124:261743

AB The title salts, useful as surfactants and bacteriostats, are prepared by hydrolysis of R1CONR2R3CN [COR1 = C8-22 (un)saturated aliphatic acyl; R2 = H, C1-3 linear or branched alkyl; R3 = C1-5 linear or branched alkylene] in the presence of basic substances, followed by treatment of the resulting R1CONR2R3CO2M (M = cation) (I) with metal hydrides, by treatment of the resulting I with oxidizing agents, or by hydrogenation of the resulting I in the presence of hydrogenation catalysts. The salts are further treated with mineral acids to adjust pH at 1-5 to give the corresponding acyl amino acids. These methods provide the products with no offensive odor and coloration. A mixture of 70 g N-lauroyl- β -aminopropionitrile, H₂O, and an aqueous KOH solution was stirred at 90° for 15 h and the crude N-lauroyl- β -alaninate (II) obtained was further treated with NaBH₄ at 60° for 2 h to give 250 g aqueous solution containing 32.8% II with APHA color 100. After removal of NH₃ from the solution

by evaporation, no offensive odor was detected.

IT 7440-02-0, Nickel, uses

RL: CAT (Catalyst use); USES (Uses)
 (Raney; preparation of (higher acyl)amino acids (salts) by hydrolysis of amidonitriles followed by chemical purification)

RN 7440-02-0 HCPLUS

CN Nickel (8CI, 9CI) (CA INDEX NAME)

Ni

IT 7440-05-3, Palladium, uses 7440-06-4,
 Platinum, uses

RL: CAT (Catalyst use); USES (Uses)
 (preparation of (higher acyl)amino acids (salts) by hydrolysis of amidonitriles followed by chemical purification)

RN 7440-05-3 HCPLUS

CN Palladium (8CI, 9CI) (CA INDEX NAME)

Pd

RN 7440-06-4 HCPLUS

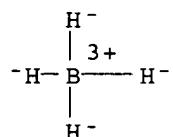
CN Platinum (8CI, 9CI) (CA INDEX NAME)

Pt

IT 11138-47-9, Sodium perborate 16940-66-2, Sodium
borohydride
RL: RCT (Reactant); RACT (Reactant or reagent)
(preparation of (higher acyl)amino acids (salts) by hydrolysis of
amidonitriles followed by chemical purification)
RN 11138-47-9 HCPLUS
CN Perboric acid, sodium salt (8CI, 9CI) (CA INDEX NAME)

*** STRUCTURE DIAGRAM IS NOT AVAILABLE ***

RN 16940-66-2 HCPLUS
CN Borate(1-), tetrahydro-, sodium (8CI, 9CI) (CA INDEX NAME)



● Na^+

L33 ANSWER 15 OF 36 HCAPLUS COPYRIGHT 2006 ACS on STN
ACCESSION NUMBER: 1995:964911 HCAPLUS
DOCUMENT NUMBER: 124:96495
TITLE: Process for removing oils and greases from industrial
wastewaters
INVENTOR(S): Guess, Robert G.
PATENT ASSIGNEE(S): Romar Technologies, Inc., USA
SOURCE: U.S., 12 pp. Cont.-in-part of U.S. Ser. No. 160,240,
abandoned.
CODEN: USXXAM
DOCUMENT TYPE: Patent
LANGUAGE: English
FAMILY ACC. NUM. COUNT: 6
PATENT INFORMATION:

PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
US 5462670	A	19951031	US 1994-218891	19940328
US 5122279	A	19920616	US 1991-682129	19910408
US 5298168	A	19940329	US 1992-893978	19920603
US 5545331	A	19960813	US 1995-386700	19950210
PRIORITY APPLN. INFO.:			US 1991-682129	A2 19910408
			US 1992-893978	A3 19920603
			US 1993-160240	B2 19931202
			US 1994-218891	A2 19940328

AB A process for removing dissolved oils and greases from an aqueous solution
which

also may contain dissolved heavy **metals** is provided wherein the
aqueous solution is mixed with a source of ferrous ion and dithionite ion in a
first step at acidic pH to reduce and permit removal of solid heavy
metal, is present and to sep. oils and greases from the aqueous solution
The solution from the first step is reacted in a second step with hydroxide
slurries obtained from third and fourth steps. A second step solution from
the second step is reacted in a third step with an alkali composition and a
third solution Optionally, the third solution is reacted with a chelating
agent

for iron and an oxidizer in a fourth step. A solution of chelated iron from
the fourth step, when practical is disposed of. Oils and greases are
recovered from the first step such as by skimming, and the heavy
metals are recovered. The method provides effective treatment of
contaminated groundwater, landfill leachate, sewage, mine drainage, slag
pile drainage, mine tailings drainage, plating wastes, and etching wastes.

IT 7429-90-5, Aluminum, uses 7439-89-6, Iron, uses

7440-66-6, Zinc, uses 16940-66-2, Sodium
borohydride

RL: NUU (Other use, unclassified); USES (Uses)
(recovering heavy **metals**, oils and greases from
industrial wastewaters in multistage process)

RN 7429-90-5 HCAPLUS

CN Aluminum (8CI, 9CI) (CA INDEX NAME)

A1

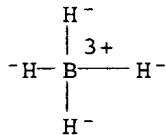
RN 7439-89-6 HCAPLUS
CN Iron (7CI, 8CI, 9CI) (CA INDEX NAME)

Fe

RN 7440-66-6 HCPLUS
CN Zinc (7CI, 8CI, 9CI) (CA INDEX NAME)

Zn

RN 16940-66-2 HCPLUS
CN Borate(1-), tetrahydro-, sodium (8CI, 9CI) (CA INDEX NAME)



● Na^+

IT 7439-92-1P, Lead, preparation 7440-02-0P, Nickel, preparation 7440-05-3P, Palladium, preparation 7440-06-4P, Platinum, preparation 7440-22-4P, Silver, preparation 7440-31-5P, Tin, preparation 7440-48-4P, Cobalt, preparation 7440-50-8P, Copper, preparation 7440-57-5P, Gold, preparation
RL: PUR (Purification or recovery); PREP (Preparation)
(recovering heavy metals, oils and greases from industrial wastewaters in multistage process)

RN 7439-92-1 HCPLUS
CN Lead (8CI, 9CI) (CA INDEX NAME)

Pb

RN 7440-02-0 HCPLUS
CN Nickel (8CI, 9CI) (CA INDEX NAME)

Ni

RN 7440-05-3 HCPLUS
CN Palladium (8CI, 9CI) (CA INDEX NAME)

Pd

RN 7440-06-4 HCPLUS
CN Platinum (8CI, 9CI) (CA INDEX NAME)

Pt

RN 7440-22-4 HCAPLUS
CN Silver (8CI, 9CI) (CA INDEX NAME)

Ag

RN 7440-31-5 HCAPLUS
CN Tin (8CI, 9CI) (CA INDEX NAME)

Sn

RN 7440-48-4 HCAPLUS
CN Cobalt (8CI, 9CI) (CA INDEX NAME)

Co

RN 7440-50-8 HCAPLUS
CN Copper (7CI, 8CI, 9CI) (CA INDEX NAME)

Cu

RN 7440-57-5 HCAPLUS
CN Gold (8CI, 9CI) (CA INDEX NAME)

Au

L33 ANSWER 16 OF 36 HCAPLUS COPYRIGHT 2006 ACS on STN
 ACCESSION NUMBER: 1995:563523 HCAPLUS
 DOCUMENT NUMBER: 122:295785
 TITLE: Ore leaching with non-cyanide redox solution for recovery of precious **metals**
 INVENTOR(S): Martenson, Irvin; Mathis, Samuel A.
 PATENT ASSIGNEE(S): USA
 SOURCE: U.S., 6 pp.
 CODEN: USXXAM
 DOCUMENT TYPE: Patent
 LANGUAGE: English
 FAMILY ACC. NUM. COUNT: 1
 PATENT INFORMATION:

PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
US 5401296	A	19950328	US 1994-266563	19940628
CA 2194137	AA	19960111	CA 1995-2194137	19950324
WO 9600801	A1	19960111	WO 1995-US3734	19950324
W: AU, BR, CA, CN, MX, RU				
AU 9521949	A1	19960125	AU 1995-21949	19950324
ZA 9502471	A	19960415	ZA 1995-2471	19950327
PRIORITY APPLN. INFO.:			US 1994-266563	A 19940628
			WO 1995-US3734	W 19950324

AB The aqueous slurry with 20-25% ore powder (size -250 mesh) for leaching of Au, Ag, and/or **Pt**-group **metals** is based on the aqueous solution containing NaBr 1.5-3.5, NaCl 3-6%, K2S2O8 for the redox potential of 700-900 mV, and HCl for the pH of 3-4. The resulting leach liquor is pumped to a holding tank, and reacted with aqueous solution of NaOH and NaBH4 at pH of 8.2-8.6 to precipitate the precious **metals** for recovery by filtration, followed by conventional separation and purification. The liquor from filtration is

electrochem. treated with replenishment of the salts, and recycled to the leaching stage. The K2S2O8 is preferably com. Oxone triple salt. The redox leaching process is suitable for recovery of **Pt**, Ir, and **Pd** from volcanic ash.

IT 7440-05-3P, Palladium, preparation 7440-06-4P,
 Platinum, preparation 7440-22-4P, Silver, preparation
 7440-57-5P, Gold, preparation
 RL: PUR (Purification or recovery); PREP (Preparation)
 (leaching of; ore leaching with non-cyanide redox solution for recovery of precious **metals**)
 RN 7440-05-3 HCAPLUS
 CN Palladium (8CI, 9CI) (CA INDEX NAME)

Pd

RN 7440-06-4 HCAPLUS
 CN Platinum (8CI, 9CI) (CA INDEX NAME)

Pt

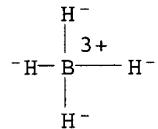
RN 7440-22-4 HCAPLUS
 CN Silver (8CI, 9CI) (CA INDEX NAME)

Ag

RN 7440-57-5 HCAPLUS
CN Gold (8CI, 9CI) (CA INDEX NAME)

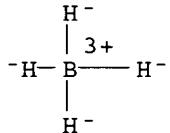
Au

IT 16940-66-2, Sodium borohydride (NaBH4)
RL: MOA (Modifier or additive use); USES (Uses)
(redox solution containing; ore leaching with non-cyanide redox solution for
recovery of precious metals)
RN 16940-66-2 HCAPLUS
CN Borate(1-), tetrahydro-, sodium (8CI, 9CI) (CA INDEX NAME)



● Na⁺

L33 ANSWER 17 OF 36 HCPLUS COPYRIGHT 2006 ACS on STN
 ACCESSION NUMBER: 1994:196570 HCPLUS
 DOCUMENT NUMBER: 120:196570
 TITLE: Rhodium, platinum, and palladium
 recovery from new and spent automotive catalysts
 AUTHOR(S): Wu, Kuo Ying Amanda; Wisecarver, Keith D.; Abraham, Martin A.; Takach, Nicholas; Yang, Ningsheng
 CORPORATE SOURCE: Dep. Chem. Eng., Univ. Tulsa, Tulsa, OK, 74104, USA
 SOURCE: Precious Metals (1993), 17th, 343-9
 CODEN: PRCMEU; ISSN: 8756-0917
 DOCUMENT TYPE: Journal
 LANGUAGE: English
 AB The recovery of Pt, Pa, and Rh automotive catalysts by leaching HCl + H₂SO₄ with the addition of F- and H₂O₂ including pretreatment with Na borohydride, sodium formate and NaOH was studied. The recovery of metals mainly depends on the catalyst type and occurs in the following sequence Pd>Pt >Rh. Rh recovery of 90% or better can be achieved from both new and spent catalysts using pretreatment with sodium formate or NaOH.
 IT 16940-66-2, Sodium borohydride
 RL: PROC (Process)
 (recovery of palladium and platinum and rhodium from spent automotive catalyst by leaching with acids and pretreatment with)
 RN 16940-66-2 HCPLUS
 CN Borate(1-), tetrahydro-, sodium (8CI, 9CI) (CA INDEX NAME)



● Na⁺

IT 7440-05-3P, Palladium, preparation 7440-06-4P, Platinum, preparation
 RL: PUR (Purification or recovery); PREP (Preparation)
 (recovery of, from spent automotive catalyst by pretreatment and acid leaching)
 RN 7440-05-3 HCPLUS
 CN Palladium (8CI, 9CI) (CA INDEX NAME)

Pd

RN 7440-06-4 HCPLUS
 CN Platinum (8CI, 9CI) (CA INDEX NAME)

Pt

Valenrod 10_721479

L33 ANSWER 18 OF 36 HCAPLUS COPYRIGHT 2006 ACS on STN
 ACCESSION NUMBER: 1994:57151 HCAPLUS
 DOCUMENT NUMBER: 120:57151
 TITLE: Recovery of noble metals from solutions produced in oxidation reactions involving alkyl nitrates by treatment with reducing agents
 INVENTOR(S): Landscheidt, Heinz; Klausener, Alexander; Blank, Heinz Ulrich
 PATENT ASSIGNEE(S): Bayer A.-G., Germany
 SOURCE: Ger. Offen., 6 pp.
 CODEN: GWXXBX
 DOCUMENT TYPE: Patent
 LANGUAGE: German
 FAMILY ACC. NUM. COUNT: 1
 PATENT INFORMATION:

PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
DE 4137965	A1	19930527	DE 1991-4137965	19911119
EP 546324	A1	19930616	EP 1992-119030	19921106
EP 546324	B1	19950419		
R: AT, BE, CH, DE, FR, GB, IT, LI, NL				
AT 121316	E	19950515	AT 1992-119030	19921106
JP 05239567	A2	19930917	JP 1992-329900	19921117
PRIORITY APPLN. INFO.:			DE 1991-4137965	A 19911119

AB Wastewaters containing metals or metal salts are treated with a reducing agent and an additive to reduce the metals, which are recovered in the metallic form. Suitable reducing agents include H₂, CO, or a complex hydride, preferably a boron hydride of Group I or II metals, especially Na, K, Li, or Zn. The additive is a primary or secondary amino group which can react with nitrous acid, e.g., ammonia, an ammonium salt, ammonium chloride, Me ammonium chloride, urea, or amidosulfonic acid. The metallic form can be changed to a salt form by oxidation with an oxidizing agent in an aqueous acid solution, e.g., HCl or acetic

acid. Suitable oxidizing agents include chlorates or nitrates of alkali or alkaline earth metals or ammonium ion, HNO₃, or H₂O₂. The method is suitable for recovering Pt group metals, especially

Pd.

IT 7647-10-1, Palladium chloride
 RL: CAT (Catalyst use); USES (Uses)
 (catalysts, for Me acrylate reaction with Me nitrite, recovery of)
 RN 7647-10-1 HCAPLUS
 CN Palladium chloride (PdCl₂) (6CI, 8CI, 9CI) (CA INDEX NAME)

Cl-Pd-Cl

IT 7440-05-3P, Palladium, preparation
 RL: PREP (Preparation)
 (recovery of, from oxidation reaction mixts.)
 RN 7440-05-3 HCAPLUS
 CN Palladium (8CI, 9CI) (CA INDEX NAME)

Pd

IT 13762-51-1 16940-66-2 16949-15-8,

Valenrod 10_721479

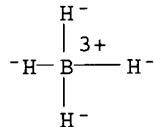
Lithium boron hydride 17611-70-0

RL: USES (Uses)

(reducing agents, in **recovery** of noble metal
catalysts)

RN 13762-51-1 HCPLUS

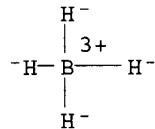
CN Borate(1-), tetrahydro-, potassium (8CI, 9CI) (CA INDEX NAME)



● K⁺

RN 16940-66-2 HCPLUS

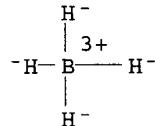
CN Borate(1-), tetrahydro-, sodium (8CI, 9CI) (CA INDEX NAME)



● Na⁺

RN 16949-15-8 HCPLUS

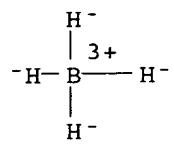
CN Borate(1-), tetrahydro-, lithium (8CI, 9CI) (CA INDEX NAME)



● Li⁺

RN 17611-70-0 HCPLUS

CN Borate(1-), tetrahydro-, zinc (2:1) (9CI) (CA INDEX NAME)



● $1/2 \text{ Zn}^{2+}$

L33 ANSWER 19 OF 36 HCAPLUS COPYRIGHT 2006 ACS on STN

ACCESSION NUMBER: 1994:12294 HCAPLUS

DOCUMENT NUMBER: 120:12294

TITLE: Recovery of platinum-group metals
from metal foil support catalysts

INVENTOR(S): Ito, Hidetoshi; Eto, Yoshuki

PATENT ASSIGNEE(S): Nissan Motor, Japan

SOURCE: Jpn. Kokai Tokkyo Koho, 3 pp.
CODEN: JKXXAF

DOCUMENT TYPE: Patent

LANGUAGE: Japanese

FAMILY ACC. NUM. COUNT: 1

PATENT INFORMATION:

PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
JP 05212296	A2	19930824	JP 1992-17614	19920203
PRIORITY APPLN. INFO.:			JP 1992-17614	19920203

AB The recovery involves (1) soaking the metal foil-supported catalysts in solns. containing NaOH \geq 20, NaHSO₄ 1-10, and Na boron hydride \geq 0.1 weight%, (2) dissolving the Al₂O₃- and CeO₂ coating layers (as catalyst active layer), and (3) recovering the Pt-group metals as platinum black. The method can recover the metals efficiently, and the method is applied to catalysts for such as exhaust gas purification

IT 1344-28-1, Alumina, uses

RL: USES (Uses)
(coat layers, catalysts containing, metal foil-support, recovery of platinum-group metals from)

RN 1344-28-1 HCAPLUS

CN Aluminum oxide (Al₂O₃) (8CI, 9CI) (CA INDEX NAME)

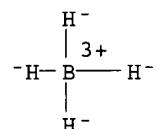
*** STRUCTURE DIAGRAM IS NOT AVAILABLE ***

IT 16940-66-2

RL: PROC (Process)
(dissoln. of alumina and ceria coat layers in solns. containing, in recovery recovery of platinum-group metals from metal foil-support catalysts)

RN 16940-66-2 HCAPLUS

CN Borate(1-), tetrahydro-, sodium (8CI, 9CI) (CA INDEX NAME)

● Na⁺IT 7440-05-3P, Palladium, preparation 7440-06-4P,
Platinum, preparation

RL: PUR (Purification or recovery); PREP (Preparation)
(recovery of, from metal foil-support catalysts containing alumina coat layers)

RN 7440-05-3 HCAPLUS

Valenrod 10_721479

CN Palladium (8CI, 9CI) (CA INDEX NAME)

Pd

RN 7440-06-4 HCAPLUS

CN Platinum (8CI, 9CI) (CA INDEX NAME)

Pt

Valenrod 10_721479

L33 ANSWER 20 OF 36 HCPLUS COPYRIGHT 2006 ACS on STN
ACCESSION NUMBER: 1992:637544 HCPLUS
DOCUMENT NUMBER: 117:237544
TITLE: Precious metal leaching with iodine-iodide
solutions
INVENTOR(S): Sloan, Hilbert
PATENT ASSIGNEE(S): Shapiro, Nelson H., USA
SOURCE: U.S., 11 pp. Cont. of U.S. Ser. No. 69,100, abandoned.
CODEN: USXXAM
DOCUMENT TYPE: Patent
LANGUAGE: English
FAMILY ACC. NUM. COUNT: 1
PATENT INFORMATION:

PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
US 5137700	A	19920811	US 1989-342587	19890424
PRIORITY APPLN. INFO.:			US 1987-69100	B1 19870702
AB	Precious metals (especially Au, Ag, Pd, and Pt) are recovered from electroplated scrap or articles by immersion into aqueous etching bath containing I 1, NH4I (or HI) 4, and water 10 parts. The precious metals are oxidized by I, form a soluble complex with the iodide, and can be precipitated from the solution by addition of a weak reducing agent (especially hydroxylamine or NaBH4) and recovered by filtration. The filtrate is oxidized with H2O2 or (NH4)S2O8 to precipitate elemental I for recycling to the etching stage.			
IT	7440-05-3P, Palladium, preparation 7440-06-4P, Platinum, preparation 7440-22-4P, Silver, preparation 7440-57-5P, Gold, preparation RL: PUR (Purification or recovery); PREP (Preparation) (recovery of, from electroplated scrap, leaching with iodine-iodide solution for)			
RN	7440-05-3 HCPLUS			
CN	Palladium (8CI, 9CI) (CA INDEX NAME)			

Pd

RN 7440-06-4 HCPLUS
CN Platinum (8CI, 9CI) (CA INDEX NAME)

Pt

RN 7440-22-4 HCPLUS
CN Silver (8CI, 9CI) (CA INDEX NAME)

Ag

RN 7440-57-5 HCPLUS
CN Gold (8CI, 9CI) (CA INDEX NAME)

Au

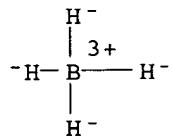
IT 16940-66-2, Sodium borohydride

RL: PROC (Process)

(reducing agent, precious metal recovery with,
after leaching with aqueous iodine-iodide solution)

RN 16940-66-2 HCPLUS

CN Borate(1-), tetrahydro-, sodium (8CI, 9CI) (CA INDEX NAME)



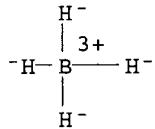
● Na^+

L33 ANSWER 21 OF 36 HCPLUS COPYRIGHT 2006 ACS on STN
 ACCESSION NUMBER: 1992:595687 HCPLUS
 DOCUMENT NUMBER: 117:195687
 TITLE: Recovery of palladium from tetraamminepalladium dichloride solutions by reduction
 INVENTOR(S): Suzuki, Masahito; Nagai, Makoto
 PATENT ASSIGNEE(S): N. E. Chemcat Corp., Japan
 SOURCE: Jpn. Kokai Tokkyo Koho, 3 pp.
 CODEN: JKXXAF
 DOCUMENT TYPE: Patent
 LANGUAGE: Japanese
 FAMILY ACC. NUM. COUNT: 1
 PATENT INFORMATION:

PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
JP 04165030	A2	19920610	JP 1990-287703	19901025
PRIORITY APPLN. INFO.:			JP 1990-287703	19901025
AB	[Pd(NH ₃) ₄]Cl ₂ solns. are added to solns. containing a strong alkaline reducer. Agglomeration-free Pd powders are prepared without loss.			
IT	7440-05-3P, Palladium, preparation			
	RL: PREP (Preparation)			
	(recovery of powdered, from tetraamminepalladium chloride, by reduction with alkaline reducer)			
RN	7440-05-3 HCPLUS			
CN	Palladium (8CI, 9CI) (CA INDEX NAME)			

Pd

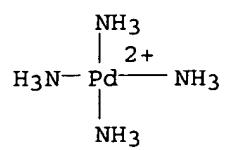
IT 16940-66-2, Sodium borohydride
 RL: RCT (Reactant); RACT (Reactant or reagent)
 (reduction by, of tetraamminepalladium chloride, for recovery of powdered palladium)
 RN 16940-66-2 HCPLUS
 CN Borate(1-), tetrahydro-, sodium (8CI, 9CI) (CA INDEX NAME)



● Na⁺

IT 13815-17-3
 RL: RCT (Reactant); RACT (Reactant or reagent)
 (reduction of, by alkaline reducer, for recovery of powdered palladium)
 RN 13815-17-3 HCPLUS
 CN Palladium(2+), tetraammine-, dichloride, (SP-4-1)- (9CI) (CA INDEX NAME)

Valenrod 10_721479



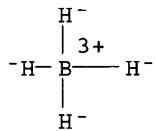
●2 Cl⁻

L33 ANSWER 22 OF 36 HCPLUS COPYRIGHT 2006 ACS on STN
 ACCESSION NUMBER: 1991:189660 HCPLUS
 DOCUMENT NUMBER: 114:189660
 TITLE: Reducing agent for recovery of precious **metals**
 from cyanide solutions
 INVENTOR(S): Shoda, Toshiaki
 PATENT ASSIGNEE(S): Nippon Electro Plating Engineers K. K., Japan
 SOURCE: Jpn. Kokai Tokkyo Koho, 3 pp.
 CODEN: JKXXAF
 DOCUMENT TYPE: Patent
 LANGUAGE: Japanese
 FAMILY ACC. NUM. COUNT: 1
 PATENT INFORMATION:

PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
JP 02163327	A2	19900622	JP 1988-316335	19881216
PRIORITY APPLN. INFO.:			JP 1988-316335	19881216
AB	The loaded cyanide feed solution (especially from leaching) is reduced with powdered			
	Al, powdered Zn, N ₂ H ₄ , alkali metal salt of HB, H ₂ PO ₃ , and/or dithionous acid, and/or a carbonyl compound to recover the precious metal . Thus, the aqueous solution (pH 13) containing 50 g KCN/L and 4.3 g Au/L was mixed with powdered Al 0.2, HCHO 1.0, and N ₂ H ₄ 0.1M, and showed the residual Au of 2.5 after 1 day and 1.6 mg/L after 5 days.			
IT	7440-57-5P, Gold, preparation RL: PUR (Purification or recovery); PREP (Preparation) (recovery of, from cyanide solution, reducing agents for)			
RN	7440-57-5 HCPLUS			
CN	Gold (8CI, 9CI) (CA INDEX NAME)			

Au

IT 16940-66-2 7429-90-5, Aluminum, uses and miscellaneous
 7440-66-6, Zinc, uses and miscellaneous
 RL: PROC (Process)
 (reducing agent, gold **recovery** with, from aqueous cyanide solution)
 RN 16940-66-2 HCPLUS
 CN Borate(1-), tetrahydro-, sodium (8CI, 9CI) (CA INDEX NAME)

● Na⁺

RN 7429-90-5 HCPLUS
 CN Aluminum (8CI, 9CI) (CA INDEX NAME)

Al

RN 7440-66-6 HCAPLUS
CN Zinc (7CI, 8CI, 9CI) (CA INDEX NAME)

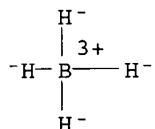
Zn

IT 7440-50-8, Copper, uses and miscellaneous
RL: USES (Uses)
(reducing agent, gold recovery with, from cyanide solution)
RN 7440-50-8 HCAPLUS
CN Copper (7CI, 8CI, 9CI) (CA INDEX NAME)

Cu

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L33 ANSWER 23 OF 36 HCPLUS COPYRIGHT 2006 ACS on STN
ACCESSION NUMBER: 1991:16564 HCPLUS
DOCUMENT NUMBER: 114:16564
TITLE: Separation of rhodium(III) from solutions, containing
iridium(IV) and -(III)
AUTHOR(S): Volkov, A. A.; Khain, V. S.
CORPORATE SOURCE: Ukhtinsk. Ind. Inst., Ukhta, USSR
SOURCE: Zhurnal Neorganicheskoi Khimii (1990), 35(9), 2220-3
CODEN: ZNOKAQ; ISSN: 0044-457X
DOCUMENT TYPE: Journal
LANGUAGE: Russian
AB A mixture of Ir(IV) as IrCl₆²⁻ or Ir(III) as IrCl₆³⁻ with Rh(III) as
Rh(OH)₆³⁻ was separated in strong acidic media (pH < 1) by NaBH₄. The method
can be used in refining Pt-group metals and for their
concentration with high yield and purity.
IT 16940-66-2, Sodium tetrahydroborate
RL: RCT (Reactant); RACT (Reactant or reagent)
(reduction by, of trivalent rhodium or tetravalent or trivalent iridium in
separation of rhodium from iridium)
RN 16940-66-2 HCPLUS
CN Borate(1-), tetrahydro-, sodium (8CI, 9CI) (CA INDEX NAME)



● Na^+

L33 ANSWER 24 OF 36 HCPLUS COPYRIGHT 2006 ACS on STN
 ACCESSION NUMBER: 1990:408958 HCPLUS
 DOCUMENT NUMBER: 113:8958
 TITLE: Chemical separation for platinum from rhodium
 INVENTOR(S): Zhang, Jian; Xu, Ying; Zhang, Ji
 PATENT ASSIGNEE(S): Xian Non-Ferrous Metals Institute, Peop. Rep. China
 SOURCE: Faming Zhuanli Shenqing Gongkai Shuomingshu, 5 pp.
 CODEN: CNXXEV
 DOCUMENT TYPE: Patent
 LANGUAGE: Chinese
 FAMILY ACC. NUM. COUNT: 1
 PATENT INFORMATION:

PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
CN 1034143	A	19890726	CN 1988-100131	19880111
CN 1011599	B	19910213		

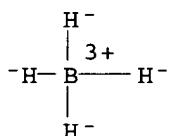
PRIORITY APPLN. INFO.: CN 1988-100131 19880111
 AB In electrolysis of Pt-Rh alloy, a small amount of H₂O₂ is added to the electrolyte solution for increasing the dissolving of Pt-Rh alloy electrodes, and the resulting electrolyte solution is then mixed with KCl to form Pt compound precipitate. The Pt compound precipitate is reduced with NaBH₄ to recover ≥99% pure Pt. Thus, waste Pt-Rh alloy chips were pressed to form electrodes, immersed into 8N HCl, placed in an electrolysis cell with magnetic stirrer, added 6 mL H₂O₂, and then electrolyzed at 180 mA/cm² for 3 h. The electrolyte solution containing 72.568 g/L Rh and 8.063 g/L Pt was mixed with 120% KCl, precipitated and filtered to obtain yellow precipitate, which was then washed with 10% KCl and 5% HCl solns., dried, and reduced with NaBH₄ to recover 99.5% pure Pt. The filtrate was passed through an ion-exchange resin bed to recover 99.9% pure Rh.

IT 16940-66-2

RL: USES (Uses)
 (reducing agent, for platinum recovery, in electrolysis of platinum-rhodium alloy)

RN 16940-66-2 HCPLUS

CN Borate(1-), tetrahydro-, sodium (8CI, 9CI) (CA INDEX NAME)



● Na⁺

IT 7440-06-4P, Platinum, preparation

RL: PREP (Preparation)
 (separation of, from platinum-rhodium alloy, by electrolysis, hydrogen peroxide addition in)

RN 7440-06-4 HCPLUS

CN Platinum (8CI, 9CI) (CA INDEX NAME)

Valenrod 10_721479

Pt

L33 ANSWER 25 OF 36 HCPLUS COPYRIGHT 2006 ACS on STN
 ACCESSION NUMBER: 1989:621064 HCPLUS
 DOCUMENT NUMBER: 111:221064
 TITLE: Recovering of noble **metal** from solution
 INVENTOR(S): Yamazaki, Hiroshi; Tateno, Yoshihisa
 PATENT ASSIGNEE(S): Tanaka Noble Metal Industrial Co., Ltd., Japan
 SOURCE: Jpn. Kokai Tokkyo Koho, 3 pp.
 CODEN: JKXXAF
 DOCUMENT TYPE: Patent
 LANGUAGE: Japanese
 FAMILY ACC. NUM. COUNT: 1
 PATENT INFORMATION:

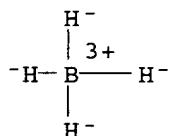
PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
JP 01083682	A2	19890329	JP 1987-241755	19870926
PRIORITY APPLN. INFO.:			JP 1987-241755	19870926
AB	A method for recovering a noble metal contained in a solution at least partially as a complex involves electrolysis to reduce the noble metal ions followed by recovery as well as to decompose the complex and recovering the noble metal ions formed by the decomposition using a reducing agent.			
IT	7440-05-3P, Palladium, preparation 7440-06-4P, Platinum , preparation RL: PREP (Preparation) (recovery of, by electrolysis and reducing agents)			
RN	7440-05-3 HCPLUS			
CN	Palladium (8CI, 9CI) (CA INDEX NAME)			

Pd

RN 7440-06-4 HCPLUS
 CN Platinum (8CI, 9CI) (CA INDEX NAME)

Pt

IT 16940-66-2, Sodium borohydride
 RL: PRP (Properties)
 (reducing agents, in **recovery** of noble metals)
 RN 16940-66-2 HCPLUS
 CN Borate(1-), tetrahydro-, sodium (8CI, 9CI) (CA INDEX NAME)

● Na^+

Valenrod 10_721479

L33 ANSWER 26 OF 36 HCAPLUS COPYRIGHT 2006 ACS on STN
 ACCESSION NUMBER: 1989:599088 HCAPLUS
 DOCUMENT NUMBER: 111:199088
 TITLE: Recovery of platinum-group metals
 from spent catalysts
 INVENTOR(S): Ezawa, Nobuyasu
 PATENT ASSIGNEE(S): Tanaka Noble Metal Industrial Co., Ltd., Japan
 SOURCE: Jpn. Kokai Tokkyo Koho, 5 pp.
 CODEN: JKXXAF
 DOCUMENT TYPE: Patent
 LANGUAGE: Japanese
 FAMILY ACC. NUM. COUNT: 1
 PATENT INFORMATION:

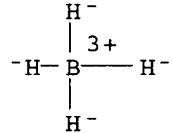
PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
JP 01108323	A2	19890425	JP 1987-264961	19871020
PRIORITY APPLN. INFO.:			JP 1987-264961	19871020

AB Spent catalysts of supported type containing Pt-group metals and/or their oxides are treated to recover the metal values. The process comprises: (1) reduction with NaBH4; (2) washing for Na removal; (3) chlorination leaching with aqueous HCl and sparged Cl; (4) leaching with aqueous HCl; (5) fluidized-bed treatment for cathodic recovery of Pt-group metals on C particles in an electrolysis cell; and (6) recirculating the spent solns. for addnl. leaching.

IT 16940-66-2, Sodium borohydride (NaBH4)
 RL: PROC (Process)
 (reduction with, of spent catalysts containing metal oxides in platinum-group metal recovery)

RN 16940-66-2 HCAPLUS

CN Borate(1-), tetrahydro-, sodium (8CI, 9CI) (CA INDEX NAME)



● Na⁺

L33 ANSWER 27 OF 36 HCAPLUS COPYRIGHT 2006 ACS on STN
 ACCESSION NUMBER: 1989:410611 HCAPLUS
 DOCUMENT NUMBER: 111:10611
 TITLE: Recovery of platinum-group metal
 from spent electrode
 INVENTOR(S): Sato, Hiroyasu; Maeda, Osamu
 PATENT ASSIGNEE(S): Tanaka Noble Metal Industrial Co., Ltd., Japan
 SOURCE: Jpn. Kokai Tokkyo Koho, 3 pp.
 CODEN: JKXXAF
 DOCUMENT TYPE: Patent
 LANGUAGE: Japanese
 FAMILY ACC. NUM. COUNT: 1
 PATENT INFORMATION:

PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
JP 01008228	A2	19890112	JP 1987-163182	19870630
JP 2575715	B2	19970129		

PRIORITY APPLN. INFO.: JP 1987-163182 19870630
 AB A spent electrode from NaCl electrolysis is treated to recover Pt
 -group metals by reduction of the PdO₂, PtO₂, and/or RhO₂ layer on
 the surface with a reducing agent (especially HCHO) followed by leaching.

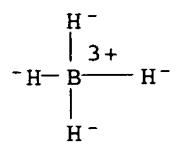
Thus,
 a spent electrode from Ti mesh coated with 6:4 (mol) Pt-PdO₂
 layer was reduced with aqueous 0.5% N₂H₂.2HCl, and then leached with 6:1
 HCl-HNO₃ mixture at 40-80°. The leached electrode was recoated with
 Pt and PdO₂, and then used in NaCl electrolysis.
 IT 7440-05-3P, Palladium, preparation 7440-06-4P,
 Platinum, preparation
 RL: PUR (Purification or recovery); PREP (Preparation)
 (recovery of, from spent electrode in salt electrolysis, reduction and
 leaching in)
 RN 7440-05-3 HCAPLUS
 CN Palladium (8CI, 9CI) (CA INDEX NAME)

Pd

RN 7440-06-4 HCAPLUS
 CN Platinum (8CI, 9CI) (CA INDEX NAME)

Pt

IT 16940-66-2
 RL: PROC (Process)
 (reduction with, for platinum-group metal
 recovery from spent salt-electrolysis electrode)
 RN 16940-66-2 HCAPLUS
 CN Borate(1-), tetrahydro-, sodium (8CI, 9CI) (CA INDEX NAME)



● Na^+

L33 ANSWER 28 OF 36 HCPLUS COPYRIGHT 2006 ACS on STN
 ACCESSION NUMBER: 1989:158128 HCPLUS
 DOCUMENT NUMBER: 110:158128
 TITLE: Recovery of platinum-group metals
 from spent electrodes
 INVENTOR(S): Sato, Hiroyasu; Maeda, Osamu
 PATENT ASSIGNEE(S): Tanaka Noble Metal Industrial Co., Ltd., Japan
 SOURCE: Jpn. Kokai Tokkyo Koho, 3 pp.
 CODEN: JKXXAF
 DOCUMENT TYPE: Patent
 LANGUAGE: Japanese
 FAMILY ACC. NUM. COUNT: 1
 PATENT INFORMATION:

PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
JP 63270421	A2	19881108	JP 1987-104133	19870427
PRIORITY APPLN. INFO.:			JP 1987-104133	19870427

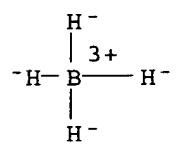
AB Spent NaCl-electrolysis electrodes coated with a Pt-group metal oxide are treated to recover Pt, Ir, and Ru. The oxide coating is reduced, leached with aqueous solution containing HCl and HNO₃ to recover a part of Pt and Ir, and then electrolyzed in a fluoride bath for coating separation. Suitable reducing agents are N₂H₄, NaBH₄, and sulfites, and the fluoride is HF and/or NH₄HF₂. Thus, spent Ti electrodes had a coating with 88:12:5 mol mixture of Pt oxide, Ir₂O₃, and Ru oxide. The electrodes were dipped for 1 h in aqueous 0.2% N₂H₄·2HCl, and then for 1.5 h in 6:1 HCl:HNO₃ mixture at 60-70°. The treated electrodes as anode were electrolyzed at 0.2 A/dm² and 4.2 V for 10 min in aqueous bath containing 5 weight% each of HF and NH₄HF₂. The Pt, Ir, and Ru were recovered.

IT 7440-06-4P, Platinum, preparation
 RL: PUR (Purification or recovery); PREP (Preparation)
 (recovery of, from spent electrolysis electrodes, reduction of oxide
 coating for electrochem.)
 RN 7440-06-4 HCPLUS
 CN Platinum (8CI, 9CI) (CA INDEX NAME)

Pt

IT 16940-66-2
 RL: PROC (Process)
 (reduction with, of metal oxide coating on spent electrode,
 recovery of platinum-group metals after)
 RN 16940-66-2 HCPLUS
 CN Borate(1-), tetrahydro-, sodium (8CI, 9CI) (CA INDEX NAME)

Valenrod 10_721479



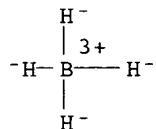
● Na^+

L33 ANSWER 29 OF 36 HCPLUS COPYRIGHT 2006 ACS on STN
 ACCESSION NUMBER: 1989:158127 HCPLUS
 DOCUMENT NUMBER: 110:158127
 TITLE: Recovery of platinum-group metals
 from spent electrodes
 INVENTOR(S): Sato, Hiroyasu; Maeda, Osamu
 PATENT ASSIGNEE(S): Tanaka Noble Metal Industrial Co., Ltd., Japan
 SOURCE: Jpn. Kokai Tokkyo Koho, 3 pp.
 CODEN: JKXXAF
 DOCUMENT TYPE: Patent
 LANGUAGE: Japanese
 FAMILY ACC. NUM. COUNT: 1
 PATENT INFORMATION:

PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
JP 63270420	A2	19881108	JP 1987-104132	19870427
PRIORITY APPLN. INFO.:			JP 1987-104132	19870427
AB	Spent NaCl-electrolysis electrodes having a Pt-group metal oxide coating are treated to recover Pt, In, and Ru. The oxide coating is reduced, leached with aqueous solution containing HCl and HNO ₃ to recover a part of Pt and In, and then contacted with a mixture of HNO ₃ , H ₂ SO ₄ , and HF for dissoln. Suitable reducing agents are N ₂ H ₄ , N ₂ H ₄ , NaBH ₄ , and sulfites. Thus, spent Ti electrodes had a coating with 88:12:5 mol mixture of Pt oxide, In ₂ O ₃ , and Ru oxide. The electrodes were treated for 1 h in aqueous 0.2% N ₂ H ₄ .2HCl, for 1.5 h in 6:1 HCl:HNO ₃ solution at 60-70°, and then for 10 min at 50-60° in a solution containing HNO ₃ 2, H ₂ SO ₄ 2, HF 1, and water 5 volume parts to recover Pt, Ir, and Ru.			
IT	7440-06-4P, Platinum, preparation			
	RL: PUR (Purification or recovery); PREP (Preparation) (recovery of, from spent electrodes, reduction of oxide coatings for leaching in)			
RN	7440-06-4 HCPLUS			
CN	Platinum (8CI, 9CI) (CA INDEX NAME)			

Pt

IT 16940-66-2
 RL: PROC (Process)
 (reduction with, of oxide coating on spent electrode, recovery of platinum-group metals after)
 RN 16940-66-2 HCPLUS
 CN Borate(1-), tetrahydro-, sodium (8CI, 9CI) (CA INDEX NAME)

● Na⁺

Valenrod 10_721479

L33 ANSWER 30 OF 36 HCAPLUS COPYRIGHT 2006 ACS on STN
 ACCESSION NUMBER: 1986:37545 HCAPLUS
 DOCUMENT NUMBER: 104:37545
 TITLE: Winning precious **metals** from geological
 brines
 INVENTOR(S): Luebbe, Ray H., Jr.; Wemhoff, Mark
 PATENT ASSIGNEE(S): Environmental Technology (U.S.), Inc., USA
 SOURCE: Eur. Pat. Appl., 22 pp.
 CODEN: EPXXDW
 DOCUMENT TYPE: Patent
 LANGUAGE: English
 FAMILY ACC. NUM. COUNT: 1
 PATENT INFORMATION:

PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
EP 159842	A2	19851030	EP 1985-302382	19850404
R: DE, FR, GB				
PRIORITY APPLN. INFO.:			US 1984-601792	A 19840419
			US 1984-610385	A 19840515

AB Geothermal and oil-field brines are contacted with an inorg. glassy material for selective adsorption of precious **metals**, especially after surface pretreatment by reduction. The **metal**-loaded glass is leached (e.g. with aqua regia) and recycled. Hydrophobic organic plastics and clays (especially attapulgite) are used as adsorbents. The brines are optionally pretreated with an oxidizing agent to enhance recovery of precious **metals** at less than +200 mV (redox) to avoid sorption of S and base elements. Thus, oil-field brine 5 gal was circulated at 2.4 L/min for 16 h through a glass chromatog. column packed with glass beads of 4 mm diameter. The beads were washed with deionized water 125 mL of pH 9.5, and 5 mL of cyanogen iodide solution was added. Recovery of Ag was only 0.03 mg/L, but was increased to 0.3 mg/L by extraction with aqua regia in 1:1 solution
 IT 1309-48-4, uses and miscellaneous 1344-28-1, uses and
 miscellaneous
 RL: USES (Uses)
 (activated, for adsorption of precious **metals** from brines)
 RN 1309-48-4 HCAPLUS
 CN Magnesium oxide (MgO) (9CI) (CA INDEX NAME)

Mg—O

RN 1344-28-1 HCAPLUS
 CN Aluminum oxide (Al2O3) (8CI, 9CI) (CA INDEX NAME)

*** STRUCTURE DIAGRAM IS NOT AVAILABLE ***
 IT 7440-05-3P, preparation 7440-06-4P, preparation
 7440-22-4P, preparation 7440-57-5P, preparation
 RL: PUR (Purification or recovery); PREP (Preparation)
 (recovery of, from brines, glassy sorbents for)
 RN 7440-05-3 HCAPLUS
 CN Palladium (8CI, 9CI) (CA INDEX NAME)

Pd

RN 7440-06-4 HCAPLUS
 CN Platinum (8CI, 9CI) (CA INDEX NAME)

Pt

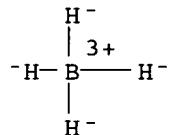
RN 7440-22-4 HCAPLUS
CN Silver (8CI, 9CI) (CA INDEX NAME)

Ag

RN 7440-57-5 HCAPLUS
CN Gold (8CI, 9CI) (CA INDEX NAME)

Au

IT **16971-29-2P**
RL: PREP (Preparation)
(reduction with, of glassy sorbent for precious metal
recovery from brines)
RN 16971-29-2 HCAPLUS
CN Borate(1-), tetrahydro- (8CI, 9CI) (CA INDEX NAME)



Valenrod 10_721479

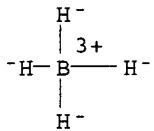
L33 ANSWER 31 OF 36 HCPLUS COPYRIGHT 2006 ACS on STN
ACCESSION NUMBER: 1985:508225 HCPLUS
DOCUMENT NUMBER: 103:108225
TITLE: Recovery of platinum metals
especially ruthenium
PATENT ASSIGNEE(S): Daido Steel Co., Ltd., Japan
SOURCE: Jpn. Kokai Tokkyo Koho, 3 pp.
CODEN: JKXXAF
DOCUMENT TYPE: Patent
LANGUAGE: Japanese
FAMILY ACC. NUM. COUNT: 1
PATENT INFORMATION:

PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
JP 60092433	A2	19850524	JP 1983-199154	19831026
PRIORITY APPLN. INFO.:			JP 1983-199154	19831026

AB Spent electrodes or catalysts are fused with KOH + KNO₃ (or K₂S₂O₇), stirred in water, and filtered. The filtrate is adjusted to pH 5-6, and filtered. The precipitate is dissolved in dilute HCl, filtered, and the filtrate reduced with NaBH₄ to obtain the Pt-group metals.
Thus, spent electrodes 15 g were processed with KOH + KNO₃, to recover 23 mg metal containing Ru 87 and Ir 3.8%. The residue was heated at 800° to recover 56 mg sublimate containing 70% RuO₂.

IT 16940-66-2
RL: PROC (Process)
(reduction with, in processing of spent catalysts and electrodes for platinum-group metal recovery)

RN 16940-66-2 HCPLUS
CN Borate(1-), tetrahydro-, sodium (8CI, 9CI) (CA INDEX NAME)



● Na⁺

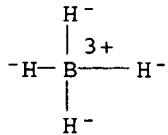
L33 ANSWER 32 OF 36 HCPLUS COPYRIGHT 2006 ACS on STN
 ACCESSION NUMBER: 1983:581262 HCPLUS
 DOCUMENT NUMBER: 99:181262
 TITLE: Oxygen scavenging
 PATENT ASSIGNEE(S): Kurita Kogyo Co., Ltd., Japan
 SOURCE: Jpn. Kokai Tokkyo Koho, 3 pp.
 CODEN: JKXXAF
 DOCUMENT TYPE: Patent
 LANGUAGE: Japanese
 FAMILY ACC. NUM. COUNT: 1
 PATENT INFORMATION:

PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
JP 58079590	A2	19830513	JP 1981-177765	19811105
JP 59032195	B4	19840807		

PRIORITY APPLN. INFO.: JP 1981-177765 19811105
 AB An O-containing water is treated with a Pd catalyst in the presence of NaBH4 and/or N2H4 (or its hydrate). The Pd catalyst is loaded on an anion exchanger. Thus, a Pd chloride was treated with Lewatit OC 1045 [78769-27-4] to prepare a Pd catalyst. Boiler water was mixed with NaBH4 and then passed through a column containing the Pd-anion exchanger. The dissolved O in the eluent was 0.1 mg/L after 200 L water/h-L resin.
 IT 7440-05-3, uses and miscellaneous
 RL: CAT (Catalyst use); USES (Uses)
 (catalyst, for oxygen scavenging from boiler water, Lewatit OC 1045 support for)
 RN 7440-05-3 HCPLUS
 CN Palladium (8CI, 9CI) (CA INDEX NAME)

Pd

IT 16940-66-2
 RL: OCCU (Occurrence)
 (in oxygen scavenging in boiler water)
 RN 16940-66-2 HCPLUS
 CN Borate(1-), tetrahydro-, sodium (8CI, 9CI) (CA INDEX NAME)



● Na⁺

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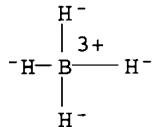
L33 ANSWER 33 OF 36 HCPLUS COPYRIGHT 2006 ACS on STN
ACCESSION NUMBER: 1983:164610 HCPLUS
DOCUMENT NUMBER: 98:164610
TITLE: The treatment of spent electroless nickel baths by autocatalytic reduction
AUTHOR(S): Parker, Konrad
CORPORATE SOURCE: Park Ridge, IL, 60068, USA
SOURCE: Plating and Surface Finishing (1983), 70(3), 60-2
CODEN: PSFMDH; ISSN: 0360-3164
DOCUMENT TYPE: Journal
LANGUAGE: English
AB The autocatalytic decomposition of some spent electroless Ni baths using NaBH4 and/or Pd ions can reduce the Ni concentration to ≤ 10 mg/L. The chemical of the plating solution appears to be a controlling factor.
IT 7440-02-0P, preparation
RL: PUR (Purification or recovery); PREP (Preparation)
(recovery of, from spent electroless bath, by autocatalytic reduction)
RN 7440-02-0 HCPLUS
CN Nickel (8CI, 9CI) (CA INDEX NAME)

Ni

IT 7647-10-1 16940-66-2
RL: RCT (Reactant); RACT (Reactant or reagent)
(reduction by, of spent electroless bath, for nickel recovery)
RN 7647-10-1 HCPLUS
CN Palladium chloride (PdCl2) (6CI, 8CI, 9CI) (CA INDEX NAME)

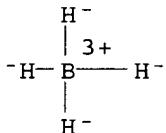
Cl-Pd-Cl

RN 16940-66-2 HCPLUS
CN Borate(1-), tetrahydro-, sodium (8CI, 9CI) (CA INDEX NAME)



● Na⁺

L33 ANSWER 34 OF 36 HCAPLUS COPYRIGHT 2006 ACS on STN
 ACCESSION NUMBER: 1982:620285 HCAPLUS
 DOCUMENT NUMBER: 97:220285
 TITLE: Applications for sodium borohydride
 in precious metal recovery and
 recycle
 AUTHOR(S): Medding, G. L.; Lander, J. A.
 CORPORATE SOURCE: Thiokol/Ventron Div., Danvers, MA, 01923, USA
 SOURCE: Precious Met. [Proc. Int. Precious Met. Inst. Conf.],
 5th (1982), Meeting Date 1981, 3-10. Editor(s): Zysk,
 Edward D. Pergamon: Willowdale, Ont.
 CODEN: 48RYAT
 DOCUMENT TYPE: Conference
 LANGUAGE: English
 AB NaBH4 recovery operations are used in precious metal recovery
 applications. The most established precious metal recovery use
 is Ag removal from spent photog. fixing solns. Ag recovery of >99% at
 >95% purity can be achieved with this process. Applications have been
 developed for reduction of AgCl via a process which is capable of achieving
 extremely pure Ag with >99% recovery of Ag from scrap. Applications are
 being developed in the areas of Pt, Pd, and Rh
 recovery from spent catalyst materials. NaBH4 is being used com. for
 recovery of most of the precious metals.
 IT 16940-66-2
 RL: PROC (Process)
 (in precious metal recovery)
 RN 16940-66-2 HCAPLUS
 CN Borate(1-), tetrahydro-, sodium (8CI, 9CI) (CA INDEX NAME)



● Na⁺

IT 7440-06-4P, preparation 7440-22-4P, preparation
 RL: PUR (Purification or recovery); PREP (Preparation)
 (recovery of, sodium borohydride in)
 RN 7440-06-4 HCAPLUS
 CN Platinum (8CI, 9CI) (CA INDEX NAME)

Pt

RN 7440-22-4 HCAPLUS
 CN Silver (8CI, 9CI) (CA INDEX NAME)

Ag

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L33 ANSWER 35 OF 36 HCPLUS COPYRIGHT 2006 ACS on STN
ACCESSION NUMBER: 1982:168042 HCPLUS
DOCUMENT NUMBER: 96:168042
TITLE: Sodium borohydride controls heavy metal
discharge
AUTHOR(S): Cook, Michael M.; Lander, Joseph A.
CORPORATE SOURCE: Ventron Div., Thiokol Corp., Danvers, MA, USA
SOURCE: Pollution Engineering (1981), 13(12), 36-8
CODEN: PLENBW; ISSN: 0032-3640
DOCUMENT TYPE: Journal
LANGUAGE: English
AB Discharge of heavy metals into sewer systems and open waters is
a major concern to both publicly owned treatment works and industries.
Using NaBH4 as an alternative chemical treatment has enabled a number of plants
to meet wastewater discharge limits in a cost-effective manner. It
provides an effective and economical method for the removal of pollutants
or for the recovery of precious metals, including Co, Cu, Au,
Ir, Pb, Hg, Ni, Pd, Pt, Rh, and Hg. The metal
is chemical reduced to the elemental state, usually as a compact precipitate

The NaBH4 process is irreversible, extremely rapid, and results in nearly
total removal of the dissolved metals even at low initial concentration

IT 7439-92-1P, preparation 7440-02-0P, preparation
7440-05-3P, preparation 7440-06-4P, preparation
7440-22-4P, preparation 7440-48-4P, preparation
7440-50-8P, preparation 7440-57-5P, preparation
RL: PREP (Preparation)
(recovery of, from wastewater, precipitation by sodium
borohydride in)

RN 7439-92-1 HCPLUS
CN Lead (8CI, 9CI) (CA INDEX NAME)

Pb

RN 7440-02-0 HCPLUS
CN Nickel (8CI, 9CI) (CA INDEX NAME)

Ni

RN 7440-05-3 HCPLUS
CN Palladium (8CI, 9CI) (CA INDEX NAME)

Pd

RN 7440-06-4 HCPLUS
CN Platinum (8CI, 9CI) (CA INDEX NAME)

Pt

RN 7440-22-4 HCPLUS
CN Silver (8CI, 9CI) (CA INDEX NAME)

Valenrod 10_721479

Ag

RN 7440-48-4 HCAPLUS
CN Cobalt (8CI, 9CI) (CA INDEX NAME)

Co

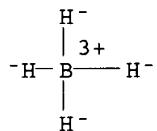
RN 7440-50-8 HCAPLUS
CN Copper (7CI, 8CI, 9CI) (CA INDEX NAME)

Cu

RN 7440-57-5 HCAPLUS
CN Gold (8CI, 9CI) (CA INDEX NAME)

Au

IT 16940-66-2
RL: PROC (Process)
(wastewater treatment by, metal precipitation in)
RN 16940-66-2 HCAPLUS
CN Borate(1-), tetrahydro-, sodium (8CI, 9CI) (CA INDEX NAME)



● Na⁺

L33 ANSWER 36 OF 36 HCPLUS COPYRIGHT 2006 ACS on STN
 ACCESSION NUMBER: 1975:485606 HCPLUS
 DOCUMENT NUMBER: 83:85606
 TITLE: Aromatic hydrogenation using sodium borohydride
 reduced transition metal supported catalysts
 INVENTOR(S): Benson, Robert F.
 PATENT ASSIGNEE(S): Texaco Inc., USA
 SOURCE: U.S., 5 pp.
 CODEN: USXXAM
 DOCUMENT TYPE: Patent
 LANGUAGE: English
 FAMILY ACC. NUM. COUNT: 1
 PATENT INFORMATION:

PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
US 3869521	A	19750304	US 1972-319574	19721229
PRIORITY APPLN. INFO.:			US 1971-158581	A1 19710630
AB Supported transition metal catalysts are provided by a method comprising impregnating a catalyst support with a solution comprising a catalyst metal ion; at least partially drying the saturated catalyst support; admixing the dried catalyst support with sodium borohydride aqueous solution until the evolution of gas ceases, thereby obtaining an active supported metal catalyst; and thereafter, recovering the active catalyst.				
IT	7440-02-0, uses and miscellaneous			
	RL: CAT (Catalyst use); USES (Uses)			
	(catalysts, for demethylation of methylnaphthalene, manufacture of)			
RN	7440-02-0 HCPLUS			
CN	Nickel (8CI, 9CI) (CA INDEX NAME)			

Ni

IT 7440-06-4, uses and miscellaneous
 RL: CAT (Catalyst use); USES (Uses)
 (catalysts, for hydrogenation of benzene, manufacture of)
 RN 7440-06-4 HCPLUS
 CN Platinum (8CI, 9CI) (CA INDEX NAME)

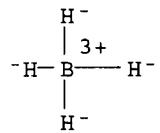
Pt

IT 7440-05-3, uses and miscellaneous
 RL: CAT (Catalyst use); USES (Uses)
 (catalysts, for hydrogenation of toluene, manufacture of)
 RN 7440-05-3 HCPLUS
 CN Palladium (8CI, 9CI) (CA INDEX NAME)

Pd

IT 16940-66-2
 RL: RCT (Reactant); RACT (Reactant or reagent)
 (reduction by, in preparation of hydrogenation catalysts)
 RN 16940-66-2 HCPLUS
 CN Borate(1-), tetrahydro-, sodium (8CI, 9CI) (CA INDEX NAME)

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● Na^+

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=> => d stat que 153
L1      795 SEA FILE=REGISTRY ABB=ON  PLU=ON  TETRAHYDROBORATE OR HYDROBORA
              TE
L2      11 SEA FILE=REGISTRY ABB=ON  PLU=ON  SODIUM BOROHYDRATE?/CN OR
              LITHIUM BOROHYDR?/CN OR POTASSIUM BOROHYDR?/CN
L3      104214 SEA FILE=REGISTRY ABB=ON  PLU=ON  METAL OR METALS OR MAGNESIUM
L4      128483 SEA FILE=REGISTRY ABB=ON  PLU=ON  BORATE
L5      15523 SEA FILE=HCAPLUS ABB=ON  PLU=ON  L1 OR TETRAHYDROBORATE OR
              HYDROBORATE
L6      SEL PLU=ON L2 1- CHEM :      52 TERMS
L7      15422 SEA FILE=HCAPLUS ABB=ON  PLU=ON  L6
L8      15583 SEA FILE=HCAPLUS ABB=ON  PLU=ON  L7 OR (SODIUM OR LITHIUM OR
              POTASSIUM) (2A) BOROHYDR?
L9      4137887 SEA FILE=HCAPLUS ABB=ON  PLU=ON  L3 OR METAL OR MAGNESIUM
L10     166954 SEA FILE=HCAPLUS ABB=ON  PLU=ON  L4 OR BORATE
L12     15254 SEA FILE=HCAPLUS ABB=ON  PLU=ON  L10 (L) L9
L14     38473 SEA FILE=HCAPLUS ABB=ON  PLU=ON  L10 (L) (RCT/RL OR RACT/RL)
L15     250013 SEA FILE=HCAPLUS ABB=ON  PLU=ON  L9 (L) (RCT/RL OR RACT/RL)
L16     2579 SEA FILE=HCAPLUS ABB=ON  PLU=ON  L12 AND L14
L17     1952 SEA FILE=HCAPLUS ABB=ON  PLU=ON  L16 AND L15
L18     104576 SEA FILE=REGISTRY ABB=ON  PLU=ON  PALLADIUM/BI
L19     208 SEA FILE=REGISTRY ABB=ON  PLU=ON  PT/MF
L21     513099 SEA FILE=HCAPLUS ABB=ON  PLU=ON  L18 OR L19 OR PALLADIUM OR
              PLATINUM OR PD OR PT
L25     3041 SEA FILE=HCAPLUS ABB=ON  PLU=ON  (L5 OR L8) (L) (PREP?/RL OR
              PREP OR PREPARTION)
L26     78 SEA FILE=HCAPLUS ABB=ON  PLU=ON  L17 AND L25
L27     11 SEA FILE=HCAPLUS ABB=ON  PLU=ON  L26 AND L21
L28     9 SEA FILE=HCAPLUS ABB=ON  PLU=ON  L27 AND PD=<NOVEMBER 25, 2003
L30     524 SEA FILE=HCAPLUS ABB=ON  PLU=ON  (L5 OR L8) (L) (PURIFI?/RL OR
              PURIFI? OR RECOVER?)
L31     152 SEA FILE=HCAPLUS ABB=ON  PLU=ON  L30 AND L10 AND L9
L32     38 SEA FILE=HCAPLUS ABB=ON  PLU=ON  L31 AND L9 AND L21
L33     36 SEA FILE=HCAPLUS ABB=ON  PLU=ON  L32 NOT (L27 OR L28)
L44     294 SEA FILE=HCAPLUS ABB=ON  PLU=ON  "SUDA S"/AU OR "SUDA SEIJIRAU"
              /AU
L45     1522 SEA FILE=HCAPLUS ABB=ON  PLU=ON  LI Z ?/AU OR "LI ZHOUPEENG"/AU
L46     52 SEA FILE=HCAPLUS ABB=ON  PLU=ON  "IWASE Y"/AU OR ("IWASE
              YASUTOSHI"/AU OR "IWASE YASUYOSHI"/AU)
L47     8 SEA FILE=HCAPLUS ABB=ON  PLU=ON  ("MORIGASAKI N"/AU OR
              "MORIGASAKI NOBUTO"/AU)
L48     7 SEA FILE=HCAPLUS ABB=ON  PLU=ON  L47 NOT (L33 OR L27 OR L28)
L49     5 SEA FILE=HCAPLUS ABB=ON  PLU=ON  L44 AND L45 AND L46
L50     44 SEA FILE=HCAPLUS ABB=ON  PLU=ON  L44 AND (L45 OR L46)
L51     5 SEA FILE=HCAPLUS ABB=ON  PLU=ON  L45 AND L46
L52     46 SEA FILE=HCAPLUS ABB=ON  PLU=ON  L48 OR L49 OR L50 OR L51
L53     46 SEA FILE=HCAPLUS ABB=ON  PLU=ON  L52 NOT (L33 OR L27 OR L28)
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L53 ANSWER 1 OF 46 HCAPLUS COPYRIGHT 2006 ACS on STN
ACCESSION NUMBER: 2005:1249878 HCAPLUS
DOCUMENT NUMBER: 144:276900
TITLE: Development of the direct borohydride fuel cell
AUTHOR(S): Li, Z. P.; Liu, B. H.; Arai, K.; Suda, S.
CORPORATE SOURCE: Materials & Energy Research Institute Tokyo, Ltd.,
Chino-shi, Nagano, 391-0301, Japan
SOURCE: Journal of Alloys and Compounds (2005), 404-406,
648-652
CODEN: JALCEU; ISSN: 0925-8388
PUBLISHER: Elsevier B.V.
DOCUMENT TYPE: Journal; General Review
LANGUAGE: English
AB A review on development of the direct borohydride fuel cells. The theor. energy conversion efficiency of the direct borohydride fuel cell (0.91) is larger than that of the polymer electrolyte membrane fuel cell (0.83). Recently, the cell power d. 290 mW/cm² has been achieved. A five-cell stack with an effective area of 67 cm² has demonstrated power of 110 W when the operation temperature is 60°, even if the stack operation started at room temperature without humidification.
REFERENCE COUNT: 22 THERE ARE 22 CITED REFERENCES AVAILABLE FOR THIS RECORD. ALL CITATIONS AVAILABLE IN THE RE FORMAT

L53 ANSWER 2 OF 46 HCAPLUS COPYRIGHT 2006 ACS on STN
ACCESSION NUMBER: 2005:1249877 HCAPLUS
TITLE: Production of sodium borohydride by using dynamic
behaviors of protide at the extreme surface of
magnesium particles
AUTHOR(S): Suda, S.; Morigasaki, N.;
Iwase, Y.; Li, Z. P.
CORPORATE SOURCE: Department of Environmental and Chemical Engineering,
Chemical Energy Laboratory, Kogakuin University,
2665-1, Nakano-machi, Hachioji-shi, Tokyo, 192-0015,
Japan
SOURCE: Journal of Alloys and Compounds (2005), 404-406,
643-647
CODEN: JALCEU; ISSN: 0925-8388
PUBLISHER: Elsevier B.V.
DOCUMENT TYPE: Journal
LANGUAGE: English
AB An advanced process for the production of sodium borohydride (NaBH4) as a
hydrogen storage material was developed, which applied the dynamic
hydriding and dehydriding behaviors of protide (H-) in Mg-H system under
transitional temperature conditions. An abundant natural resource named borax
(Na2B4O7·10H2O) and the anhydrous sodium metaborate (NaBO2) recovered
from the "spent fuel" as NaBO2·4H2O were used as the starting material in
the present process. Powder-state Mg played an important role in the
transitional hydriding and dehydriding process where the gaseous hydrogen
was converted to protide at the extreme surface of Mg to form NaBH4 in
exchange with the simultaneous transition of oxygen in NaBO2 to form MgO.
In the present process, the protide as the most reactive state among the
four states of hydrogen is applied for the synthesis of NaBH4, which can
exist in metal-hydrogen complexes, such as NaAlH4 and NaBH4. The NaBH4
yield was reached higher than 90% by a single batch process but was found
to be largely dependent on the rate of temperature change and the particle
size,
i.e., the sp. surface area of Mg particles.
REFERENCE COUNT: 31 THERE ARE 31 CITED REFERENCES AVAILABLE FOR THIS
RECORD. ALL CITATIONS AVAILABLE IN THE RE FORMAT

L53 ANSWER 3 OF 46 HCAPLUS COPYRIGHT 2006 ACS on STN

ACCESSION NUMBER: 2005:491812 HCAPLUS

DOCUMENT NUMBER: 143:176139

TITLE: Performance improvement of a micro borohydride fuel cell operating at ambient conditions

AUTHOR(S): Liu, B. H.; Li, Z. P.; Arai, K.; Suda, S.

CORPORATE SOURCE: Department of Environmental and Chemical Engineering, Kogakuin University, Hachioji, Tokyo, 192-0015, Japan

SOURCE: Electrochimica Acta (2005), 50(18), 3719-3725
CODEN: ELCAAV; ISSN: 0013-4686

PUBLISHER: Elsevier B.V.

DOCUMENT TYPE: Journal

LANGUAGE: English

AB In this study, aqueous borohydride solns. were employed to fuel a micro cell. Electrochem. performance of the micro borohydride fuel cell was tested at ambient conditions without any auxiliary facilities. Electrochem. impedance spectroscopy analyses were performed to characterize the cell performance. Both anion and cation exchange membranes were tried to sep. the fuel from the cathode. Membrane properties were found to be a decisive factor for cell performance. A maximum power d. of 40 mW/cm² at room temperature was achieved when the Nafion NRE211 membrane was used. Hydrogen evolution at the anode side resulting from the competitive hydrolysis reaction influenced cell performance by preventing transfer of the electrolyte. The cell also demonstrated promising performance even when a silver cathode was used.

REFERENCE COUNT: 14 THERE ARE 14 CITED REFERENCES AVAILABLE FOR THIS RECORD. ALL CITATIONS AVAILABLE IN THE RE FORMAT

L53 ANSWER 4 OF 46 HCPLUS COPYRIGHT 2006 ACS on STN
 ACCESSION NUMBER: 2004:1060839 HCPLUS
 DOCUMENT NUMBER: 142:25940
 TITLE: Production of alkali metal tetrahydroborates by
 hydrogenation of alkali metal metaborates via alkaline
 earth metal hydride intermediates
 INVENTOR(S): Suda, Seijirau; Li, Zhoupeng;
 Iwase, Yasuyoshi; Morigasaki, Nobuto
 PATENT ASSIGNEE(S): Japan
 SOURCE: U.S. Pat. Appl. Publ., 9 pp., Cont.-in-part of U.S.
 Ser. No. 558,828.
 CODEN: USXXCO
 DOCUMENT TYPE: Patent
 LANGUAGE: English
 FAMILY ACC. NUM. COUNT: 2
 PATENT INFORMATION:

PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
US 2004249215	A1	20041209	US 2003-721479	20031125
JP 2004224684	A2	20040812	JP 2003-72466	20030317
PRIORITY APPLN. INFO.:			US 2000-558828	A2 20000426
			JP 2002-344426	A 20021127
			JP 2003-72466	A 20030317

AB Alkali metal tetrahydroborates are produced in high yield by reaction of a suitable alkali metal borate (preferably an alkali metal metaborate) with a finely ground alkaline earth metal (preferably magnesium, ground to average particle diameter $\leq 100 \mu\text{m}$) in the presence of H₂ at below the equilibrium pressure for stable alkaline earth metal hydride formation. The solid reactants are mixed with hydrogen at $\leq 450^\circ$, and heated to 500-650 $^\circ$. Coke-oven gas is a suitable hydrogen source. Alkali metal tetrahydroborates have been suggested as a convenient source of fuel-cell-grade hydrogen.

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L53 ANSWER 5 OF 46 HCAPLUS COPYRIGHT 2006 ACS on STN
ACCESSION NUMBER: 2004:899747 HCAPLUS
DOCUMENT NUMBER: 143:175994
TITLE: Sodium hydroborate as a hydrogen storing material
AUTHOR(S): Suda, Seijirou; **Morigasaki, Nobuto**; Iwase,
Yasuyoshi; Li. Zhou Peng
CORPORATE SOURCE: Dep. of Engineering, Kogakuin University, Japan
SOURCE: Kagaku Sochi (2004), 46(10), 64-71
CODEN: KASOB7; ISSN: 0368-4849
PUBLISHER: Kogyo Chosakai
DOCUMENT TYPE: Journal; General Review
LANGUAGE: Japanese
AB A review with 0 refs.

L53 ANSWER 6 OF 46 HCAPLUS COPYRIGHT 2006 ACS on STN
 ACCESSION NUMBER: 2004:446908 HCAPLUS
 DOCUMENT NUMBER: 140:409212
 TITLE: Method for producing tetrahydroborates
 INVENTOR(S): Suda, Seijirau; Li, Zhoupeng;
 Iwase, Yasuyoshi; Morigasaki, Nobuto
 PATENT ASSIGNEE(S): Materials & Energy Research Institute Tokyo, Ltd.,
 Japan
 SOURCE: Eur. Pat. Appl., 11 pp.
 CODEN: EPXXDW
 DOCUMENT TYPE: Patent
 LANGUAGE: English
 FAMILY ACC. NUM. COUNT: 2
 PATENT INFORMATION:

PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
EP 1424310	A2	20040602	EP 2003-27149	20031126
R: AT, BE, CH, DE, DK, ES, FR, GB, GR, IT, LI, LU, NL, SE, MC, PT, IE, SI, LT, LV, FI, RO, MK, CY, AL, TR, BG, CZ, EE, HU, SK				
JP 2004224684	A2	20040812	JP 2003-72466	20030317
JP 2002-344426 A 20021127				
JP 2003-72466 A 20030317				

PRIORITY APPLN. INFO.:
 AB Tetrahydroborates, such as NaBH₄, LiBH₄, and KBH₄, are prepared by reacting a powdery mixture containing a borate and magnesium in a hydrogen atmospheric under pressure at 500-650°. The hydrogen source can be coke oven gas. The reaction can be carried out in the presence of a hydrogenation catalyst. The starting materials have an average particle diameter of < 100 μ m.

Valenrod 10_721479

L53 ANSWER 7 OF 46 HCPLUS COPYRIGHT 2006 ACS on STN
ACCESSION NUMBER: 2004:306656 HCPLUS
DOCUMENT NUMBER: 141:210022
TITLE: Metal-hydrogen complex compounds as hydrogen storage
materials
AUTHOR(S): Suda, S.; Iwase, Y.;
Morigasaki, N.; Li, Z.-P.
CORPORATE SOURCE: Department of Environmental & Chemical Engineering,
Kogakuin University, Hachioji, 192-0015, Japan
SOURCE: Advanced Materials for Energy Conversion II,
Proceedings of a Symposium held during the TMS Annual
Meeting, 2nd, Charlotte, NC, United States, Mar.
14-18, 2004 (2004), Meeting Date 2004, 123-133.
Editor(s): Chandra, Dhanesh; Bautista, Renato G.;
Schlapbach, Lous. Minerals, Metals & Materials
Society: Warrendale, Pa.
CODEN: 69FHAF; ISBN: 0-87339-574-3
DOCUMENT TYPE: Conference
LANGUAGE: English
AB A new production process for NaBH4 as H storage material was developed by
using NaBO2, according to $NaBO_2 + 2Mg + 2H_2 \rightarrow NaBH_4 + 2MgO$ at a
fast reaction rate by changing the temperature of the system dynamically. The
process uses the transitional states of H at the surface of Mg when it is
hydrogenated and dehydrogenated at a constant rate of temperature change,
 dT/dt .
The rate of conversion is dependent on dT/dt and is regulated by the
extent of oxidation of the Mg particles. The role of H- in $Mg \cdot (2H^-)$
under transient conditions, $Mg + H_2 \rightarrow MgH_2 \rightarrow Mg \cdot (2H^-)$, is discussed with regard to the exchange reaction between
NaBO2 and NaBH4.
REFERENCE COUNT: 30 THERE ARE 30 CITED REFERENCES AVAILABLE FOR THIS
RECORD. ALL CITATIONS AVAILABLE IN THE RE FORMAT

L53 ANSWER 8 OF 46 HCPLUS COPYRIGHT 2006 ACS on STN
ACCESSION NUMBER: 2004:83205 HCPLUS
DOCUMENT NUMBER: 140:256177
TITLE: Evaluation of alkaline borohydride solutions as the
fuel for fuel cell
AUTHOR(S): Li, Z. P.; Liu, B. H.; Arai, K.; Asaba, K.;
Suda, S.
CORPORATE SOURCE: Department of Environmental & Chemical Engineering,
Kogakuin University, Hachioji-shi, Tokyo, 192-0015,
Japan
SOURCE: Journal of Power Sources (2004), 126(1-2), 28-33
CODEN: JPSODZ; ISSN: 0378-7753
PUBLISHER: Elsevier Science B.V.
DOCUMENT TYPE: Journal
LANGUAGE: English
AB A fuel cell was assembled using alkaline borohydride solns. as the fuel. The
sp. gr., sp. viscosity, and m.ps. of the borohydride solns. were
evaluated. With increasing the NaBH4 concentration, the sp. gr. was decreased
but the viscosity was increased consistently. The effects of NaBH4 and
NaOH concns. on the cell and electrode polarizations were investigated.
The increase of NaBH4 concentration improved the anode performance only a
little
but significantly increased the cathode polarization due to the BH4-
crossover. As a result, the cell polarization increased when using highly
concentrated borohydride solns. Based on the Nernst's equation, the increase
of
NaOH concentration is favorable for the anode, but detrimental to the cathode
polarization. In principle, the cell polarization has nothing to do with
the NaOH concentration according to the cell reaction $\text{NaBH}_4 + \text{O}_2 = \text{NaBO}_2 + 2\text{H}_2\text{O}$.
However, the exptl. results showed that the high concentration of NaOH affected
the cell polarization at higher current densities.
REFERENCE COUNT: 12 THERE ARE 12 CITED REFERENCES AVAILABLE FOR THIS
RECORD. ALL CITATIONS AVAILABLE IN THE RE FORMAT

Valenrod 10_721479

L53 ANSWER 9 OF 46 HCAPLUS COPYRIGHT 2006 ACS on STN
ACCESSION NUMBER: 2003:601030 HCAPLUS
DOCUMENT NUMBER: 139:309993
TITLE: Protide compounds in hydrogen storage systems
AUTHOR(S): Li, Z. P.; Liu, B. H.; Arai, K.; Morigazaki, N.; Suda, S.
CORPORATE SOURCE: Department of Environmental and Chemical Engineering,
Kogakuin University, Hachioji-shi, Tokyo, 192-0015,
Japan
SOURCE: Journal of Alloys and Compounds (2003), 356-357,
469-474
CODEN: JALCEU; ISSN: 0925-8388
PUBLISHER: Elsevier Science B.V.
DOCUMENT TYPE: Journal
LANGUAGE: English
AB Based on 3 chemical states of hydrogen, H-, H0 and H+, a triangular H energy system is proposed. The energy transfer between the 3 states is discussed. Expts. prove that NaBH4 as a H- carrier can release its energy directly through a borohydride fuel cell or generate H2 for polymer electrolyte membrane fuel cells. The used fuel, Na meta-borate, can be converted back to NaBH4 through a reaction with MgH2.
REFERENCE COUNT: 13 THERE ARE 13 CITED REFERENCES AVAILABLE FOR THIS RECORD. ALL CITATIONS AVAILABLE IN THE RE FORMAT

L53 ANSWER 10 OF 46 HCPLUS COPYRIGHT 2006 ACS on STN
ACCESSION NUMBER: 2003:458475 HCPLUS
DOCUMENT NUMBER: 139:216877
TITLE: A Fuel Cell Development for Using Borohydrides as the Fuel
AUTHOR(S): Li, Z. P.; Liu, B. H.; Arai, K.; Suda, S.
CORPORATE SOURCE: Department of Environmental and Chemical Engineering,
Kogakuin University, Nakano-machi 2665-1,
Hachioji-shi, Tokyo, 192-0015, Japan
SOURCE: Journal of the Electrochemical Society (2003), 150(7),
A868-A872
CODEN: JESOAN; ISSN: 0013-4651
PUBLISHER: Electrochemical Society
DOCUMENT TYPE: Journal
LANGUAGE: English
AB A fuel cell was developed using borohydride solns. as the fuel. The cell consisted of an anode made of a Zr-Ni alloy, a cathode made of Pt/C, and a Na⁺ form Nafion membrane as the electrolyte. The borohydride-fueled cell showed an open-circuit voltage of 1.3 V, compared with 1.0 V for a hydrogen gas-fueled one. The anode exhibited a small polarization property compared with the cathode. The cathode polarization was the main reason for the cell voltage drop with increasing currents. When a Nafion membrane was used as the electrolyte, cations (Na⁺) were the charge carrier in it. Compared with Nafion 112 membrane, Nafion 117 membrane demonstrated a considerable resistance to borohydride crossover and resulted in acceptable cell performance. However, there are several problems such as H₂ evolution during operation, BH-4 crossover, NaOH accumulation at the cathode, and NaBO₂ accumulation at the anode in recent systems. Further effort is needed to develop the fuel cell using borohydrides as the fuel.
REFERENCE COUNT: 15 THERE ARE 15 CITED REFERENCES AVAILABLE FOR THIS RECORD. ALL CITATIONS AVAILABLE IN THE RE FORMAT

L53 ANSWER 11 OF 46 HCAPLUS COPYRIGHT 2006 ACS on STN

ACCESSION NUMBER: 2003:324965 HCAPLUS

DOCUMENT NUMBER: 139:103130

TITLE: Preparation of potassium borohydride by a mechano-chemical reaction of saline hydrides with dehydrated borate through ball milling

AUTHOR(S): Li, Zhou Peng; Liu, Bin Hong; Morigasaki, Nobuto; Suda, Sejirau

CORPORATE SOURCE: Department of Environmental and Chemical Engineering, Kogakuin University, Hachioji-shi, Tokyo, 192-0015, Japan

SOURCE: Journal of Alloys and Compounds (2003), 354(1-2), 243-247

CODEN: JALCEU; ISSN: 0925-8388

PUBLISHER: Elsevier Science B.V.

DOCUMENT TYPE: Journal

LANGUAGE: English

AB A convenient method was developed to synthesize potassium borohydride by a mechanochem. reaction of saline hydrides with dehydrated borates using a planetary ball mill. Among the tested saline hydrides, MgH₂ was the most effective reactant for borohydride formation. In order to improve the borohydride yield, it was necessary to add an excess amount of MgH₂. When a 35% excess of MgH₂ was added, potassium borohydride yield reached 100%. The borohydride formation was strongly dependent on the water content in reactants. When the water content in KBO₂ samples was >24.8 weight%

(equivalent to KBO₂·1.5H₂O), no borate was converted into borohydride by the mechanochem. reaction.

REFERENCE COUNT: 8 THERE ARE 8 CITED REFERENCES AVAILABLE FOR THIS RECORD. ALL CITATIONS AVAILABLE IN THE RE FORMAT

L53 ANSWER 12 OF 46 HCPLUS COPYRIGHT 2006 ACS on STN
ACCESSION NUMBER: 2003:99330 HCPLUS
DOCUMENT NUMBER: 138:371547
TITLE: A study on recycle of borohydride liquid fuel
AUTHOR(S): **Morigasaki, Nobuto; Tanisawa, Kazuhiro;**
Li, Zhoupeng; Suda, Seijirau
CORPORATE SOURCE: Graduate School of Applied Chemistry and Chemical
Engineering, Kogakuin University, Tokyo, Japan
SOURCE: Kogakuin Daigaku Kenkyu Hokoku (2002), 93, 55-59
CODEN: KDKHAY; ISSN: 0368-5098
PUBLISHER: Kogakuin Daigaku
DOCUMENT TYPE: Journal
LANGUAGE: Japanese
AB Borohydrides can be formed by the reaction of anhydrous metaborate with Mg hydride. However, the borohydride conversion rate was influenced by the metaborate source. Anhydrous sodium metaborate can be produced by dehydrating the $\text{NaBO}_2 \cdot 4\text{H}_2\text{O}$ crystals from used fuel or the commercialized $\text{NaBO}_2 \cdot 4\text{H}_2\text{O}$. The NaBH_4 conversion rate based on NaBO_2 from used fuel was lower than that based on the anhydrous borate form the commercialized $\text{NaBO}_2 \cdot 4\text{H}_2\text{O}$. It was because the sample obtained from used fuel was not completely dehydrated.

L53 ANSWER 13 OF 46 HCPLUS COPYRIGHT 2006 ACS on STN
ACCESSION NUMBER: 2003:99329 HCPLUS
DOCUMENT NUMBER: 138:371546
TITLE: The electric discharge mechanism of the liquid type
fuel cell using borohydride complex ion
AUTHOR(S): Arai, Kunihito; Chiku, Satoshi; Li, Zhoupeng
; Suda, Seijirau
CORPORATE SOURCE: Materials & Energy Research Institute Tokyo, Ltd.,
Japan
SOURCE: Kogakuin Daigaku Kenkyu Hokoku (2002), 93, 49-53
CODEN: KDKHAY; ISSN: 0368-5098
PUBLISHER: Kogakuin Daigaku
DOCUMENT TYPE: Journal
LANGUAGE: Japanese
AB In this liquid type fuel cell using borohydride complex ion (called "BFC"),
it was found that alkaline type reactions had occurred at each electrode,
though this fuel cell used the pos. ion exchange membrane (i.e. an acid
membrane) as the electrolyte. These reactions are involved in the fact
that K⁺ ion which exists so much at an anode side moved in the electrolyte
as an elec. charge carrier. Moreover, when we used the fluorinated metal
hydride as an anode catalyst, we obtained potential as much as Pd.

L53 ANSWER 14 OF 46 HCPLUS COPYRIGHT 2006 ACS on STN
ACCESSION NUMBER: 2003:99328 HCPLUS
DOCUMENT NUMBER: 138:355708

TITLE: A review on the methods of synthesis of alkali metal borohydrides

AUTHOR(S): Iwase, Yasuyoshi; Li, Zhoupeng;
Suda, Seijirau

CORPORATE SOURCE: Materials & Energy Research Institute Tokyo, Ltd.,
Japan

SOURCE: Kogakuin Daigaku Kenkyu Hokoku (2002), 93, 43-48
CODEN: KDKHAY; ISSN: 0368-5098

PUBLISHER: Kogakuin Daigaku

DOCUMENT TYPE: Journal; General Review

LANGUAGE: Japanese

AB A review of articles and patents on the synthesis methods of alkali metal borohydrides to explore the required conditions for new and more economical methods of producing the alkali metal borohydrides. Based on this overview and our experiment results, we determined the conditions to synthesize

borohydrides (BH_4^-) from borate salts: (a) reaction proceeds under anhydrous condition or in aprotic solvent and (b) borate salts react with other protide compds. (H^-) to form BH_4^- , but never form BH_4^- with proton.

L53 ANSWER 15 OF 46 HCPLUS COPYRIGHT 2006 ACS on STN
ACCESSION NUMBER: 2002:972505 HCPLUS
DOCUMENT NUMBER: 138:347880
TITLE: Preparation of sodium borohydride by the reaction of
MgH₂ with dehydrated borax through ball milling at
room temperature
AUTHOR(S): Li, Z. P.; Morigazaki, N.; Liu, B. H.;
Suda, S.
CORPORATE SOURCE: Department of Environmental and Chemical Engineering,
Kogakuin University, Hachioji-shi, Tokyo, 192-0015,
Japan
SOURCE: Journal of Alloys and Compounds (2003), 349(1-2),
232-236
CODEN: JALCEU; ISSN: 0925-8388
PUBLISHER: Elsevier Science B.V.
DOCUMENT TYPE: Journal
LANGUAGE: English
AB A convenient method was developed to synthesize NaBH₄ by the reaction of
MgH₂ with Na₂B₄O₇ through ball milling at room temperature. To improve the Na
borohydride yield, Na compds. were added to compensate the Na
insufficiency in reactants when MgH₂ instead of NaH was used as the
reducing agent. Na₂CO₃ addition was better than NaOH or Na₂O₂ addition in
increasing the borohydride yield.
REFERENCE COUNT: 5 THERE ARE 5 CITED REFERENCES AVAILABLE FOR THIS
RECORD. ALL CITATIONS AVAILABLE IN THE RE FORMAT

L53 ANSWER 16 OF 46 HCPLUS COPYRIGHT 2006 ACS on STN
ACCESSION NUMBER: 2002:95069 HCPLUS
DOCUMENT NUMBER: 136:312397
TITLE: Effects of surface structure of fluorinated AB₂ alloys
on their electrodes and battery performances
AUTHOR(S): Li, Z. P.; Liu, B. H.; Hitaka, K.;
Suda, S.
CORPORATE SOURCE: Department of Environmental and Chemical Engineering,
Kogakuin University, Hachioji, 192-0015, Japan
SOURCE: Journal of Alloys and Compounds (2002), 330-332,
776-781
CODEN: JALCEU; ISSN: 0925-8388
PUBLISHER: Elsevier Science S.A.
DOCUMENT TYPE: Journal
LANGUAGE: English
AB Fluorination technique effectively improve the activation property and
rate discharge capabilities of AB₂ alloys. In this paper, the performance
of sealed batteries using fluorinated AB₂ alloys were investigated.
Compared with the sealed battery using untreated AB₂ alloys, those using
fluorinated AB₂ alloys showed better activation property, longer cycle
life and higher rate discharge capacity. The rate discharge capacity of
the sub-C sealed battery using duplicated fluorination treated AB₂ alloys
reached up to 2.3 AH at 4C (10.4A). The batteries using fluorinated AB₂
alloys sustained over 300 cycles at 0.2C rate, but its discharge capacity
degraded quickly when being cycled at 1C rate.
REFERENCE COUNT: 13 THERE ARE 13 CITED REFERENCES AVAILABLE FOR THIS
RECORD. ALL CITATIONS AVAILABLE IN THE RE FORMAT

Valenrod 10_721479

L53 ANSWER 17 OF 46 HCPLUS COPYRIGHT 2006 ACS on STN

ACCESSION NUMBER: 2001:300367 HCPLUS

DOCUMENT NUMBER: 135:95104

TITLE: Fluorinated metal hydrides for the catalytic hydrolysis of metal-hydrogen complexes

AUTHOR(S): Suda, S.; Sun, Y.-M.; Uchida, M.; Liu, B.-H.; Mormitu, S.; Arai, K.; Zhou, Y.; Tsukamoto, N.; Candra, Y.; Li, Z.-P.

CORPORATE SOURCE: Chemical Energy Laboratory, Kogakuin University, Tokyo, 192-0015, Japan

SOURCE: Metals and Materials International (2001), 7(1), 73-75
CODEN: MMIECY

PUBLISHER: Korean Institute of Metals and Materials

DOCUMENT TYPE: Journal

LANGUAGE: English

AB This paper is related to the hydrogen storage (H-storage) system composed of an alkaline solution that contains metal-hydrogen complex ions. The BH₄-complex ions stabilized when dissolved in aqueous KOH or NaOH to form homogeneous solns. However, it is necessary to apply an appropriate catalyst for a high rate generation of hydrogen under atmospheric pressure and temperature conditions. The authors intended to develop catalysts suitable for the hydrolysis of BH₄-complex ion solns. and found that Mg₂Ni exhibits excellent catalytic functions in the hydrolysis, in particular, after the fluorination treatment (F-treatment). In this work, the effects of F-treatment on the rates of hydrolysis were evaluated for the hydrided and unhydrided Mg₂Ni. No hydrogenation in addition to F-treatment is required in view of hydrolysis kinetics and cost-effectiveness.

REFERENCE COUNT: 14 THERE ARE 14 CITED REFERENCES AVAILABLE FOR THIS RECORD. ALL CITATIONS AVAILABLE IN THE RE FORMAT

L53 ANSWER 18 OF 46 HCPLUS COPYRIGHT 2006 ACS on STN
ACCESSION NUMBER: 2001:206623 HCPLUS
DOCUMENT NUMBER: 134:313607
TITLE: Catalytic generation of hydrogen by applying
fluorinated-metal hydrides as catalysts
AUTHOR(S): Suda, S.; Sun, Y.-M.; Liu, B.-H.; Zhou, Y.;
Morimitsu, S.; Arai, K.; Tsukamoto, N.; Uchida, M.;
Candra, Y.; Li, Z.-P.
CORPORATE SOURCE: Chemical Energy Laboratory, Department of
Environmental and Chemical Engineering, Kogakuin
University, Hachioji, 192-0015, Japan
SOURCE: Applied Physics A: Materials Science & Processing
(2001), 72(2), 209-212
CODEN: APAMFC; ISSN: 0947-8396
PUBLISHER: Springer-Verlag
DOCUMENT TYPE: Journal
LANGUAGE: English
AB Metal-hydrogen complexes such as NaAlH₄, KBH₄, and NaBH₄ are known as high
H-content materials. The highly reactive natures of these materials
against moist air and water can be easily stabilized in aqueous KOH and NaOH
solns. Accordingly, it is required to develop catalysts suitable for
generating hydrogen from the stabilized metal-hydrogen complexes in alkaline
solns. This work is aimed at developing catalysts that can generate
hydrogen from such solns. with considerably high kinetics under moderate
temperature and pressure conditions. We have found that Mg₂Ni, a typical
high-temperature hydriding alloy, exhibits excellent functions as a catalyst
for
the hydrolysis of BH₄--ion-containing solns. The fluorination-treatment
(F-treatment) effects on granular particles of Mg₂Ni and Mg₂NiH₄ are
reported in this paper.
REFERENCE COUNT: 39 THERE ARE 39 CITED REFERENCES AVAILABLE FOR THIS
RECORD. ALL CITATIONS AVAILABLE IN THE RE FORMAT

L53 ANSWER 19 OF 46 HCAPLUS COPYRIGHT 2006 ACS on STN

ACCESSION NUMBER: 2001:88882 HCAPLUS

DOCUMENT NUMBER: 134:298330

TITLE: Effects of fluorination of AB2-type alloys and of mixing with AB5-type alloys on the charge-discharge characteristics

AUTHOR(S): Higuchi, E.; Toyoda, E.; Li, Z. P.;

Suda, S.; Inoue, H.; Nohara, S.; Iwakura, C.

CORPORATE SOURCE: Faculty of Engineering, Department of Environmental and Chemical Engineering, Chemical Energy Lab, Kogakuin University, Hachioji-shi, Tokyo, 192-0015, Japan

SOURCE: Electrochimica Acta (2001), 46(8), 1191-1194

CODEN: ELCAAV; ISSN: 0013-4686

PUBLISHER: Elsevier Science Ltd.

DOCUMENT TYPE: Journal

LANGUAGE: English

AB The effects of fluorination (F-treatment) of AB2-type Laves-phase Zr0.9Ti0.1V0.2Mn0.6Co0.1Ni1.1 (AB2) alloys and of mixing with untreated AB5-type MnNi3.68Co0.78Mn0.36Al0.28 (AB5) alloys on their charge-discharge characteristics were investigated. The F-treatment of the AB2 alloy greatly increased initial activation rate and discharge capacity and slightly improved high-rate dischargeability. The mixing of the F-treated AB2 alloy with untreated AB5 (<25 μ m) alloy decreased somewhat the maximum discharge capacity compared to the fluorinated AB2 alloy but it greatly improved the high-rate dischargeability. In particular, the high-rate dischargeability for the AB2(F)-AB5 (<25 μ m) mixture pos. deviated from the estimated values based on the maximum discharge capacity for both alloys

and

their content, suggesting that the AB5 alloy can work as a catalyst for dehydriding the AB2 alloy hydride besides as a hydrogen reservoir; a synergistic effect appeared by the mixing.

REFERENCE COUNT: 17 THERE ARE 17 CITED REFERENCES AVAILABLE FOR THIS RECORD. ALL CITATIONS AVAILABLE IN THE RE FORMAT

L53 ANSWER 20 OF 46 HCPLUS COPYRIGHT 2006 ACS on STN
ACCESSION NUMBER: 2000:147662 HCPLUS
DOCUMENT NUMBER: 132:200265
TITLE: Effects of fluorination temperature on surface structure and electrochemical properties of AB2 electrode alloys
AUTHOR(S): Li, Z. P.; Higuchi, E.; Liu, B. H.; Suda, S.
CORPORATE SOURCE: Department of Environmental and Chemical Engineering, Kogakuin University, Hachioji, 192-0015, Japan
SOURCE: Electrochimica Acta (2000), 45(11), 1773-1779
CODEN: ELCAAV; ISSN: 0013-4686
PUBLISHER: Elsevier Science Ltd.
DOCUMENT TYPE: Journal
LANGUAGE: English
AB In order to accelerate the activation of AB2 electrode alloys, a new fluorination treatment has been developed which removes oxides from alloy surfaces and implants metallic Ni into the alloy surface layer. The behavior of Ni during fluorination treatment depends strongly on the treatment temperature. The constituent Ni of the alloy dissolves into the treatment solution at a low temperature, but Ni complex ion in treatment solution is chemically reduced by constituent elements at a high temperature. The temperature is the key factor of the fluorination treatment. The electrochem. properties were improved through increasing the treatment temperature due to the Ni implantation into surface layer.
REFERENCE COUNT: 10 THERE ARE 10 CITED REFERENCES AVAILABLE FOR THIS RECORD. ALL CITATIONS AVAILABLE IN THE RE FORMAT

Valenrod 10_721479

L53 ANSWER 21 OF 46 HCPLUS COPYRIGHT 2006 ACS on STN
ACCESSION NUMBER: 2000:6032 HCPLUS
DOCUMENT NUMBER: 132:95674
TITLE: Electrochemical properties and characteristics of a
fluorinated AB₂-alloy
AUTHOR(S): Li, Z. P.; Higuchi, E.; Liu, B. H.;
Suda, S.
CORPORATE SOURCE: Department of Environmental & Chemical Engineering,
Kogakuin University, Hachioji, 192-0015, Japan
SOURCE: Journal of Alloys and Compounds (1999), 293-295,
593-600
CODEN: JALCEU; ISSN: 0925-8388
PUBLISHER: Elsevier Science S.A.
DOCUMENT TYPE: Journal
LANGUAGE: English
AB A series of fluorination techniques have been developed in this laboratory to
improve the surface condition of AB₂ alloy. The fluorination was found
effective for removing oxide layer and implanting catalytic Ni to the
alloy surfaces. The developed fluorination techniques improved
considerably the initial activation characteristics, decreased the
electrochem. reaction impedance and increased the rate capacity of AB₂
electrodes.
REFERENCE COUNT: 7 THERE ARE 7 CITED REFERENCES AVAILABLE FOR THIS
RECORD. ALL CITATIONS AVAILABLE IN THE RE FORMAT

L53 ANSWER 22 OF 46 HCPLUS COPYRIGHT 2006 ACS on STN
ACCESSION NUMBER: 1999:482282 HCPLUS
DOCUMENT NUMBER: 131:274162
TITLE: Fluorinated Laves-AB2 alloy as the negative electrode
in Ni-MH battery
AUTHOR(S): Higuchi, Eiji; Miyoshi, Hirokazu; Sugimoto, Hiroyuki;
Li, Zhoupeng; Suda, Seijirau
CORPORATE SOURCE: Grad. Sch. Ind. Chem., Kogakuin Univ., Tokyo,
163-8677, Japan
SOURCE: Kogakuin Daigaku Kenkyu Hokoku (1999), 86, 69-75
CODEN: KDKHAY; ISSN: 0368-5098
PUBLISHER: Kogakuin Daigaku
DOCUMENT TYPE: Journal
LANGUAGE: Japanese

AB AB5 hydriding alloys as neg. electrode materials have been well developed for Ni-MH battery. Many scientists have paid attention to AB2 alloy due to the high-capacity of AB2 alloy compared with AB5 alloy. However, AB2 alloy usually showed poor activation and kinetic properties. It is possible to overcome these shortcomings by surface modification through fluorination. Aiming for practical use, we investigated the performances of practical electrode of fluorinated AB2 alloy to improve the rate of discharge capability and cycle life. It is found that the connection of lead line to electrode conductor of the electrode, and the pore size of the Ni foam exerted the big influence on electrode properties.

Valenrod 10_721479

L53 ANSWER 23 OF 46 HCPLUS COPYRIGHT 2006 ACS on STN
ACCESSION NUMBER: 1999:482280 HCPLUS
DOCUMENT NUMBER: 131:274160
TITLE: Effects of the fluorination treatment on the surface structure and electrochemical properties of Zr-based AB2-type hydriding alloys
AUTHOR(S): Okutsu, Akira; Nakamura, Kazuhiro; Li, Zhoupeng; Suda, Seijirau
CORPORATE SOURCE: Grad. Sch. Ind. Chem., Kogakuin Univ., Tokyo, 163-8677, Japan
SOURCE: Kogakuin Daigaku Kenkyu Hokoku (1999), 86, 57-62
CODEN: KDKHAY; ISSN: 0368-5098
PUBLISHER: Kogakuin Daigaku
DOCUMENT TYPE: Journal
LANGUAGE: Japanese
AB AB5 alloys have been used in Ni-MH battery as neg. electrode materials. In this work, the activation and reactivity of Zr-based AB2 alloys whose capacities were larger than AB5 alloys, have been improved by fluorination. The effects of fluorination technique combining with plenary ball milling on electrochem. properties have been investigated to improve performances of AB2 alloys for developing the high power secondary battery of elec. vehicle. It was found that oxides can be removed from alloy surfaces while metallic Ni can be implanted to alloy surfaces so that the sp. surface area of the fluorinated AB2 alloy was increased. The electrochem. activation, high rate charge and discharge capability were improved through surface modification by fluorination.

L53 ANSWER 24 OF 46 HCPLUS COPYRIGHT 2006 ACS on STN
ACCESSION NUMBER: 1998:544811 HCPLUS
DOCUMENT NUMBER: 129:163814
TITLE: Surface passivation of metal hydrides for applications
AUTHOR(S): Suda, S.; Li, Z.-P.; Sun, Y.-M.;
Liu, B.-H.; Gao, X.-P.
CORPORATE SOURCE: Department of Environmental & Chemical Engineering,
Kogakuin University, Tokyo, 192-0015, Japan
SOURCE: Materials Research Society Symposium Proceedings
(1998), 513(Hydrogen in Semiconductors and Metals),
25-36
CODEN: MRSPDH; ISSN: 0272-9172
PUBLISHER: Materials Research Society
DOCUMENT TYPE: Journal; General Review
LANGUAGE: English
AB A review with many refs. Properties and characteristics of hydriding
alloys are strongly dependent on surface compns. and morphologies. For
instance, oxides such as La₂O₃ on AB₅ alloys and ZrO₂ on AB₂, AB, and
body-centered-cubic (BCC) alloys act as the barriers for the conversion of
mol. and ionic hydrogen to atomic hydrogen at the surface, thus reducing the
kinetics in both the gas-solid and electrochem. reactions. Alloy surfaces
chemical treated by an aqueous F-ion containing solution have been developed
to solve
such problems. F-treated surfaces exhibit significantly improved
characteristics in regard to the hydrogen uptakes and the protection
against impurities and electrolyte solution. In addition, highly conductive
metallic Ni layers can be formed on the surface of the alloy particles by
the fluorination. The authors report the properties and characteristics
of fluorinated hydriding alloys, mainly of a typical AB₂ Laves phase
material which represents the difficult activation characteristics and
poor long-term durability during electrochem. charge/discharge cycles.
REFERENCE COUNT: 34 THERE ARE 34 CITED REFERENCES AVAILABLE FOR THIS
RECORD. ALL CITATIONS AVAILABLE IN THE RE FORMAT

L53 ANSWER 25 OF 46 HCPLUS COPYRIGHT 2006 ACS on STN
ACCESSION NUMBER: 1998:51782 HCPLUS
DOCUMENT NUMBER: 128:117289
TITLE: The effects of fluorination solution composition on
the electrochemical properties of
Zr0.9Ti0.1V0.2Mn0.6Co0.1Ni1.1 alloy
AUTHOR(S): Higuchi, Eiji; Sakasita, Masaya; Li, Zhoupeng
; Suda, Seijirau
CORPORATE SOURCE: Dep. Ind. Chem., Kogakuin Univ., Tokyo, 163-91, Japan
SOURCE: Kogakuin Daigaku Kenkyu Hokoku (1997), 83, 27-30
CODEN: KDKHAY; ISSN: 0368-5098
PUBLISHER: Kogakuin Daigaku
DOCUMENT TYPE: Journal
LANGUAGE: Japanese
AB Effects of components of fluorination treatment solution on discharge
properties of Zr-based AB₂-type alloy Zr0.9Ti0.1V0.2Mn0.6Co0.1Ni1.1 were
studied. It was found that Ni amount implanted on the hydriding alloy
surface showed significant effects on the discharge capacity. The element
dissoln. in the alloy during the treatment also showed similar effects.

L53 ANSWER 26 OF 46 HCPLUS COPYRIGHT 2006 ACS on STN
ACCESSION NUMBER: 1998:51781 HCPLUS
DOCUMENT NUMBER: 128:117288
TITLE: The effects of fluorination treatment on the
electrochemical properties of Zr-based AB2-type
hydriding alloys
AUTHOR(S): Okutsu, Akira; Sakasita, Masaya; Li, Zhoupeng
; Suda, Seijirau
CORPORATE SOURCE: Dep. Ind. Chem., Kogakuin Univ., Tokyo, 163-91, Japan
SOURCE: Kogakuin Daigaku Kenkyu Hokoku (1997), 83, 23-26
CODEN: KDKHAY; ISSN: 0368-5098
PUBLISHER: Kogakuin Daigaku
DOCUMENT TYPE: Journal
LANGUAGE: Japanese
AB Fluorination treatment effects on discharge properties of Zr-based
AB2-type hydriding alloys were studied. It was found that electrochem.
activation properties could be improved and discharge capacity could be
increased by implanting Ni on the surfaces of the hydriding alloys through
the fluorination treatment.

Valenrod 10_721479

L53 ANSWER 27 OF 46 HCPLUS COPYRIGHT 2006 ACS on STN

ACCESSION NUMBER: 1997:456292 HCPLUS

DOCUMENT NUMBER: 127:164304

TITLE: Fluorination mechanism and its effects on the electrochemical properties of metal hydrides

AUTHOR(S): Sakashita, M.; Li, Z. P.; Suda, S.

CORPORATE SOURCE: Department of Chemical Engineering, Kogakuin University, Nakano-machi 2665-1, Hachioji-shi, Tokyo, 192, Japan

SOURCE: Journal of Alloys and Compounds (1997), 253-254, 500-505

CODEN: JALCEU; ISSN: 0925-8388

PUBLISHER: Elsevier

DOCUMENT TYPE: Journal

LANGUAGE: English

AB Fluorination of hydriding alloys has been found effective for improving durability and initial discharge characteristics of hydride electrodes in Ni-MH rechargeable batteries. However, it has also been found that it decreases the discharge capacity because of the fluoride formed on the surface which decreases the elec. conductivity of the electrode. A more advanced

technique has been developed in this laboratory to implant metallic Ni in the fluoride layer to form a functionally graded surface layer. The proposed fluorination technique considerably improves the durability and initial activation characteristics of AB₅-types of hydride, electrodes and also the initial activation characteristics of AB₂ electrodes. The fluorination was found effective for removing the oxide layer which acts as an impedance to electrochem. hydrogen uptake.

REFERENCE COUNT: 7 THERE ARE 7 CITED REFERENCES AVAILABLE FOR THIS RECORD. ALL CITATIONS AVAILABLE IN THE RE FORMAT

L53 ANSWER 28 OF 46 HCPLUS COPYRIGHT 2006 ACS on STN
ACCESSION NUMBER: 1996:430623 HCPLUS
DOCUMENT NUMBER: 125:92484
TITLE: An investigation of Ca-based hydride electrode materials
AUTHOR(S): Li, Z. P.; Matsuoka, T.; Suda, S.
CORPORATE SOURCE: Dep. Chem. Eng., Kogakuin Univ., Tokyo, 192, Japan
SOURCE: Vacuum (1996), 47(6-8, Proceedings of the 13th International Vacuum Congress and the 9th International Conference on Solid Surfaces, 1995), 893-897
CODEN: VACUAV; ISSN: 0042-207X
PUBLISHER: Elsevier
DOCUMENT TYPE: Journal
LANGUAGE: English
AB The effects of La-substitution for Ca in Ca_{1-x}La_xNi_{5.7}Al_{0.3} on metallog., structure and electrochem. properties have been investigated. It was found that Ca and Al segregated over the grain boundaries of the alloys. Based on La substituted CaNi₅-type alloys, we studied the Zr and Ti addition effects as well. It was found that Zr-Ni phases or Ti-Ni phases deposited on the boundaries of AB₅ phase improved the anti-corrosion property of the alloys. Annealing improves the electrochem. durability of the Ca_{1-x}La_xNi_{4.7}Al_{0.3} alloys by reducing A₂B₇ phase and Ca segregation on the grain boundary.

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L53 ANSWER 29 OF 46 HCPLUS COPYRIGHT 2006 ACS on STN
ACCESSION NUMBER: 1996:199099 HCPLUS
DOCUMENT NUMBER: 124:265514
TITLE: A new family of hydride electrode materials based on
CaNi5-type alloys
AUTHOR (S): Matsuoka, Takashi; Li, Zhoupeng; Uchida,
Masaki; Suda, Seijirau
CORPORATE SOURCE: Kogakuin Univ., Tokyo, 163-91, Japan
SOURCE: Kogakuin Daigaku Kenkyu Hokoku (1995), 78, 85-7
CODEN: KDKHAY; ISSN: 0368-5098
PUBLISHER: Kogakuin Daigaku
DOCUMENT TYPE: Journal
LANGUAGE: Japanese
AB Charge-discharge characteristic of $Ca_{1-x}AxNi_5-y-zMyNz$ ($A = La, Ce, Mg, Sr$;
 $M = Sn, Si, Co$; $N = Al$; $0 \leq x, y, z \leq 1$) hydride electrode were
examined. The discharge capacity of the Ca-based alloys was 300-350 mAh/g at
discharge current of 150 mA/g and was lower than that of CaNi5. In
comparison to the M and/or N substitution, A substitution was effective on
the improvement of cycling durability and kept the discharge capacity at a
high level. The most durable electrodes were obtained with $A = La$ and $x =$
 0.6 .

L53 ANSWER 30 OF 46 HCPLUS COPYRIGHT 2006 ACS on STN
ACCESSION NUMBER: 1996:106282 HCPLUS
DOCUMENT NUMBER: 124:180993
TITLE: The Ni content and F-treatment effects on the
electrochemical cycling behavior of $Zr0.5Ti0.5Ni_x$
alloys
AUTHOR(S): Li, Zhoupeng; Takiguchi, Katsutoshi;
Suda, Seijirau
CORPORATE SOURCE: Kogakuin Univ., Tokyo, 163-91, Japan
SOURCE: Kogakuin Daigaku Kenkyu Hokoku (1995), 78, 81-4
CODEN: KDKHAY; ISSN: 0368-5098
PUBLISHER: Kogakuin Daigaku
DOCUMENT TYPE: Journal
LANGUAGE: English
AB The alloys were prepared through arc melting of pure metals and pulverized
to 26 μm and the powders were treated with F ion-containing solution. With
increasing of Ni content in the alloy of $x = 1.25-1.675$, the discharge
capacity increased and the electrochem. activation became easier, but with
further addition of Ni discharge capacity decreased. The alloys of high
content of Ni ($x = 1.5-1.675$) had higher discharge capacity but the alloys
of lower capacity had better durability. F-treatment of alloys containing low
Ni as $x \leq 1.25$ decreased discharge capacity because of decrease of
Ni content at the surface.

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L53 ANSWER 31 OF 46 HCPLUS COPYRIGHT 2006 ACS on STN
ACCESSION NUMBER: 1996:106281 HCPLUS
DOCUMENT NUMBER: 124:180992
TITLE: Effects of hydriding-dehydriding cycling on the P-C-T
and electrochemical properties of Ca1-
 χ La_xNi4.7Al0.3 alloys
AUTHOR(S): Li, Zhoupeng; Suda, Seijirau
CORPORATE SOURCE: Kogakuin Univ., Tokyo, 163-91, Japan
SOURCE: Kogakuin Daigaku Kenkyu Hokoku (1995), 78, 77-80
CODEN: KDKHAY; ISSN: 0368-5098
PUBLISHER: Kogakuin Daigaku
DOCUMENT TYPE: Journal
LANGUAGE: English
AB Electrodes were made through mixing the alloys of $x = 0-0.6$ with Ni and
PTFE in weight ratio of 5:15:1 and through packing in Ni grid and pressing
under 4 ton/cm² to pellets of 13 mm dia. in 1.5 mm thickness. The cycling
was done for composition of charge of 200 mA/g of 3 h and discharge of 150 mA/g
in 6M KOH of 20° to 1.0 V vs. Ni/Ni hydroxide electrode. For XRD
pellets of alloy/PTFE of 5/1 weight/weight were prepared. After 40 cycles of
hydriding/dehydriding at 40° no CaH₂ signal was found. From P-C-T
properties degradation of capacity was not observed after 50 cycles.
La-substituted Ca-based alloys showed strong anti-disproportionation
tendency to stabilize the lattice.

L53 ANSWER 32 OF 46 HCAPLUS COPYRIGHT 2006 ACS on STN
ACCESSION NUMBER: 1996:71040 HCAPLUS
DOCUMENT NUMBER: 124:152479
TITLE: The annealing effect on the durabilities of La
substituted CaNi5-type alloys
AUTHOR(S): Li, Zhoupeng; Nagai, Susumu; Suda,
Seijirau
CORPORATE SOURCE: Kogakuin Univ., Tokyo, 163-91, Japan
SOURCE: Kogakuin Daigaku Kenkyu Hokoku (1995), 78, 71-5
CODEN: KDKHAY; ISSN: 0368-5098
PUBLISHER: Kogakuin Daigaku
DOCUMENT TYPE: Journal
LANGUAGE: English
AB The annealing effects of Ca-La series alloys on homogenization and
durability improvement were investigated, using x-ray diffraction,
microg., and electrochem. measurement. The annealing was good for
homogenization but showed no effect on the electrochem. durability of the
La-free Ca-based alloy. When La-substituted CaNi5 type alloy was annealed
at 900°C for 3 h., the phase structure changed a little, but the
electrochem. durability was improved. Higher temperature and longer time
annealing was effective for the electrochem. cycling durability of the
Ca-La series alloys. The discharge capacity of the annealed alloys
decreased because of the vaporizing of Ca during annealing for a long
time.

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L53 ANSWER 33 OF 46 HCPLUS COPYRIGHT 2006 ACS on STN
ACCESSION NUMBER: 1996:44265 HCPLUS
DOCUMENT NUMBER: 124:94143
TITLE: The annealing effect on the durabilities of
La-substituted CaNi₅-type alloys
AUTHOR(S): Li, Z. P.; Suda, S.
CORPORATE SOURCE: Tokyo, 192, Japan
SOURCE: Journal of Alloys and Compounds (1995), 231(1-2),
835-40
CODEN: JALCEU; ISSN: 0925-8388
PUBLISHER: Elsevier
DOCUMENT TYPE: Journal
LANGUAGE: English
AB The annealing effects of the Ca-La series of alloys on homogenization and durability improvement have been investigated using X-ray diffraction, microg. and electrochem. measurement. It was found that annealing has no effect on the electrochem. durability of the La-free Ca-containing alloy although it can be well homogenized. The phase structure of La-substituted CaNi₅-type alloys was only changed a little by annealing the alloy at 900°C for 3h, but the electrochem. durability could be improved. A high temperature, longer anneal of the Ca-La series of alloys enabled better electrochem. cycling durability to be attained. The discharge capacity of the annealed alloys was decreased a little because of the vaporization of Ca when annealing for a long time.

L53 ANSWER 34 OF 46 HCPLUS COPYRIGHT 2006 ACS on STN
ACCESSION NUMBER: 1996:44250 HCPLUS
DOCUMENT NUMBER: 124:121950
TITLE: A new family of hydride electrode materials based on
CaNi5-type alloys
AUTHOR(S): Li, Z. P.; Suda, S.
CORPORATE SOURCE: Tokyo, 1992, Japan
SOURCE: Journal of Alloys and Compounds (1995), 231(1-2),
751-4
CODEN: JALCEU; ISSN: 0925-8388
PUBLISHER: Elsevier
DOCUMENT TYPE: Journal
LANGUAGE: English
AB A new family of low-cost hydride electrode materials based on CaNi5 was studied. CaNi5-series alloys showed the same charge-discharge potential as RNi5-type alloys (R = rare earth metal). A CaNi5 electrode can reach the capacity of 400 mA-h/g but only can last a few cycles in an alkaline electrolyte. The extra high capacity of CaNi5 might be brought by Ca corrosion. The CaNi5-type alloys after element substitution demonstrated a discharge capacity of 300-350 mA-h/g at a discharge current of 150 mA/g. Comparing with B-side substitution in AB5, A-side substitution is more effective on the durability improvement and keeps the discharge capacity at a high value. Rare earth and alkaline earth substitution can dramatically improve the cycling durability. This is attributed to the lattice stabilization of the alloy and size stabilization of the hydride particles as well as anti-corrosion improvement of the alloy. Zr and Ti addition in Ca0.4La0.6Ni4.7Al0.3 can improve the durability further. The A-rich non-stoichiometric alloy (AB5-type Ca-based alloy) showed higher discharge capacity but poorer durability compared with a stoichiometric compound

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L53 ANSWER 35 OF 46 HCPLUS COPYRIGHT 2006 ACS on STN
ACCESSION NUMBER: 1996:44220 HCPLUS
DOCUMENT NUMBER: 124:122029
TITLE: Effects of hydriding-dehydriding cycling on P-C-T and
electrochemical properties of La-Ca-Ni-Al alloys
AUTHOR(S): Li, Z. P.; Suda, S.
CORPORATE SOURCE: Tokyo, 192, Japan
SOURCE: Journal of Alloys and Compounds (1995), 231(1-2),
594-7
CODEN: JALCEU; ISSN: 0925-8388
PUBLISHER: Elsevier
DOCUMENT TYPE: Journal
LANGUAGE: English
AB The effect of hydriding-dehydriding cycling on electrochem. capacity,
durability, and pressure-composition-temperature (P-C-T) characteristics of
Ca_{1-x}La_xNi_{4.7}Al_{0.3} alloys (x = 0, 0.2, and 0.6) was investigated at
40°. The La-substituted alloys showed a smaller electrochem.
capacity degradation than CaNi₅ after 40 hydriding-dehydriding cycles at
40°. Compared with uncycled alloys, La-substituted Ca-based alloys
showed better durability with the exception of La-free CaNi_{4.7}Al_{0.3}.

L53 ANSWER 36 OF 46 HCPLUS COPYRIGHT 2006 ACS on STN
ACCESSION NUMBER: 1995:649767 HCPLUS
DOCUMENT NUMBER: 123:95902
TITLE: Electrochemical properties of the chemically treated
LaNi4.7Al0.3 and MnNi3.5Co0.7Al0.8
AUTHOR(S): Yan, D. -Y.; Li, Z. -P.; Suda, S.
CORPORATE SOURCE: Department Chemical Engineering, Kogakuin University,
Hachioji, 192, Japan
SOURCE: Transactions of the Materials Research Society of
Japan (1994), 18B, 1217-19
CODEN: TMRJE3; ISSN: 1382-3469
DOCUMENT TYPE: Journal
LANGUAGE: English
AB The effects of the chemical treatment by an aqueous solution which contains
fluorine
(F-treatment) on the electrochem. properties of LaNi4.7Al0.3 and
MnNi3.5Co0.7Al0.8 alloys was investigated. The F-treatment improved the
electrochem. activation characteristics of LaNi4.7Al0.3. The treated
MnNi3.5Co0.7Al0.8 improved significantly the charge discharge cycle life.

L53 ANSWER 37 OF 46 HCPLUS COPYRIGHT 2006 ACS on STN
ACCESSION NUMBER: 1995:649766 HCPLUS
DOCUMENT NUMBER: 123:95901
TITLE: Activation of Zr0.5Ti0.5V0.75Ni1.25 electrodes by hot
alkaline solutions
AUTHOR(S): Yan, Deyi; Sandrock, Gary; Li, Zhoupeng;
Wang, Xianglong; Suda, Seijirau
CORPORATE SOURCE: Department Chemical Engineering, Kogakuin University,
Hachioji, 192, Japan
SOURCE: Transactions of the Materials Research Society of
Japan (1994), 18B, 1213-16
CODEN: TMRJE3; ISSN: 1382-3469
DOCUMENT TYPE: Journal
LANGUAGE: English
AB Hot alkaline treatment of the multiphase alloy Zr0.5Ti0.5V0.75Ni1.25 results
in significant surface corrosion, with at least some of the corrosion
products being soluble in the treating solution. The corrosion process
generates
H that is largely absorbed by the alloy during treatment, resulting in
heterogeneous particle cracking and significant electrochem. discharge
capacity (C0) even before the 1st deliberate charge. The discharge
capacity (C1) after the 1st deliberate charge is significantly increased
by hot alkaline treatment. C0 is a direct measure of alloy corrosion during
the hot alkaline treatment and should be minimized in practice. C1 is a
measure of electrochem. activation rate and should be maximized in
practice. Another important parameter is the maximum capacity Cmax. The
effects of treatment time and temperature, as well as solution (KOH vs. NaOH)
on
C0, C1 and Cmax are presented, along with some long-time cyclic capacity
data.

L53 ANSWER 38 OF 46 HCPLUS COPYRIGHT 2006 ACS on STN
ACCESSION NUMBER: 1995:649765 HCPLUS
DOCUMENT NUMBER: 123:95900
TITLE: The effects of F-treatment level on the
electrochemical behavior of LaNi4.7Al0.3
AUTHOR(S): Li, Z. P.; Yan, D. Y.; Suda, S.
CORPORATE SOURCE: Department Chemical Engineering, Kogakuin University,
Hachijio, 192, Japan
SOURCE: Transactions of the Materials Research Society of
Japan (1994), 18B, 1209-12
CODEN: TMRJE3; ISSN: 1382-3469
DOCUMENT TYPE: Journal
LANGUAGE: English
AB The F-treatment has good effects on improving the activation of
LaNi4.7Al0.3 electrodes. Short-term treatments are good for higher
capacities and lower charge potentials. It also can improve the electrode
kinetic properties. Appropriate F-treatments can improve the
galvanostatic cycling durabilities of the LaNi4.7Al0.3 electrodes without
decreasing their kinetic properties. A phys. model is proposed to explain
these results.

L53 ANSWER 39 OF 46 HCPLUS COPYRIGHT 2006 ACS on STN
ACCESSION NUMBER: 1995:487772 HCPLUS
DOCUMENT NUMBER: 122:295255
TITLE: The fluoride protection mechanism during CaNi5
electrode cycling in the F- ion containing KOH
electrolyte
AUTHOR(S): Li, ZhouPeng; Matsuoka, Takashi; Sandrock,
Gary; Uchida, Masaki; Suda, Seijirau
CORPORATE SOURCE: Kogakuin Univ., Tokyo, 163-91, Japan
SOURCE: Kogakuin Daigaku Kenkyu Hokoku (1994), 77, 77-9
CODEN: KDKHAY; ISSN: 0368-5098
PUBLISHER: Kogakuin Daigaku
DOCUMENT TYPE: Journal
LANGUAGE: Japanese
AB F- ion addition in KOH electrolyte can improve the cycling durability of
CaNi5 electrode dramatically. Using phys. and electrochem. analyses, the
durability improvement mechanism was investigated. The fluoride layers
are formed on the surfaces of CaNi5 particles when discharging in
F--containing KOH solution. The produced CaF2 film is more protective than the
Ca(OH)2 formed during cycling in KOH solution; therefore, the cycling
durability of CaNi5 electrode was improved.

L53 ANSWER 40 OF 46 HCPLUS COPYRIGHT 2006 ACS on STN
ACCESSION NUMBER: 1995:487771 HCPLUS
DOCUMENT NUMBER: 123:37119
TITLE: The effect of element substitution in CaNi5 on the durability improvement of CaNi5-type electrodes
AUTHOR(S): Li, ZhouPeng; Kanazashi, Masato; Sandrock, Gary; Uchida, Masaki; Suda, Seijirau
CORPORATE SOURCE: Kogakuin Univ., Tokyo, 163-91, Japan
SOURCE: Kogakuin Daigaku Kenkyu Hokoku (1994), 77, 71-5
CODEN: KDKHAY; ISSN: 0368-5098
PUBLISHER: Kogakuin Daigaku
DOCUMENT TYPE: Journal
LANGUAGE: Japanese
AB The effects of Ni substitution with Al and Co as well as Ca substitution with La in CaNi5 on the durability improvement when cycling in KOH solution were investigated. Co substitution improved the cycling durability in late cycles. Ca substitution with La also improved the cycling durability in KOH solution appreciably. However, Al substitution has no effect on the durability improvement. The structural and metallog. analyses of these alloys have been done in order to understand the durability improvement mechanism.

L53 ANSWER 41 OF 46 HCPLUS COPYRIGHT 2006 ACS on STN
ACCESSION NUMBER: 1995:469996 HCPLUS
DOCUMENT NUMBER: 122:269997
TITLE: Electrochemical durability of Ca-based alloys
AUTHOR(S): Li, Z. P.; Suda, S.
CORPORATE SOURCE: Dep. Chem. Eng., Kogakuin Univ. Nakano, Tokyo, 192,
Japan
SOURCE: Electrochimica Acta (1995), 40(4), 467-71
CODEN: ELCAAV; ISSN: 0013-4686
PUBLISHER: Elsevier
DOCUMENT TYPE: Journal
LANGUAGE: English
AB In order to improve the cycling durability of Ca-based alloy in KOH solution as hydride electrode in Ni-H battery, we investigated the substitution effects by using Al and Co to substitute nickel as well as alkaline earth metals and rare earth elements to calcium in CaNi₅. It was found that rare earth and alkaline earth substitution can significantly improve the cycling durability during electrochem. charging and discharging. Comparing with A-site substitution in AB₅, B-site substitution by Al or Co shows little improvement on the durability. The structural and metallog. analyses have been done to understand the substitution effects.

L53 ANSWER 42 OF 46 HCPLUS COPYRIGHT 2006 ACS on STN
ACCESSION NUMBER: 1995:357547 HCPLUS
DOCUMENT NUMBER: 122:138069
TITLE: The inorganic compound coating effect on the durability of CaNi5 electrode
AUTHOR(S): Li, Z. P.; Suda, S.
CORPORATE SOURCE: Dep. Chem. Eng., Kogakuin Univ., Tokyo, 192, Japan
SOURCE: Proceedings - Electrochemical Society (1994), 94-27(Hydrogen and Metal Hydride Batteries), 78-84
CODEN: PESODO; ISSN: 0161-6374
PUBLISHER: Electrochemical Society
DOCUMENT TYPE: Journal
LANGUAGE: English
AB F- ion addition in KOH electrolyte can improve the electrochem. cycling durability of CaNi5 electrode dramatically. The durability improvement mechanism was investigated through phys. and electrochem. analyses. Fluoride layers will be formed on the surfaces of CaNi5 particles during cycling CaNi5 electrode in F- ion containing KOH electrolyte. The formed calcium fluoride is more protective against the corrosion of strong alkaline solution than calcium hydroxide that was formed when the electrode was cycling in pure KOH electrolyte. With increasing the F- ion content in the electrolyte, the durability of CaNi5 electrode can be improved very much but the discharge capacity was decreased because of the ion conductivity loss.

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L53 ANSWER 43 OF 46 HCPLUS COPYRIGHT 2006 ACS on STN
ACCESSION NUMBER: 1995:357541 HCPLUS
DOCUMENT NUMBER: 122:138065
TITLE: The metallurgical and electrochemical investigations
of Ca-La-(NiAl)₅ alloys
AUTHOR(S): Li, Z. P.; Suda, S.
CORPORATE SOURCE: Dep. Chem. Eng., Kogakuin Univ., Tokyo, 192, Japan
SOURCE: Proceedings - Electrochemical Society (1994),
94-27 (Hydrogen and Metal Hydride Batteries), 16-23
CODEN: PESODO; ISSN: 0161-6374
PUBLISHER: Electrochemical Society
DOCUMENT TYPE: Journal
LANGUAGE: English

AB The effects of lanthanum substitution for calcium in Ca_{1-x}La_xNi_{4.7}Al_{0.3} on metallog., structure and electrochem. properties. have been investigated. It was found that cored dendritic structure existed and some segregation of Ca and La occurred on solidification. Lanthanum substitution can dramatically decrease the amount of A₂B₇ phases, which was caused by high vapor pressure of Ca and non-equilibrium solidification of the alloy. From the results of X-ray lattice parameter measurement, the compactness of Ca-La-(NiAl)₅ alloys linearly increases with increasing the La substitution amount. La substitution can obviously improve the electrochem. durability of the alloy by lattice stabilization and anti-corrosion capability improvement in strong alkaline solution

L53 ANSWER 44 OF 46 HCPLUS COPYRIGHT 2006 ACS on STN
ACCESSION NUMBER: 1994:609064 HCPLUS
DOCUMENT NUMBER: 121:209064
TITLE: Rate capacity determination mechanism of F-treated
LaNi4.7Al0.3 electrode
AUTHOR(S): Li, Zhoupeng; Yan, Deyi; Suda,
Seijirau
CORPORATE SOURCE: Zhejiang Univ., Hangzhou, Peop. Rep. China
SOURCE: Kogakuin Daigaku Kenkyu Hokoku (1993), 75, 113-22
CODEN: KDKHAY; ISSN: 0368-5098
DOCUMENT TYPE: Journal
LANGUAGE: English
AB A math. model of a rate capacity determination mechanism of LaNi4.7Al0.3 as a
hydride electrode is established, based on the hydride electrode kinetics
and capacity degradation mechanism. Rate cycling behavior of the F-treated
electrode was investigated. Treatment with F enhances its durability.

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L53 ANSWER 45 OF 46 HCPLUS COPYRIGHT 2006 ACS on STN
ACCESSION NUMBER: 1994:609063 HCPLUS
DOCUMENT NUMBER: 121:209063
TITLE: Effects of F treatment on durability of LaNi_{4.7}Al_{0.3}
electrode
AUTHOR(S): Li, Zhoupeng; Yan, Deyi; Suda,
Seijirau
CORPORATE SOURCE: Zhejiang Univ., Hangzhou, Peop. Rep. China
SOURCE: Kogakuin Daigaku Kenkyu Hokoku (1993), 75, 107-12
CODEN: KDKHAY; ISSN: 0368-5098
DOCUMENT TYPE: Journal
LANGUAGE: English
AB Treatment of LaNi_{4.7}Al_{0.3} as a hydride electrode with F to coat it with fluoride enhances its charge-discharge cycle durability. However, the coating layer is finely cracked, when exposed to the cycles for an extended period, which decreases capacity of the electrode.

L53 ANSWER 46 OF 46 HCPLUS COPYRIGHT 2006 ACS on STN
ACCESSION NUMBER: 1994:609062 HCPLUS
DOCUMENT NUMBER: 121:209062
TITLE: Electrode kinetics of F-treated LaNi_{4.7}Al_{0.3}
electrodes
AUTHOR(S): Li, Zhoupeng; Yan, Deyi; Suda,
Seijirau
CORPORATE SOURCE: Zhejiang Univ., Hangzhou, Peop. Rep. China
SOURCE: Kogakuin Daigaku Kenkyu Hokoku (1993), 75, 99-106
CODEN: KDKHAY; ISSN: 0368-5098
DOCUMENT TYPE: Journal
LANGUAGE: English
AB An attempt was made to correlate rate capacity of LaNi_{4.7}Al_{0.3} as a
hydride electrode with discharge c.d. The electrode reaction is
controlled by the electrochem. reaction occurring on the catalytic sites
in a high current discharge region. Treatment with F enhances kinetic
characteristics of the electrode.

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